The Importance of Surface IrO\textsubscript{x} in Stabilizing RuO\textsubscript{2} for Oxygen Evolution

María Escudero-Escribano,*,1,2,3 Anders F. Pedersen,2 Elisa A. Paoli,2 Rasmus Frydendal,2 Daniel Friebel,3 Paolo Malacrida,2 Jan Rossmeisl,1 Ifan E. L. Stephens,*,2,4 Ib Chorkendorff*,2

1Nano-Science Centre, Department of Chemistry, University of Copenhagen, Universitetsparken 5, DK-2100 Copenhagen, Denmark
2Department of Physics, Fysikvej, Building 312, Technical University of Denmark (DTU), DK-2800 Kgs. Lyngby, Denmark
3SUNCAT Center for Interface Science and Catalysis, Department of Chemical Engineering, Stanford University, 443 Via Ortega, Stanford, CA 94305, United States
4Department of Materials, Imperial College London, 2.03b, Royal School of Mines, London SW72AZ, England

ABSTRACT:

The high precious metal loading and high overpotential of the oxygen evolution reaction (OER) prevents the widespread utilization of polymer electrolyte membrane (PEM) water electrolyzers. Herein we explore the OER activity and stability in acidic electrolyte of a combined IrO\textsubscript{x}/RuO\textsubscript{2} system consisting of RuO\textsubscript{2} thin films with sub-monolayer (1, 2 and 4 Å) amounts of IrO\textsubscript{x} deposited on top. Operando extended X-ray absorption fine structure (EXAFS) on the Ir L-3 edge revealed a rutile type IrO\textsubscript{2} structure with some Ir sites occupied by Ru, IrO\textsubscript{x} being at the surface of the RuO\textsubscript{2} thin film. We monitor corrosion on IrO\textsubscript{x}/RuO\textsubscript{2} thin films by combining electrochemical quartz crystal microbalance (EQCM) with inductively coupled mass spectrometry (ICP-MS). We elucidate the importance of sub-monolayer surface IrO\textsubscript{x} in minimizing Ru dissolution. Our work shows that we can tune the surface properties of active OER catalysts such as RuO\textsubscript{2}, aiming to achieve higher electrocatalytic stability in PEM electrolyzers.
1. INTRODUCTION

There is an urgent need to develop a sustainable economy based on renewable energy.\textsuperscript{1,2} Electrocatalytic reactions taking place at energy conversion devices such as fuel cells and electrolyzers may play a key role in the transition towards a sustainable future.\textsuperscript{3} In particular, electrolysis of water allows the production of renewable hydrogen as an energy carrier.\textsuperscript{4–6} Polymer electrolyte membrane (PEM) electrolyzers are highly attractive due to their high efficiency, as they can operate at much higher current densities (1-3 A cm\textsuperscript{-2}) than traditional alkaline electrolyzers (around 0.2 A cm\textsuperscript{-2}).\textsuperscript{7–9} Furthermore, they are more amenable towards the small scale storage intermittent electricity from renewable resources. However, due to the acidic conditions at which PEM electrolyzers operate, the electrode materials at both anode and cathode are based on precious metals like Pt and Ir.\textsuperscript{7}

The slow kinetics of the oxygen evolution reaction (OER) at the electrolyzer anode causes significant potential losses (overpotential).\textsuperscript{10,11} To date, only oxides based on Ir and Ru present reasonable activity and stability for OER in acidic electrolyte.\textsuperscript{12–18} High loadings of precious metal electrocatalysts are needed to minimize the overpotential, preventing the widespread utilization of PEM electrolyzers. In contrast, the cathodic reaction, the hydrogen evolution reaction (HER), has been widely investigated in the past decade. Negligible loadings of Pt can sustain enormous hydrogen evolution currents;\textsuperscript{19–22} and numerous electrode materials have been developed, including highly active non-precious metal catalysts based on sulfides and phosphides,\textsuperscript{23,24} although there are some stability issues here as well.\textsuperscript{25}

