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Reaction of Acetylenedicarboxylic Acid Made Easy: High-Pressure Route for Polymerization

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Supporting Information

ABSTRACT: A breakthrough has been achieved in improving the efficiency of solid-state polymerization of acetylenedicarboxylic acid (ADCA). Traditional solid-state polymerization of ADCA is marked by long exposure times of γ-radiation (>10 days) and very low yields (around 5.5%). We have been able to perform a reaction to an n = 8 oligomer, as confirmed by matrix-assisted laser desorption/ionization-time of flight, in less than 2 min by employing ~6 GPa of pressure. We have determined the crystal structure of ADCA on increasing pressure to (5.2 GPa) to provide insight into the process of polymerization with Pixel calculations supporting our evaluation of the polymerization process.

INTRODUCTION

Organic polymers play an important role in day-to-day life (e.g., polyethylene, nylon, Teflon, etc.). Solid state synthesis of polymers is of great interest, as the polymers synthesized by this route might have different architectures and properties compared to those obtained via solution methods. Solid-state synthesis of acetylenes is of special interest as it may selectively form conducting polyacetylene polymers1−4 that are used in a range of electronics as much needed components for electronics (e.g., solar cells, organic light emitting diodes etc.). A polymer of acetylene is limited in terms of its solubility and functionality, i.e., lack of functional groups on backbone, and scientists are investigating the formation of various derivatives of polyacetylene using functionalized acetylenes as monomers. One such monomer of interest is the acetylenedicarboxylic acid (ADCA), whose solid-state polymerization is considered a problematic task. Usanmuz et al. attempted solid-state polymerization of ADCA by irradiating it with Co-60 γ-rays. No polymer was formed in the first 100 h over which the sample was exposed to a radiation dose of approximately 2 Mrad. They were able to convert about 5.5% of monomer to oligomer, as confirmed by matrix-assisted laser desorption/ionization-time of flight, in less than 2 min by employing ~6 GPa of pressure. We have determined the crystal structure of ADCA on increasing pressure to (5.2 GPa) to provide insight into the process of polymerization with Pixel calculations supporting our evaluation of the polymerization process.
solid-state reactions came from Schmidt who proposed the topochemical mechanism based around the distances between the reaction centers. The rule states that the centers should be between 3.6 and 4.1 Å before the reaction can take place.38,39 Other theories based on molecular motion were proposed by Kaupp using atomic force microscopy (AFM) to account for the observations that did not obey Schmidt’s Rule.40

Herein, we discuss the structural changes that occur in ADCA before reaction to the polymer product as well as providing insight into the product of the reaction using matrix-assisted laser desorption/ionization-time of flight (MALDI-TOF).

**EXPERIMENTAL SECTION**

**Materials Used.** ADCA was purchased from Sigma-Aldrich and was recrystallized from acetone. Petroleum ether (35:60) (Alfa Aesar) was used as a pressure-transmitting medium in the high-pressure experiments.

**High-Pressure Experiments.** High-pressure studies were performed using a Merrill-Bassett diamond anvil cell (DAC) with 600 μm culet diamonds glued onto tungsten carbide backing disks. A 250 μm thick tungsten foil gasket was indented to ~90 μm, and a 200 μm hole was drilled in the indented portion of the gasket to create a sample chamber. A few rubies were placed in the sample chamber to allow in situ pressure measurement via the ruby fluorescence technique.41

For the structure determination by single-crystal X-ray diffraction, we loaded three crystals in different orientations simultaneously together with the rubies. We saw an improvement of 30% from 54% to 74% for pressure points 4.29 and 4.85 GPa (55% completeness to 69% after merging). OLEX2 was used to refine the low temperature model to the high-pressure data sets.33 All non-hydrogen atoms were anisotropically refined, and the hydrogen atom was geometrically placed on both oxygen atoms with their occupancies fixed at the values obtained from the single-crystal experiment described above, and coordinates were allowed to optimize. Convergence was defined with medium quality, i.e., when the maximum changes in total energy, displacement, and gradient were 2 × 10−3, 5 × 10−3 Å, and 4 × 10−3 Ha Å−1, respectively. Brillouin zone integrations were performed by Monkhorst-Pack k-point sampling at intervals of 0.08 Å−1.

