New Materials for Hydrogen Evolution

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Publication date:
2008

Document Version
Early version, also known as pre-print

Citation (APA):
New Materials for Hydrogen Evolution

Ph.D. Thesis
April 2008

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Preface

This thesis is submitted in partial fulfillment of the requirements for the Ph.D. degree from the Technical University of Denmark. The project has been carried out at the Institute of Physics at the Technical University of Denmark, within the Interdisciplinary Center for Catalysis, ICAT and Center for Individual Nanoparticle Functionality, CINF, during the last three years under the supervision of Professor Ib Chorkendorff. Financial support was provided by the Danish Strategic Research foundation.

I would like to thank Professor Ib Chorkendorff for inspiring discussions, for making it possible to venture into the field of electrochemistry and for having a such close collaboration with the theory group of Jens K. Nørskov. Especially I would like to thank the former and current members of the electrochemistry group, Assistant Professor Thomas F. Jaramillo, Post doc. Billie Abrams and Ph.D student Peter Vesborg for inspiring discussions ranging from electrochemistry and materials science to the next planned social activity.

In connection with the build and many iterations of improving the electrochemistry lab I enjoyed the help and support of Technician John Larsen as well as Hans Christian Sørensen and Dan Schacam.

I also thank Billie Abrams, Klas Andersson and Peter Vesborg who helped to proof read this thesis.

The current and former members of ICAT,CINF,CAMD and CAMP over the years are also thanked for, not only, providing a strong basis for scientific discussion, but also for giving work an important social aspect. Finally I would like thank my beloved wife Josephine and my son Tobias for their loving support.

Jacob Lindner Bonde
København, April 4th, 2008
Abstract

This thesis primarily deals with the identification and electrochemical characterization of materials for the hydrogen evolution reaction, where protons and electrons combine to form hydrogen. Based on computational screening, a BiPt surface alloy was identified as being a potential HER candidate. The BiPt surface alloy was synthesized and identified as being more active for the HER than Pt. A core shell system composed of a Pt skin on an Au core was also investigated and found to have a HER activity that could be increased or decreased depending on the potentials previously applied to the sample. MoS$_2$ nanoparticles were identified as being promising HER catalysts. Further investigations on MoS$_2$ on Au(111) imaged by scanning tunneling microscopy (STM) proved that the active site of these particles was the edge and enabled the quantification of the HER activity. Carbon supported MoS$_2$ and WS$_2$ nanoparticles were synthesized and found to have HER activities close to that of Au(111) supported MoS$_2$. The effect of cobalt promotion of WS$_2$ and MoS$_2$ was also investigated and Co was found to promote the HER activity. Finally Mo$_3$S$_4$ clusters were found to catalyze the HER and electrochemical characterization of STM preimaged samples allowed for the quantification of the turn over frequency per Mo$_3$S$_4$ cluster.
Denne afhandling omhandler primært identifikation og elektrokemisk karakterisering af materialer til brint evolutions reaktionen, hvor protoner og elektroner kombineres til brint.


MoS$_2$ og WS$_2$ på et kul bæremateriale blev også synthetiseret og aktiviteten for brintudviklingsaktiviteten var sammenlignelig med aktiviteten af MoS$_2$ på Au(111). Effekten af kobolt promovering af MoS$_2$ og WS$_2$ blev også undersøgt og tilsætningen af kobolt viste sig at have en promoverende effekt.

Sidst blev det vist Mo$_3$S$_4$ molekuler kunne katalysere brint evolutions reaktionen og elektrokemisk karakterisering af STM afbilledede Mo$_3$S$_4$ partikler gjorde det muligt at kvantifisere aktiviteten pr. molekule.
Included Publications

Paper 1:
“Biomimetic Hydrogen Evolution”
Journal of the American Chemical Society, 127, 5308 (2005)

Paper 2:
“Computational high-throughput screening of electrocatalytic materials for hydrogen evolution” J. Greeley, T. F. Jaramillo, J. Bonde, I. Chorkendorff and J. K. Nørskov,