In order to minimize the overpotential in oxygen electrocatalysis and design more efficient materials, we need to understand the activity and stability descriptors.\textsuperscript{4,26–29} The activity and stability of RuO\textsubscript{2}, the most active OER electrocatalyst, can vary dramatically as a function of the oxidation pre-treatment. Electrochemically (anodically) grown or plasma treated RuO\textsubscript{2} catalysts are typically highly active for OER; however, they are less stable than thermally grown oxides.\textsuperscript{30–32} The stability of IrO\textsubscript{x} depends strongly on the synthesis conditions as well.\textsuperscript{33} Typically, IrO\textsubscript{x} has a lower catalytic activity than RuO\textsubscript{x}; even so, it is less prone to corrosion.\textsuperscript{34–38} This prevents the use of pure RuO\textsubscript{x} at the PEM electrolyzer anode, IrO\textsubscript{x} being the catalyst of choice for large scale applications.\textsuperscript{7,39} Enhancing the catalytic stability is crucial to maximize the durability of PEM electrolyzers.

In the 1950s, Beer discovered efficient and stable transition metal oxides, the so-called dimensionally stable anodes (DSA ®),\textsuperscript{40} which consisted of thermally decomposed oxides mixed with an activated metal, mainly Ir and Ru, and deposited on an inert substrate, typically Ti. After the DSA ® success,\textsuperscript{41} researchers have tried different combinations of metal oxides in order to find the optimal electrocatalyst composition.\textsuperscript{5} In particular, both adding IrO\textsubscript{x} to RuO\textsubscript{x} and forming Ru-Ir bimetallic oxides can result in a stability increase over RuO\textsubscript{x}.\textsuperscript{36,42–46} The activity and stability of Ru-Ir oxides are also very sensitive to the preparation method, since the surface enrichment of Ir and homogeneity of the phases varies depending on the preparation procedure.\textsuperscript{34,43–45}

Experimental studies on Ru-Ir mixed oxides show that Ru at the surface tends to dissolve, leaving an enriched Ir layer behind.\textsuperscript{36,47,48} However, the nature of the stabilizing effect is not fully understood yet.
Thermodynamic stabilization under OER conditions is challenging.\textsuperscript{49} Consequently, Ir could provide kinetic stability against further dissolution of RuO$_2$. The higher stability of IrO$_x$ as compared to RuO$_x$ have led theorists to study the surface segregation and migration energies of Ru-Ir mixed oxides and investigate alternative strategies to obtain highly active and stable OER catalysts.\textsuperscript{50} DFT calculations suggest that the active sites for OER on a rutile RuO$_2$(110) surface are terrace sites\textsuperscript{26} and dissolution takes place at under-coordinated sites. The calculated surface energies suggest that IrO$_x$ tends to sit on the surface under equilibrium conditions, since it has a lower surface energy. Moreover, Ir is most stable at the step edge sites of a stepped RuO$_2$(110) surface. This configuration is more stable than pure RuO$_2$ because of the higher dissolution potential of IrO$_2$.\textsuperscript{51} This is in agreement with the work by Danilovic et al.: they vacuum annealed and thermally oxidized a bulk metallic Ir-Ru alloy; this treatment resulted in an Ir-rich oxide surface which was stable under oxygen evolution conditions.\textsuperscript{36} A similar effect was recently observed on SrIrO$_3$,\textsuperscript{16} where an IrO$_x$ overlayer was formed by leaching of Sr. These results are consistent with Man’s proposal\textsuperscript{50} that surface IrO$_x$ can stabilise a less stable oxide, which resides in the bulk. Even so, preparing oxides that are mixed in the bulk does not allow the precise control of the surface composition.

Understanding degradation mechanisms and quantifying the stability improvement of Ru-Ir mixed oxides is challenging. Thus far, the most common methods to study the stability have been short term electrochemical measurements, such as chronopotentiometry (at a constant current density) or chronoamperometry (at a constant potential) measurements. However, these tests cannot allow an estimation of the long-term catalytic performance.\textsuperscript{52,53} In contrast, combining electrochemical methods with \textit{in situ} and \textit{operando} studies can offer valuable information on the catalyst stability.\textsuperscript{31,47,53,54} Electrochemical quartz crystal microbalance (EQCM) allows following mass changes at the electrode.\textsuperscript{53} In addition, on-line elemental analysis can be performed by combining inductive couple plasma mass spectrometry (ICP-MS) with a scanning flow cell.\textsuperscript{35,55}