**Geometry optimizations of Form III** were performed by periodic density functional theory (DFT) using the DMOL3 code46 as part of the Materials Studio modeling suite.47 The DND numerical basis set46 was used in combination with the PBE functional48 with the Tkatchenko-Scheffler correction for dispersion.49 The unit cell dimensions were held fixed at the values obtained from the single-crystal experiment described above, and coordinates were allowed to optimize. Convergence was defined with medium quality, i.e., when the maximum changes in total energy, displacement, and gradient were 2 × 10−5 Ha, 5 × 10−3 Å, and 4 × 10−3 Ha Å−1, respectively. Brillouin zone integrations were performed by Monkhorst-Pack k-point sampling at intervals of 0.08 Å−1.

**Raman Spectroscopy.** Raman spectra were collected on a Horiba Scientific Jobin Yvon Raman XploRA microscope or a ThermoFisher DXR microscope fitted with a 532 nm laser to allow monitoring of both the Raman spectrum and the fluorescence from the ruby. The latter Raman microscope required the use of an extended range grating to observe the ruby fluorescence. Spectra were analyzed using LabSpec6 or Omnic and Origin2015.

**Single-Crystal X-ray Diffraction.** X-ray diffraction intensities were collected using a Bruker Apex II diffractometer with an Incoatec IμS microsource (λ = 0.71073 Å). The diffraction from the three crystals was separated using a reciprocal lattice determination tool and indexed with three separate orientation matrices. The data for each crystal were integrated separately using SAINT in APEX3 software utilizing the in-built dynamic masking procedure. Absorption correction was conducted with SADABS.50 After data-processing the data for each crystal at each pressure point were merged in XPREP, culminating in a 30% improvement in the data completeness; only two crystals were used for pressure points 4.29 and 4.85 GPa (55% completeness to 69% after merging). OLEX2 was used to refine the low temperature model to the high-pressure data sets.33 All non-hydrogen atoms were anisotropically refined, and the hydrogen atom was geometrically placed on both oxygen atoms with their occupancies set to 0.5, which was identified in previous work44,45 and supported by our observations. For each data set, some data have been omitted due to overlap with diamond or the beamstop. Crystallographic tables can be found in Table S1I of the Supporting Information.

**Periodic DFT Calculations.** Geometry optimizations of Form III were performed by periodic density functional theory (DFT) using the DMOL3 code46 as part of the Materials Studio modeling suite.47 The DND numerical basis set46 was used in combination with the PBE functional48 with the Tkatchenko-Scheffler correction for dispersion.49 The unit cell dimensions were held fixed at the values obtained from the single-crystal experiment described above, and coordinates were allowed to optimize. Convergence was defined with medium quality, i.e., when the maximum changes in total energy, displacement, and gradient were 2 × 10−3, 5 × 10−3 Å, and 4 × 10−3 Ha Å−1, respectively. Brillouin zone integrations were performed by Monkhorst-Pack k-point sampling at intervals of 0.08 Å−1.

**Pixel Calculations.** Pixel calculations, as incorporated in the CLP suite of programs, were performed on each data set.51,52 Ordered models were used for this process; therefore we kept the hydrogen atoms attached to the oxygen with more single bond character. Electron densities were calculated using Gaussian09 at the MP2/6-31G** level. The condensation level was set at the default, 4. Visualization of the interactions was made possible through the use of processPixel.54 The energies can be found in Tables S1I–S9 in the Supporting Information.
**RESULTS AND DISCUSSION**

ADCA is one of the simplest alkyne molecules that possess hydrogen-bonding capabilities. A polymer of ADCA will contain many carboxylic acid functional groups and hence may find applications for the delivery of chemicals such as pharmaceuticals, pigments, agrochemicals, etc. The carboxylic acid groups of ADCA are disordered leading to C—O distances that are shorter and C=C—O distances slightly longer than normal which confirms previous literature reports. Our refinements fix the occupancy of the hydrogen atoms to be 50%. ADCA forms carboxylic acid dimers (IN0) that continue into chains through the structure along the [1 0 1] direction (Figure 1). These chains form a layer that is offset from those above and below giving a close-packed structure when viewed down the hydrogen-bonded chain length (Figure 1b,d). Each of these chains is surrounded by six neighboring chains (IN1—INS) where the acetylene groups of the molecules are in close proximity to one another, giving rise to the potential formation of a polymer through compression of these interactions (Figure 1d). At ambient pressure the distance between the acetylene groups is approximately 4 Å, which is at the top end for reactive center separation proposed by Schmidt rule for solid-state reactions.