Paper 3:
“Identification of Active Edge Sites for Electrochemical $H_2$ Evolution from MoS$_2$ Nanocatalysts”
T. F. Jaramillo, K. P. Jorgensen, J. Bonde, J. H. Nielsen, S. Horch and I. Chorkendorff,
Science, 316, 100 (2007)

Paper 4:
“Hydrogen Evolution on Nano-particulate Transition Metal Sulfides”
J. Bonde, P. G. Moses, T. F. Jaramillo, J. K. Nørskov and I. Chorkendorff,
Accepted Faraday Discussions, (2008)

Paper 5:
“Hydrogen Evolution on supported incomplete cubane-type $[Mo_3S_4]^{4+}$ electrocatalyst”
T. F. Jaramillo, J. Bonde, J. Zang, B. L. Ooi, J. Ulstrup and I. Chorkendorff,
## Contents

1 Introduction 1

2 Electrocatalysis 3
   2.1 Electrochemistry and catalysis 3
   2.2 The Hydrogen Evolution Reaction 7
      2.2.1 The Reaction Pathways 7
      2.2.2 Interpreting Tafel Slopes 8
      2.2.3 The exchange current density 8
      2.2.4 The exchange current density and Volcanoes 9
      2.2.5 The HER Catalyst 10

3 Experimental methods and setups 13
   3.1 Electrochemical Methods and the Setups 13
      3.1.1 Voltammetry 13
      3.1.2 The Electrochemical Setups 16
   3.2 X-ray Photoelectron Spectroscopy 17

4 Reducing the use of Pt 19
   4.1 The Descriptor for HER reactivity 19
   4.2 Tuning the Activity of electrocatalysts 20
      4.2.1 Nanoparticles 20
      4.2.2 Overlayers 20
      4.2.3 Alloys 21

5 Pt based systems 23
   5.1 Computational Screening and a BiPt surface alloy 23
      5.1.1 Screening procedure 23
      5.1.2 BiPt the non intuitive candidate 25
      5.1.3 Hydrogen oxidation and CO tolerance 27
   5.2 PtAu core shell nanoparticles 30

6 Biomimetic systems 35
   6.1 HER on Metal Sulfides 37
      6.1.1 Initial Studies on MoS$_2$ 38
6.1.2 Measuring the area of MoS$_2$ .......................... 39
6.2 MoS$_2$ on Au(111) ........................................... 40
6.3 MoS$_2$ on Carbon .............................................. 44
   6.3.1 Preparation methods .......................... 44
   6.3.2 Measuring the area by oxidation ................. 45
   6.3.3 XPS analysis ........................................ 47
6.4 The other metal sulfides ................................. 48
6.5 Promotion of MoS$_2$ and WS$_2$ .......................... 50
6.6 Mo$_3$S$_4$ .................................................. 54
6.7 Hydrogen oxidation and metal sulfides ............... 56

7 Summary and outlook ........................................ 57

Bibliography ............................................... 59
Chapter 1

Introduction

In recent years the focus on renewable energy sources have been significantly increased due to increasing oil prices and the growing concern for global warming. Currently the moment most of the world’s energy consumption is fossil fuel based, but there is a growing interest in replacing the fossil fuels with renewable resources like solar, wind and biomass. Increasing the amount of renewable energy is not a simple task as renewable energy sources like sun and wind do not provide a stable supply of energy, since the energy is produced when the sun shines or when the wind blows.

One way of stabilizing the energy supply is to use different energy carriers that can be produced when there’s a surplus of renewable energy and used when there’s a lack of renewable energy. One suggested energy carrier is hydrogen\cite{1} which can be produced from water by electrolysis and made into power, heat and water in a fuel cell.

The catalysts used in for instance fuel cells are usually based on platinum or platinum ruthenium alloys, these catalysts are both scarce and expensive. Only 200 tons of platinum is mined every year\cite{2} and as current state of the art fuel cells use on the order of 1 gram of platinum pr kW of power\cite{3}, this corresponds to the platinum needed to produce 2 million cars pr. year each equipped with a 100 kW engine.