\textit{In situ} and \textit{operando} spectroscopy is particularly suitable to study corrosion mechanisms and monitoring the reaction products under dynamic conditions.\textsuperscript{54,56} Synchrotron based X-ray techniques have become an essential tool to elucidate solid catalysts,\textsuperscript{57–60} Miquel Salmerón is a pioneer in this area.\textsuperscript{54,61} In particular, (near) ambient pressure X-ray photoelectron spectroscopy (AP-XPS) and \textit{operando} X-ray absorption spectroscopy (XAS) allow \textit{operando} investigations of the chemical state and structure of the electrode surface and the reaction intermediates in oxygen electrocatalysis.\textsuperscript{47,59,62,63}

This article combines EQCM and ICP-MS measurements to monitor mass losses and investigate the corrosion rate of sputtered RuO$_2$ thin films decorated by sub-monolayer (1, 2 and 4 Å) amounts of IrO$_x$. Recently, we used a very similar approach to mildly stabilize MnO$_x$ films in acid under oxygen evolution conditions, using surface TiO$_x$.\textsuperscript{64} Herein, we have also used \textit{operando} grazing incidence extended X-ray absorption fine structure (GI-EXAFS) on the Ir L$_3$ edge to show that the IrO$_x$ resides at the surface of the RuO$_2$ thin film. The addition of sub-monolayer surface IrO$_x$ improves the stability of RuO$_2$ thin films.
2. EXPERIMENTAL SECTION

2.1. Preparation of thin films

The hybrid IrO$_x$/RuO$_2$ thin films, as well as 40 nm thin films of pure RuO$_2$ and IrO$_2$ were fabricated by magnetron sputtering. Sub-monolayer amounts of IrO$_x$ were deposited on top of a 40 nm RuO$_x$ thin film. Both films were deposited by sputtering of a Ru or Ir target at 30 W with the sample substrate being held at 300°C. The oxide was formed during sputtering by a reactive atmosphere at 3 mTorr and a flow of 10 sccm Ar and 4 sccm O$_2$. For IrO$_2$ thin films, a pressure of 5 mTorr, and a flow of 20 sccm, of Ar and 5 sccm of O$_2$ was used, in order to assure the formation of crystalline IrO$_2$. The sputter rate was monitored by a quartz crystal microbalance (QCM).

For the stability measurements, 40 nm RuO$_2$ thin films were deposited on an Au-coated electrochemical quartz crystal microbalance (EQCM), followed by the deposition of 1 Å, 2 Å, or 4 Å of IrO$_x$. After the deposition, the substrate was held at 300°C for 30 minutes for annealing. For X-ray diffraction (XRD) characterization, the RuO$_2$ and IrO$_2$ thin films were deposited on a glass substrate, while for X-ray absorption spectroscopy (XAS) measurements the thin films were deposited on polycrystalline Au disks.

2.2. Characterization methods

The 40 nm RuO$_2$ thin film was structurally characterized by glancing angle X-ray diffraction (GAXRD) from a sample film on a glass substrate. A PAN Analytical X’Pert PRO Diffractometer was used, equipped with a Cu Kα X-ray source with a wavelength of 1.54 Å and a monochromator on the detector. The incidence angle was set to 0.5°. The thin film morphology was characterized by scanning electron microscopy (SEM) made with an acceleration voltage of 5.00 kV and secondary electron detection using a Helios EBS3 microscope. IrO$_x$/RuO$_x$ thin films were examined by X-ray photoelectron spectroscopy (XPS). A Theta Probe system by Thermo Fischer was utilized, with a monochromatized Al K$_\alpha$ X-ray source. It is equipped with a hemispherical analyzer, which also allows angle-resolved XPS (AR-XPS) to be measured in the same system.  