In his work, Schmidt realized the importance of distance between the polymerizable groups coining the “Schmidt Rule”. According to this rule, for a photodimerization reaction the phase transition is only clear due to the discontinuity in the pressure vs volume graph (Figure 3a). The equation of state of Form I could not be refined due to the lack of data points, but Form II shows a compressibility that is commensurate with other organic substances (K’ 14.5(8), V0 450(3) Å3 using a second order Birch–Murnaghan equation of state). The compression of Form II is anisotropic with the c-axis, the least compressible (2.8%) followed by the a- (8.7%) and then the b-axis (11.1%). This can be rationalized from the orientation of the hydrogen bonded chains with respect to the
axes. The direction of the chains [1 0 1] has components of both the $a$- and $c$-axis; hence the compression along these axes would necessitate the compression of the molecule in these directions. There are no strong hydrogen bonding interactions along the $b$-axis, hence it is the most compressible direction.

For solid-state reactions to occur, the packing of the molecules is crucial and in particular the distance between reaction centers. In the crystal structure of ADCA, it is surrounded by six neighbors, with which it can interact for polymerization using its carbon−carbon triple bond. We have labeled these as interactions 0−5 (IN0-5; Figure 1d). The centroids of the molecules, as used by Pixel, lie on the carbon triple bond and hence serve as a good estimate of the distances between the centers required for polymerization to occur. Figure 4 shows how the energies of interaction change as a function of distance.

The interaction between molecules in the dimer (Int. 0) is the strongest interaction dominated by Coulombic and polarization terms that steadily becomes more repulsive as pressure is applied. The interactions that have the greatest rate of compression are Int. 2 and 3, which are more dispersive in character as the molecules interact using van der Waals interactions; these two interactions can be seen in Figure 1d. As we reach the point of reaction the distance of the molecules tends toward 3.5 Å, which is close to the ideal distance for polymerization as commented by Schmidt (3.58 Å and 3.61 Å, respectively), which suggests that these two interactions may be those to initiate that polymerization. Between 4.85 and 5.2 GPa there is a transition to Form III which we were able to identify. The refinement statistics are not as good as the previous data sets; however we have been able to identify a C-centered cell and solve the structure with $Z' = 1/2$. The residual electron density from this refinement (located at (0.5118, 0.4386, 0.5828)) is higher than normal and is observed at right angles to the molecules near the acetylene moiety. The direction of the density is toward the molecules involved in Int. 2 and 3 rather that those involved in Int. 5 and 6. This supports our observations that these two interactions may be where the reaction is initiated. To understand the changes over the transition, we geometry-optimized the structure at 5.2 GPa before performing the Pixel calculations. This was, in part, due to the quality of the data set at 5.2 GPa but also because the refined structure had $Z' = 1/2$, which may have the unintended problem of distorting the molecule due to symmetry constraints rather than through physical changes. Pixel requires a full molecule to compute the energies, and hence we lowered the symmetry back to $P2_1/n$ for this purpose. In Form III we see that the lateral interactions 2, 3, 5, and 6 all continue to compress but show a decrease in energy of $-2$ to $-4$ kJ mol$^{-1}$ to be more stabilizing. However, it is at the repulsive end of the interactions that shows a very large change. Throughout the compression the interaction of the central molecule with one transformed by $1/2 + x$, $3/2 - y$, $1/2 + z$ shows a steady destabilization but jumps significantly, by $10$ kJ mol$^{-1}$, from $6$ kJ mol$^{-1}$ to $16$ kJ mol$^{-1}$ due to the increase in repulsive energy. This interaction (Figure 5) shows the increasing proximity of the carbonyl groups of the carboxylic acid (2.607 Å, from geometry optimized structure). In the original study by Usanmuz et al. they suggested that CO or CO$_2$ is lost from the structure during the polymerization. The rapid change in repulsion energy between molecules in this interaction may signify the CO/CO$_2$ loss mechanism. Why the molecules required this amount of force to polymerize may be found in...
the work of Kaupp who proposed that it was the motion of the molecules not the distance between centers that causes the reaction to occur. With molecules able to move easily (due to the chain structure), we are able to push the molecules closer together and into an arrangement that presets it for the transition-state of the reaction, and hence it facilitates the reaction.