The aim of this project is to find and characterize alternative materials for the hydrogen evolution reaction(HER) where protons and electrons recombine to form hydrogen.

\[ 2H^+ + 2e^- \rightarrow H_2 \] (1.1)

The hydrogen evolution reaction is one of the most studied electrochemical reactions and is often used as the textbook example\cite{4} on the relation between kinetics and surface properties of electrodes. The HER is not only
interesting from a scientific point of view, due to its commercial importance for use in electrolyzers producing hydrogen.

The outline of this thesis is as follows. Chapter 2 will give a brief overview of the relation between electrochemistry and catalysis, followed by and overview of electrocatalysis with the emphasis electrocatalytic hydrogen evolution. Chapter 3 will be description of the electrochemical methods and the setups used. Chapter 4 is a short description of some the approaches that can lead to a reduction in the use of platinum. Chapter 5 is an overview of the Pt based systems that were investigated during the work on this thesis. Chapter 6 gives an overview of the metal sulfides investigated and the results of the electrochemical investigations of these systems. Finally I have attached the published and submitted papers that where written during this Ph.D study.
Chapter 2

Electrocatalysis

The insight into heterogeneous catalysis has increased tremendously in recent years, due to investigations of well defined model systems combined with theoretical DFT calculations. This has revealed the nature of both rate limiting steps and sites governing the overall activity for many different systems. With these trends it is now possible to make computational screening for new materials with higher activity, selectivity and/or stability and successful screening studies have already been reported [5, 6, 7]. Recently theoretical calculations and detailed studies of model systems have also been used in the field of electrocatalysis both for detailed studies of reaction pathways[8] and for the study of activity descriptors[9, 10, 11, 12].

2.1 Electrochemistry and catalysis

Just as in heterogeneous catalysis, electrocatalysis involves converting reactants into products without loss of the catalyst. But in electrocatalysis it is possible to drive a reaction out of equilibrium by means of the addition or subtraction of charge. In the following some basic equations relevant for the hydrogen evolution reaction will be derived, more detailed discussions about the derivation and the basics of electrode kinetics can be found in reference[4].

The electrode can transfer charge to or from a reactant. One simple example is changing the oxidation state of ions in a solution by means of a metal electrode as follows:

\[ X^{3+}(aq) + e^- (electrode) \rightarrow X^{2+}(aq) \]  

(2.1)

The current for this reaction will then be given by:

\[ i = AFk_0cX^{3+} \]  

(2.2)
Where \( A \) is the area of the electrode, \( F \) is Faraday’s constant \( k_0 \) is the rate of electron transfer and \( c_{X^{3+}} \) is the concentration of the reactant. The current is limited by the concentration of the reactants or the rate of electron transfer. The latter can controlled by applying a potential to the electrode. But in reality an electrochemical reaction can be in equilibrium and the reaction in question can have an activation barrier, just as in heterogeneous catalysis\[13\]. In general the relation between a potential and an activation barrier is given by:

\[
\Delta E = -\frac{\Delta G}{nF}
\]

(2.3)

Where \( \Delta E \) is the electrode potential, \( n \) is the number of electrons transferred and \( \Delta G \) is the the free energy of the activation barrier. As the potential of an electrode is a function of the current density drawn, the therm overpotential is defined as being the difference in electrode potential at equilibrium\( (E_e) \) and the potential at a given current density\( (E_i) \):

\[
\eta = E_i - E_e
\]

(2.4)

Thus the potential of an electrode should be defined by the change in standard free energy \( \Delta G_0 \) and the overpotential \( \eta \) as:

\[
\Delta V = \Delta G_0 \pm \eta
\]

(2.5)

In reality however the relation between potential and free energy is not that simple. There are many other sources that affect the potential such as: ohmic drop, diffusion limitations, stability of the material, etc. An important parameter in electrochemistry is the transfer coefficient determining the amount of charge transferred by means of an overpotential. Thus the change in activation energy pr. electron for a given process as function of overpotential is given by:

\[
\Delta E = -nF\eta
\]

(2.6)

In heterogeneous catalysis the rate of a reaction which is equivalent to the current density can be described by an Arrhenius type equation given by:

\[
k = \nu e^{-\frac{E_a}{RT}}
\]