2.3. In situ X-ray absorption synchrotron measurements

Synchrotron measurements were carried out at Stanford Synchrotron Radiation Lightsource (SSRL) at SLAC National Accelerator Laboratory in California, USA. Grazing incidence extended X-ray absorption fine structure (GI-EXAFS) measurements were carried out at beam line 11-2, which is a hard X-ray wiggler beam line equipped with a Si(220) monochromator in the $\varphi$=0° setting. A 100 element Ge detector was used for fluorescence detection, while ion chambers were used for measuring initial and transmitted intensities. EXAFS was measured at the Ir L$_3$-edge (11215 eV) on 1 Å and 2 Å IrO$_x$/RuO$_2$/GC samples. The analysis was carried out using the SIXPack analysis software, and the fitting was done using scattering paths from FEFF6 calculations. A $k^2$-weighted $\chi(k)$ function was used for fitting in the range of 3 – 9 Å$^{-1}$ using a Hanning filter. The fitting itself was done in R-space in the range 1.1 – 3.8 Å.
All experiments were done in an in-situ cell, based on a hanging meniscus setup developed by Daniel Friebel and co-workers, which is described elsewhere. The cell had a motorized syringe attached, so that the size of the meniscus could be controlled accurately, forming a cylindrically shaped meniscus 5 mm in diameter. The electrolyte used was 0.05 M H$_2$SO$_4$ with a Pt wire counter electrode and an Ag/AgCl reference electrode (Innovative Instruments, Inc. LF-1.6, 3.4M AgCl). To control the potentials a BioLogic potentiostat (BioLogic Scientific Instruments) was used, which was controlled by a separate PC.

2.4. Electrochemical tests

The electrochemical experiments were performed with a rotating ring disk electrode (RRDE) assembly (Pine Instruments Corporation) in a standard three-electrode glass cell, equipped with a Luggin capillary. All the electrochemical measurements were carried out in 0.05 M H$_2$SO$_4$ (Merck suprapur 96%, diluted with 18 MΩ Millipore water). The counter electrode was a carbon rod, and a Hg/HgSO$_4$ was used as a reference electrode. Activity measurements were done in N$_2$ or Ar saturated electrolyte. The oxygen evolution activity was measured by quasi stationary cyclic voltammetry (CV) in N$_2$-saturated 0.05 M H$_2$SO$_4$ in the EQCM set-up. For the stability measurements, the mass loss was monitored by an EQCM balance, which changes its resonance frequency in response to change in the electrode mass. The dissolution of electrode material was further examined by inductively coupled plasma mass spectroscopy (ICP-MS), in which samples of the electrolyte were analyzed at several stages of the stress test. The experiments were all carried out using a BioLogic VMP2 potentiostat controlled by a PC with EC-lab control software (BioLogic Science Instruments).

3. RESULTS AND DISCUSSION

The aim of this work is to improve the stability of RuO$_2$ by depositing sub-monolayer amounts of IrO$_x$ on top. Pure RuO$_2$ and IrO$_2$ thin films were prepared by reactive sputter deposition. Both thin films were structurally characterized by GA-XRD and SEM (see Supplementary Information, S.I., Figures S1-S4). According to SEM images (Figures S3-S4), IrO$_2$ thin films present rougher surfaces than RuO$_2$ thin films, exhibiting triangular-shape structures which protrude in the z direction. This could be explained by differences in the film growth during the sputter deposition. Stoezinger et al. also observed differences in morphology on epitaxially grown RuO$_2$ and IrO$_2$ films; IrO$_2$ (100) films were almost 10 times rougher than RuO$_2$ (100) films, while IrO$_2$ (110) exhibited 4 times higher roughness than RuO$_2$ (110) films. McCrory et al. observed much higher roughness for electrodeposited IrO$_2$ as well, as compared to other non-noble metal oxides.

Small amounts of IrO$_x$ were sputtered onto the surface RuO$_2$ thin films to form the IrO$_x$/RuO$_2$ surfaces. The nominal deposition thicknesses were 1, 2, 4 and 10 Å. This provides a well-defined and well-characterized system onto which model studies can be carried out in order to understand the activity
and stability trends, as well as quantify the stability enhancement. A monolayer of iridium oxide corresponds to 3.2 Å. Assuming a conservative roughness factor of 2 for RuO₂ thin films, it is unlikely that the surface is completely covered by IrOₓ, even when 4 Å are deposited. Hence we consider 1, 2 and 4 Å thick films to correspond to sub-monolayer additions.