Diamond cells are a very powerful tool for the characterization of materials under high-pressure conditions; however the volume of sample is limited for any subsequent analysis unless it is via spectroscopic or diffraction techniques. Hence we turned to the use of a Paris-Edinburgh press for the compression of a larger volume of sample so that we may analyze the product of the reaction. Two experiments were conducted, with and without pressure-transmitting medium. During one of the compression runs with hydrostatic fluid, we noted a large and sudden decrease in pressure as we reached the highest load of 60 tonnes (∼3.5 GPa) and collapse of the gasket signifying that a large reduction in volume occurred. The compression without fluid also lost pressure but over a greater period of time signifying the reaction. These samples were subsequently characterized using MALDI-TOF which can provide information on the size of the polymers produced; a typical spectrum from these runs is shown in Figure 6. The largest mass fragments were observed at 885.35, 907.33, and 919.59 m/z; lower mass fragments were avoided due to the contribution from the matrix in which the sample sits. The mass at 885.45 m/z corresponds to [M₈−CO + H]$^+$ given the mass of the monomer is 114.06 m/z. The mass at 907.33 m/z corresponds to [M₈−CO + Na]$^+$, while the mass at 919.59 corresponds to [M₈−O + Na]$^+$. We did not observe any intact [M₈ + H]$^+$ fragments, and our results indicate that the ADCA has formed oligomers comprised of eight monomer units but with loss of CO and O. It may be that heavier fragments did not readily ionize and were therefore not detected. Our observation from the MALDI-TOF that there is loss of CO does match the observations of Usanmaz et al. but via a different reaction route.

The color of the sample from both the diamond cell experiment and large volume experiment may also give us clues to the structure of the oligomers. In the seminal work of Shirikawa on conducting polymers, they noticed that depending on the conformation of the ethylene groups in the polymer, i.e., trans and cis, the color of the polymer was either silver or copper colored. The oligomers that we have produced do show either copper or brown hue rather than the metallic luster, which suggests that these are of the cis form. This corresponds with our observations that Int. 2 and 3 are involved in the polymerization. If we look at the spatial arrangement of the molecules in Figure 1d, then we can see that by polymerizing in this manner we will end up with the cis rather than trans isomers and hence a copper-colored product.

**CONCLUSION**

We have shown that we have been able to induce reaction in ADCA without the use of initiator or significant doses of radiation. The structural information that we have gained through compression has shown that the acetylene moieties are compressed to ∼3 Å before the reaction occurs. The product of the reaction has been shown to be an oligomer of n = 8 with loss of CO and O; however the method of detection may preclude the identification of higher mass fragments. From the color of the samples, we argue that the conformation of the polymer products is of the cis form rather than the trans form.

**ASSOCIATED CONTENT**

Supporting Information
The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.cgd.7b01338.

Crystallographic tables and tables of Pixel energies (PDF)

Accession Codes
CCDC 1574461–1574473 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

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**Author Contributions**

A.D. collected the data sets and produced initial analysis and draft of manuscript. A.J.U. performed the MALDI-TOF experiments and contributed to the manuscript. I.B.H. helped with the high-pressure work. C.L.B. and N.P.F. performed the large volume experiments on ADCA. I.D.H.O. completed the analysis and wrote the final manuscript.
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Notes
The authors declare no competing financial interest.
All data underpinning this publication are openly available from the University of Strathclyde KnowledgeBase at http://dx.doi.org/10.15129/06d23e6-b46e-4f6a-bdb7-df5a6ad30bc3.

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Abbreviations
ADCA, acetylenedicarboxylic acid; MALDI-TOF, matrix-assisted laser desorption/ionization-time of flight

References