(2.7)

k is the rate constant, \( \nu \) is the prefactor and \( E_a \) is the activation energy. If we assume that we have a reversible reaction, such as the oxidation and reduction of an Fe ion, we will have a net rate given by:

\[
r = k_c c_O - k_a c_R
\]

(2.8)

Where \( k_c \) is the rate constant for oxidation(anodic) and \( k_a \) is the rate constant for reduction(cathodic), \( c_R \) and \( c_O \) is the concentration of the species
being reduced or oxidized respectively. If we assume that both the oxidation and the reduction reactions are activated by the free energies of $\Delta G_c$ and $\Delta G_a$ respectively. Then the rate constant for the oxidation and the reduction will be given by:

$$k_c = \nu e^{-\Delta G_c/RT}$$

(2.9)

$$k_a = \nu e^{-\Delta G_a/RT}$$

(2.10)

In electrochemistry it is possible to change the activation barrier by adding the overpotential $\eta$ multiplied with the transfer coefficient $\alpha$. Thus the free energy activation barriers will be:

$$\Delta G_c = \Delta G_{c,eq} + \alpha_c n_e F \eta$$

(2.11)

$$\Delta G_a = \Delta G_{a,eq} - \alpha_a n_e F \eta$$

(2.12)

$\Delta G_{c,eq}$ and $\Delta G_{a,eq}$ are assumed to be the potential independent activation barriers at equilibrium and $\alpha_c + \alpha_a = 1$. The rate constants for the oxidation and the reduction will be given by:

$$k_c = \nu e^{-\Delta G_{c,eq} - \alpha_c n_e F \eta/RT}$$

(2.13)

$$k_a = \nu e^{-\Delta G_{a,eq} + \alpha_a n_e F \eta/RT}$$

(2.14)

If we introduce a new constant $k_0$ which includes all the parameters that are assumed to be potential independent and assume that the prefactors and activation barriers for the different rate constant are equal. Then the rate constants are given by:

$$k_c = k_0 e^{-\alpha_c n_e F \eta/RT}$$

(2.15)

$$k_a = k_0 e^{\alpha_a n_e F \eta/RT}$$

(2.16)

The rate constants can now be inserted in equation 2.8 and the net rate equation becomes:

$$r = \frac{i}{n_e F} = k_0(\nu e^{-\alpha_c n_e F \eta/RT} - \nu e^{\alpha_a n_e F \eta/RT})$$

(2.17)

If we convert the rate into current and introduce the exchange current density $i_0 = k_0 n_e F$ we have the classical Butler Volmer equation:

$$i = i_0(\nu e^{-\alpha_c n_e F \eta/RT} - \nu e^{\alpha_a n_e F \eta/RT})$$

(2.18)
Chapter 2 - Electrocatalysis

Figure 2.1: Tafel plot where \( \log|i| \) is plotted as a function of the overpotential. The slope in the linear region is the inverse of the Tafel slope. If the linear region is extrapolated to zero overpotential the exchange current density \( i_0 \) is given by the intercept with the y-axis.

The Butler Volmer equation is often used to analyze the kinetics and thermodynamics of electrochemical reactions by means of Tafel analysis. The Tafel analysis is used when the reaction is out of equilibrium, meaning that either the anodic or the cathodic current is dominating the current density measured. The logarithm of the current density \( \log|i| \) can then be plotted as a function of the overpotential \( \eta \) (See Fig. 2.1). This plot can then be used to evaluate the overpotential \( \eta \) as a function of the current density in the linear region of the Tafel plot given by:

\[
\eta = a + b \log|i|
\]

(2.19)

Where \( a \) is \( \log|i_0| \) and \( b \) is the Tafel slope given by:

\[
b = \frac{2.3RT}{\alpha_c n_e F}
\]

(2.20)