In order to investigate bond distances and coordination numbers of neighboring atoms of the surface, we carried out operando GI-EXAFS measurements on the Ir L₃-edge of IrOₓ/RuO₂ thin films containing 1 and 2 Å of IrOₓ. Both samples were measured as-prepared (“dry”), and the 2 Å IrOₓ/RuO₂ sample was also measured in-situ with an applied potential of 1.0 V vs. RHE. Figure 1 shows the structure resulting from the fit to the EXAFS data, consisting of a rutile type IrOₓ, with Ru occupying some of the Ir sites. Table 1 shows the fitted bond distances as well as the coordination numbers. Debye-Waller factors (σ²), overall energy tuning (∆E₀), amplitude reduction factor (S₀²), and goodness of fit parameters are shown in Table S1 in the Supporting Information (S.I.). Figure S5 shows the k²-weighted χ(k) and χ(R) extracted from the measured EXAFS spectra, as well as the fitted function. The range of k-space used for the Fourier transform is indicated as well as the fitting range in real space. During the fitting process, we were unable to distinguish between Ir and Ru neighbors, and we could not separate bond distances that were very close to each other. To avoid degeneracy between the coordination numbers, amplitude reduction factor, and Debye-Waller factors, the latter two were fixed during the fitting procedure. The Debye-Waller factors were tuned by hand to optimize the fit while maintaining physically reasonable values. The fitting analysis shows that the first coordination shell of O was fully occupied in all cases. The further coordination shells with Ir/Ru and O show a smaller occupancy than bulk rutile IrO₂. These results are consistent with IrOₓ being on the surface of the RuO₂ thin film.

Figure 1. Schematic view of the structure consistent with the fitting of the EXAFS data, a rutile type IrO₂ with some Ir sites occupied by Ru. All the panels (a-d) show the full structure transparently, and the following atoms are highlighted: (a) the central absorbing Ir (blue) and the 6 surrounding O (red) making up the first oxygen coordination shell; (b) the two nearest Ir/Ru sites; (c) the O atoms making up the second O coordination shell; and (d) the second nearest Ir/Ru sites, in this case also containing Ru (green).
Table 1. EXAFS fitting results from 1 Å and a 2 Å IrO\textsubscript{x} on RuO\textsubscript{2} thin film. The bond distances (R) and coordination numbers (N) are shown for the two dry samples, as well as the \textit{in situ} measurement on the 2 Å IrO\textsubscript{x}/RuO\textsubscript{2} sample at 1.0 V vs. RHE. For comparison the bond distances and coordination numbers are shown for bulk rutile IrO\textsubscript{2}. The first coordination shell of both O and Ir/Ru sites is fully occupied, whereas the second coordination shell has a lower occupancy than bulk IrO\textsubscript{2}, consistent with IrO\textsubscript{x} being at the surface.

<table>
<thead>
<tr>
<th>Sample</th>
<th>First oxygen</th>
<th>First Ir/Ru</th>
<th>Second oxygen</th>
<th>Second Ir/Ru</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>R\textsubscript{1} [Å]</td>
<td>N\textsubscript{1}</td>
<td>R\textsubscript{2} [Å]</td>
<td>N\textsubscript{2}</td>
</tr>
<tr>
<td>1 Å IrO\textsubscript{x}/RuO\textsubscript{2}/GC (Dry)</td>
<td>1.98±0.02</td>
<td>5.5±0.5</td>
<td>3.09±0.09</td>
<td>2.1±2.3</td>
</tr>
<tr>
<td>2 Å IrO\textsubscript{x}/RuO\textsubscript{2}/GC (Dry)</td>
<td>1.98±0.01</td>
<td>5.6±0.3</td>
<td>3.11±0.05</td>
<td>3.1±1.8</td>
</tr>
<tr>
<td>2 Å IrO\textsubscript{x}/RuO\textsubscript{2}/GC (1V)</td>
<td>2.01±0.04</td>
<td>6.1±1.3</td>
<td>3.14±0.12</td>
<td>5.7±9.7</td>
</tr>
<tr>
<td>Bulk IrO\textsubscript{2}</td>
<td>1.96-2.00</td>
<td>6</td>
<td>3.16</td>
<td>2</td>
</tr>
</tbody>
</table>

The IrO\textsubscript{x}/RuO\textsubscript{2} thin films with a nominal deposition of 10 Å were characterized by monochromatized XPS. Figure 2 shows the XPS spectra of Ru 3d (Figure 2a) and Ir 4f (Figure 2b) core level regions. The XPS spectra with variable nominal IrO\textsubscript{x} deposition exhibit similar peak shapes. For both Ru 3d and Ir 4f, the peaks positions attest the presence of the valence state IV. The fitting reveals that the metallic component is zero, only the oxides of Ru and Ir being present at the surface. The XPS spectra exhibit the characteristic features for RuO\textsubscript{2} and IrO\textsubscript{2}, with the pair of doublets for the main peak and the satellite.\textsuperscript{69,70} In all cases, the Ru 3d and Ir 4f peaks could be fitted with fixed positions and similar intensity ratios, suggesting that the chemical phase is not changing among samples with different nominal thicknesses of IrO\textsubscript{x}. Moreover, the peaks do not change for XPS spectra taken at different emission angles, indicating a good homogeneity of the samples.