This simple analysis gives two parameters the exchange current density \( i_0 \) and the Tafel slope \( \frac{2.3RT}{\alpha_c n_e F} \). The exchange current density is comparable to the rate constants used in heterogeneous catalysis and just as in heterogeneous catalysis it is governed both by kinetics and thermodynamics. The Tafel slope gives information on both the kinetics and the thermodynamics. It is often used as a fingerprint showing whether or not the individual measurements are dominated by transport limitations. The Tafel slope is governed by the temperature \( T \) and the number of electrons per unit re-
action \( n_e \) and the transfer coefficient \( \alpha \). If we assume that \( \alpha = 0.5 \) and that \( n_e = 1 \) the Tafel slope at 298.15 K is expected to be:

\[
\frac{2 \cdot 2.3RT}{F} = 118 \text{ mV/dec}
\]  

(2.21)

If the measured Tafel slope is higher it is difficult to determine the exact reason for the discrepancy as it could be transport/diffusion limitations, cell resistance or that the proposed mechanism is not not the correct one. Since the Tafel slope parameters \( \alpha \) and \( n_e \) are multiplied it is difficult to determine the exact values of \( \alpha \) and \( n_e \) from the slope alone.

As the Tafel slope and the exchange current density are linked, a wrongly measured Tafel slope can influence the extrapolated exchange currents density. A high Tafel slope will give a high exchange current density and it is thus important to make sure that the Tafel slope is in the expected region.

### 2.2 The Hydrogen Evolution Reaction

The hydrogen evolution reaction (HER) is perhaps the most studied electrochemical reaction and is of importance for applications ranging from electrodeposition and corrosion of metals to energy storage via \( \text{H}_2 \) production.

#### 2.2.1 The Reaction Pathways

The hydrogen evolution reaction is a simple reaction where two protons and two electrons combine to form hydrogen, the reaction is given by:

\[
2H^+ + 2e^- \rightarrow H_2
\]  

(2.22)

Different reaction paths have been proposed, the initial reaction where a proton and an electron react to form an adsorbed hydrogen atom, is called the Volmer reaction\[14\]:

\[
H^+ + e^- \rightarrow H_{ad}
\]  

(2.23)

This reaction step can then be followed by the Tafel reaction\[15\], which is a Langmuir-Hinselwood type reaction where two adsorbed H atoms react associatively and desorbs

\[
2H_{ad} \rightarrow H_2
\]  

(2.24)

or the Heyrovsky reaction\[16\] which is an Eley-Reidel type reaction, where a proton and an electrons react directly with an adsorbed atomic H atom to form Hydrogen.

\[
H_{ad} + H^+ + e^- \rightarrow H_2
\]  

(2.25)
The exact reaction path is not always simple to deduce and it has for instance been found that the same material can exhibit different HER reaction paths on different facets, see example in the following section.

### 2.2.2 Interpreting Tafel Slopes

Studies of single crystal Pt electrodes performed by Markovic et al [17] have shown that the reaction mechanisms can vary for the different crystal planes. They found that the Tafel slope for the HER on the Pt(110) facet was 28mV/dec. $\approx 2.3RT/2F$. From the slope they interpreted the reaction path to be the Tafel-Volmer. On Pt(100) they found Tafel slopes of 37mV/dec. $\approx 2.3RT/3F$ (low overpotentials) and 112mV/dec $\approx 2(2.3RT/F)$ (high overpotentials). From these slopes they interpreted the reaction path to be the Heyrovsky-Volmer. On the Pt(111) facet they did however not find any conclusive evidence of the exact reaction mechanism as they recorded a Tafel slope of 74mV/dec.

The HER in an electrochemical double layer on Pt(111) has recently been investigated by means of DFT calculations by Skulason et al [18], who introduced both a potential and water layers in their calculations. They were able to calculate the activation energies for both the Tafel and the Heyrovsky reactions and they found that they were comparable, with the Heyrovsky reaction being most favourable. The small difference in the calculated activation barriers explain the inconclusive observations on Pt(111) by Markovic et al.

In order to make an exact interpretation of the Tafel slope it is crucial to have the correct value of the transfer coefficient $\alpha$. Usually the transfer coefficient for the HER is assumed to be between 0 and 1 [3]. For Pt the transfer coefficient has been found to be 0.5 according to Gasteiger et al [19]. Skulason et al [18] also calculated the transfer coefficients on Pt(111) for the different HER reactions and reported values of $\approx 0.4$ and $\approx 0.65$ for the Tafel and the Heyrovsky reaction respectively.