The depth profile from angle-resolved XPS with the atomic concentration as a function of the depth on the 10 Å IrO\textsubscript{x}/RuO\textsubscript{2} thin films is illustrated in Figure 3. As can be observed, AR-XPS revealed that these surfaces do not present Ru on the topmost layers, confirming that the surface is covered with Ir as deposited. We note that the appearance of a C signal in Figure 3 is presumably due to adventitious carbon accumulated on the surface of the IrO\textsubscript{x}/RuO\textsubscript{2} thin films after air exposure. This is in agreement with the C 1s peak present in Figure 2a. The intensity of the C 1s peak was determined by imposing the 3:2 intensity ratio of the Ru 3d doublets; the position of the C 1s peak resulting from the fit is also compatible with adventitious carbon and it was not found to vary significantly with different angles, indicating a good reliability of the fit.
Figure 2. XPS spectrum of a RuO$_2$ thin film modified with 10 Å IrO$_2$ taken at 21° emission angle: (a) XPS spectrum of Ru 3d core level region; (b) XPS spectrum of Ir 4f core level region.

Figure 3. Depth profile of the surface modified 10 Å IrO$_2$ on RuO$_2$ thin film evaluated using AR-XPS.

The OER activity was measured by quasi stationary cyclic voltammetry (CV) in N$_2$-saturated 0.05 M H$_2$SO$_4$ in the EQCM set-up. Figure 4 shows the CVs showing the OER activity as the geometric current density as a function of the iR-corrected potential versus the reversible hydrogen electrode (RHE) on a 40 nm RuO$_2$ thin film, as well as RuO$_2$ thin films with 1, 2, and 4 Å of IrO$_x$ deposited on
the surface. The inset shows the overpotential required to reach a current density of 5 mA/cm². The bare RuO₂ thin film is most active with an overpotential of 360 mV, and the more IrOₓ on the surface the higher overpotential is measured up to 410 mV for 4 Å IrOₓ/RuO₂. This is in agreement with earlier works on mixtures of Ir and Ru oxides for OER showing a decrease in the activity when Ir is present.⁶,⁴²-⁴⁴ Notably, a significant drop in activity of 40 mV at 5 mA cm⁻² occurs when adding 1 Å of IrO₂ to the RuO₂ thin film. Although the overpotential increases with the amount of deposited IrOₓ, the activity drop becomes less pronounced for 2 and 4 Å of IrO₂/RuO₂. This behavior differs significantly from the linear relationship of the overpotential at 0.1 mA cm⁻² observed by Kötz et al.⁴²

**Figure 4.** Quasi stationary CVs in the EQCM set-up showing the geometric current density as a function of the potential for a 40 nm RuO₂ thin film as well as RuO₂ thin films with 1, 2 and 4 Å of IrOₓ on the surface. The inset shows the overpotential required to reach a current density of 5 mA cm⁻² versus the IrO₂ nominal deposition (in Å). The CVs were recorded in N₂-saturated 0.05 M H₂SO₄ and at 5 mV s⁻¹.

In order to investigate the stability enhancement of the surface-modified RuO₂ thin films, the IrOₓ/RuOₓ thin films were deposited on Au-covered EQCM substrates. We first studied the stability under potentiostatic conditions.⁵³ Figure S6 shows the current density and mass losses of RuO₂ and 2 Å IrOₓ onto RuO₂ thin films after a two-hour chronoamperometry test at 1.8 V vs. RHE. In this case, we could not observe any effect on the stability by adding surface IrOₓ. After one hour, the IrOₓ/RuOₓ thin films behave like pure RuO₂ thin films. We can speculate that the surface Ir atoms did not diffuse to the undercoordinated sites of the RuO₂ surface when the thin films were prepared. Another possible
explanation could be that the corrosion of RuO$_2$ is so fast at 1.8 V vs. RHE that we cannot see any stability effect by addition of sub-monolayer amounts of surface IrO$_x$.