### 2.2.3 The exchange current density

The exchange current density $i_0$ for the HER is often the figure of merit when activities of different electrocatalysts are compared to each other. The exchange current density estimate is however not better than the Tafel slope used to extrapolate the current density to zero overpotential. It is thus important to evaluate both the exchange current density and the Tafel slope when different catalysts are compared to each other. For example the exchange current density for Pt ranges over several orders of magnitude depending on the measurements and conditions used. In a recent study...
Gasteiger et al.\cite{19} reported exchange current densities of 235-600 mA/cm² on Pt nanoparticles. This is two orders of magnitude more than the 0.8-1.4 mA/cm² reported by Markovic et al.\cite{17} on Pt single crystals. The two studies were done in very different systems, Markovic et al. measured on single crystals in a rotating disc electrode using a liquid electrolyte (0.05 M H₂SO₄) and Gasteiger et al. measured the activity on a specially prepared membrane electrode assembly. Gasteiger et al. explains the discrepancy as being due to the lack of uncompensated mass transport limitations in their system. Thus care should be taken when the exchange current densities of highly active HER catalysts are evaluated.

2.2.4 The exchange current density and Volcanoes

Volcanoes have been known in heterogeneous catalysis for decades and are based on the Sabatier principle\cite{20}. Stating that optimal catalytic activity is achieved on catalyst with intermediate binding energies (heat of adsorption) of reaction intermediates. In other words if the the binding is too weak the reactant will not adsorb on the surface and if the binding is too strong the the product will not desorb from the surface. Thus an optimal catalyst is able to adsorb the reactant and desorb the product.

![Volcano plot illustrating the Sabatier principle.](image)

Figure 2.2: Volcano plot illustrating the Sabatier principle. The exchange current density is plotted as function of $\Delta G_H$. If $\Delta G_H$ is negative the activity will be governed by the rate of desorption of products and if $\Delta G_H$ is positive the rate of the adsorption of reactant will govern the activity.

The exchange current density for the HER has been studied for decades in order establish a connection between the activity and physical properties of the catalyst. It was first suggested by Parsons\cite{21} that the exchange current density for the HER was governed by the Hydrogen binding free energy $\Delta G_H$. Others have also plotted the exchange current densities as a func-
Figure 2.3: Volcano plot of the measured exchange current densities as a function of the DFT values for $\Delta G_H$. The exchange current densities are compiled from [25, 9, 24] and the calculated values $\Delta G_H$ values are from [11].

Recently this volcano has appeared in a new version where the exchange current density for different metals[9] and Pd overlayers[25] is plotted as function of DFT calculated values of $\Delta G_H$ for the corresponding metals and overlayer surfaces. One could argue that the DFT calculations for such systems are simple and thus not accurate enough to describe the complexities of the electrode surroundings. But the descriptor for the HER activity, the calculated values of $\Delta G_H$, scales with the measured values of the exchange current density, see Figure 2.3.

### 2.2.5 The HER Catalyst

According to Figure 2.3 the most active catalysts for the HER are those of the platinum group metals, the platinum group metals are also special because they are generally stable in corrosive environments[26]. The combination of stability and high activity has led to the use of platinum for various applications in corrosive environments. The catalyst that are currently used in low temperature PEM fuel cells are for instance based on platinum and platinum alloys. Unfortunately the platinum supply is limited and it is thus...
crucial to have alternatives that are just as active and stable. Especially as
the increasing focus on the hydrogen economy could lead to an increased
use of Pt to for instance fuel cells.
Chapter 3

Experimental methods and setups

The experimental methods and setups used in this thesis are a combination of surface science techniques and electrochemical measurements. This combination ensures that we are doing electrochemistry on samples that were well characterized.

3.1 Electrochemical Methods and the Setups

Electrochemical measurements where electrode currents are measured as a function of an applied potential is a strong and versatile tool. These methods can provide information of many different characteristics of a certain electrode setup, such as surface area, reaction kinetics and reaction rates.