We carried out CVs before the chronoamperometry, to investigate whether this allowed diffusion towards undercoordinated sites. Figure S7 shows little stabilizing effect after an hour test on the chronoamperometric tests after cycling either. However, during the initial potentiodynamic test, we could observe some improvements. This could indicate that potential cycling can help Ir to the undercoordinated sites.

After performing the stability tests under potentiodynamic conditions (Figures S6 and S7), we carried out stability tests under dynamic conditions. We measured the mass loss of IrO$_x$/RuO$_2$ thin films as well as pure RuO$_2$ and IrO$_2$ thin films, for comparison, while the potential was continuously cycled in an accelerated stress test. The potential was cycled from 1.23 to 1.80 V vs. RHE in N$_2$-saturated electrolyte. Figure 5a shows the geometric current density at 1.8 V versus time as a function of the time during the nine-hour test under dynamic operations. In parallel, the EQCM mass losses were recorded as a function of time during that test, as shown in Figure 5b. A positive effect due to the addition of surface IrO$_x$ to RuO$_2$ thin films can be clearly observed now. The total mass decreases continuously for the bare RuO$_2$ thin film due to its poor stability. In contrast, all the measured IrO$_x$/RuO$_2$ films seem to be rather stable. Notably, thin films covered with only 1 Å of IrO$_x$ exhibit a drastic improvement in stability as compared to RuO$_2$. Further addition of IrO$_x$ (2 Å) induce slight additional improvements. The behaviour of 4 Å IrO$_x$/RuO$_2$ thin films over the 9 h stability test is different from the other surface-modified thin films studied in this work, showing a mass gain over time. Pure IrO$_2$ exhibited a similar (although less pronounced) behavior. This may be related to the ability of the EQCM to adapt its properties while the potential is cycled.

![Figure 5](image_url)

**Figure 5.** EQCM measurements while the potential is cycled from 1.23 to 1.80 V vs. RHE at 20 mV s$^{-1}$ in N$_2$-saturated 0.05 M H$_2$SO$_4$ at room temperature for pure RuO$_2$ and IrO$_2$ thin films as well as 1, 2...
and 4 Å of IrO$_x$ deposited on RuO$_2$ thin films during nine hour stability test: (a) Geometric current density at 1.8 V (vs. RHE) as a function of time; (b) mass loss evaluated from the EQCM as a function of time.

We can hypothesize that the RuO$_2$ does not only corrode at the undercoordinated sites. For instance, if the undercoordinated sites were a factor 100 less stable than the terraces, a two-hour stability test at 1.8 V vs. RHE would also induce the corrosion at the terraces. In contrast, if we cycled between 1.23 and 1.8 V vs. RHE for 2 hours, we would only spend 10 min in the potential region above 1.75 V vs. RHE, and 20 min above 1.70 V vs. RHE, which would be equivalent to only 8% and 16% of the total time of the stability test, respectively.

We note that capacitance measurements$^{13,71}$ revealed a much higher double layer capacitance for 40 nm IrO$_2$ thin films than 40 nm RuO$_2$ thin films as well as polycrystalline RuO$_2$ and IrO$_2$ electrodes,$^{66}$ for comparison (see Figure S8 in the S.I.). This suggests that IrO$_2$ thin films exhibits a much rougher surface, in agreement with SEM images (Figure S4), and consistent with the literature.$^{13,67}$ In contrast, RuO$_2$ thin films have a very low roughness, close to polycrystalline RuO$_2$.$^{66}$ We hypothesize that submonolayer amounts of IrO$_x$ on RuO$_2$ thin films are smoother than a pure IrO$_2$ thin film, due to favorable interactions between the film and the surface.

To study and quantify the corrosion rate for RuO$_2$ and IrO$_x$/RuO$_2$ thin films, we combined EQCM measurements with electrolyte analysis by means of ICP-MS. During the stability tests described above, aliquots from the electrolyte were extracted for ICP-MS analysis after 150 cycles. This allowed a quantitative analysis of the amount of Ru and Ir dissolved in this time frame. Figure 6 shows the overpotential at 5 mA cm$^{-2}$ (from Figure 4) together with the mass losses for RuO$_2$ thin films, as well as that for 1 and 2 Å of IrO$_x$ deposited on RuO$_2$ thin films (Figures 6a and b, respectively). As can be observed in Figure 6b, Ru is responsible for the majority of the dissolved material for all samples. IrO$_x$ strongly reduces the amount of dissolved RuO$_2$, the addition of 2 Å IrO$_x$ improves the stability of RuO$_2$ by $\sim$72%. However, we could not observe any further significant change in the stability by addition of 4 Å IrO$_x$, both activity and stability values of these two samples are very similar. Remarkably, the Ir dissolution is extremely low for the IrO$_x$/RuO$_2$ as compared to the Ru dissolution.

Earlier work by Cherevko et al. shows that IrO$_x$ corrosion tends to be enhanced under potentiodynamic conditions, when the electrode is cycled; conversely, the corrosion rate of RuO$_x$ is purely a function of the applied potential, and is insensitive to cycling.$^{35}$ We speculate that cycling the potential may enhance the mobility of Ir; this could either occur via the formation of a solution phase species in a high oxidation state and its subsequent disproportionation, or simply via the surface movement of Ir atoms during oxidation and reduction cycles. The outcome would be that cycling enables the Ir atoms to reside at the undercoordinated sites.

Our results show that stable IrO$_x$/RuO$_x$ systems may be formed with IrO$_x$ exclusively at the surface. This is consistent with the notion that the stability of the oxide material is linked to the dissolution
potential of the surface atoms. At higher potentials, particularly at 1.8 V vs. RHE, the dissolution rate of Ir is significant. Since only a small amount is present, it all dissolves with time. The reason why this is not observed in the potential cycling experiments could be that the fraction of the time spent at a high enough potential for the Ir to dissolve is very small. Consequently, in order for all the Ir to dissolve, the cycling has to be extended for longer periods of time.

**Figure 6.** Activity and stability of RuO$_2$ and IrO$_x$/RuO$_2$ thin films as a function of the nominal deposition of IrO$_x$: (a) Overpotential at 5 mA cm$^{-2}$, from the first quasi stationary CV between 1.23 and 1.8 V vs. RHE at 5 mV s$^{-1}$; (b) ICP-MS mass loss for RuO$_2$ (in grey) and IrO$_2$ (in blue) evaluated after 150 cycles between 1.23 and 1.80 V vs. RHE in N$_2$-saturated 0.05 M H$_2$SO$_4$. The error bars show the standard deviation evaluated from at least four independent measurements.

4. CONCLUSIONS
We have presented the strategy of stabilizing RuO$_2$ by sub-monolayer amounts of surface IrO$_x$. EXAFS measurements revealed the formation of a rutile type IrO$_2$ structure at the surface of the thin films. We have evaluated the stability of the thin films by combination of EQCM and ICP-MS. Sub-monolayers amounts of IrO$_x$ at the surface of RuO$_2$ protect the Ru against dissolution through the higher dissolution potential of Ir. These results can explain earlier reports in the literature, which showed that bulk mixed oxides of Ir and Ru show enhanced stability relative to pure RuO$_x$.$^{36,42-45,72}$ more specifically, we suggest that this stabilization is due to surface phenomena. The addition of IrO$_x$ at the surface increases the overpotential for the OER reaction, but an optimum may be reached in terms of balancing stability and activity. Furthermore, we measured low rates of Ir dissolution at 1.8 V vs. RHE; the protection effect is limited if only a small amount of Ir is present at the surface. This approach is useful to tune stability of active OER catalysts for PEM electrolyzers.

**AUTHOR INFORMATION**

**Corresponding Authors**

María Escudero-Escribano; e-mail: maria.escudero@chem.ku.dk

Ifan E.L. Stephens; e-mail: i.stephens@imperial.ac.uk

Ib Chorkendorff; e-mail: ibchork@fysik.dtu.dk

**ASSOCIATED CONTENT**

Supporting Information available: Additional experimental characterization: X-ray diffraction, scanning electron microscopy, grazing incidence extended X-ray absorption fine structure, electrochemical stability experiments and capacitance measurements.

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Illustration of the oxygen evolution reaction and ruthenium dioxide corrosion on a RuO$_2$ (110) surface (left) and on the same surface with the under-coordinated sites decorated with IrO$_x$ (right).