3.1.1 Voltammetry

Voltammetry where the potential is ramped with a certain rate while corresponding current is recorded is a very sensitive method. It is possible to measure transfer of charge(current) at the interface between an electrode and an electrolyte. This surface sensitive technique gives the opportunity to study a wide variety of different processes at the electrode interface, but it can at the same time be difficult to identify the process that exactly led to an increase in current. If a constant potential $E$ is applied to an electrode then the steady state current can consist of different components:

\[ i = i_{\text{reaction}1} + i_{\text{corrosion}} + i_{\text{reaction}2} + i_{\text{charge}} \]  

(3.1)

The current can consist of four different current sources, two different reactions, corrosion of the electrode and charging of the electrode, making the exact interpretation difficult. The interpretation is complicated further since the potential applied to the electrode might not be the exact potential.
at the electrode-electrolyte interface. There could also be potential losses in both the electrolyte and the electrode due to resistance. One way to distinguish the different processes from each other is to apply a non constant potential. One simple method is changing the potential linearly from one potential $E_1$ to another potential $E_2$ while the current is recorded, if the potential then afterwards is changed from $E_2$ to $E_1$ (See Figure 3.1) a cyclic voltammogram(CV) can be made.

Figure 3.1: The potential ramp used to create a cyclic voltammogram, where the potential is changed between two potentials with a constant rate.
In Figure 3.2 an example of a CV is shown. This samples CV has specific characteristics. On the anodic (towards more positive potentials) and the cathodic (towards more negative potentials) sweeps there is a set of peaks centered around 0.7 V vs. NHE. This redox feature shows that charge (electrons) are removed from the electrode during the anodic sweep and added on the reverse sweep. Such a feature could be due to anodic desorption of an adsorbed species followed by cathodic adsorption of the same species. It could also be the opposite where anodic adsorption is followed by cathodic desorption. But many other possibilities for such features exist and usually spectroscopic techniques are coupled with the electrochemical setup to measure the exact source of the redox features. The second feature that can be seen in Figure 3.2 is the double layer capacitance which causes a positive shift of the current at anodic sweeps and a negative shift of the current at cathodic sweeps. The size of the redox couple peak and the double layer capacitance is a function of the scan rate as both features consist of addition and removal of a finite charge. A third feature can be seen at potentials more positive than 0.8 vs NHE. Here the current seems to have a positive shift, which could be due to an oxidation reaction of some species or simply corrosion of the electrode. A fourth feature is seen at potentials more negative than zero vs NHE. Here the current is shifted more negative indicating the possible reduction of protons in hydrogen or another reduction reaction. Many other voltammetric techniques exist, where the potential for instance can be applied as sine curve or as steps. Each of these methods are suitable for the investigations of different phenomena. An overview of different techniques and electrochemistry in general can be found in [4].
3.1.2 The Electrochemical Setups

In this thesis we have primarily used two different kinds of electrochemical setups. A setup designed for the test of membrane electrode assemblies (MEA) for fuels and a standard three electrode setup. Both setups were coupled to a Biologic WMP2 potentiostat with an impedance option.

![Diagram of the setup](image)

Figure 3.3: The setup used for the test of electrode materials on high surface area supports in a two electrode setup. The electrodes can be tested for reactivity under a variety of gases.

In Figure 3.3 the fuel cell setup is depicted. This setup was originally designed for isotopic CO exchange measurements[27] and allows for the testing of MEAs under various gases, with the possibility of having the exhaust gases analyzed by a mass spectrometer. This setup was initially used to test electrode materials that were supported on high surface area carbon and made into MEAs. The setup allows for the testing of catalyst for electrochemical reactions like the HER, the hydrogen oxidation reaction (HOR), the HOR in trace amounts of CO and the oxygen reduction reaction (ORR). It has primarily been used for the initial screening of different materials. These materials were tested for stability and ability to catalyze the wanted reactions at reasonable potentials under conditions that resembled those found in commercially available systems. The test of materials in this manner is however quite tedious. This requires careful preparation of the MEAs in order to compare the activities of different catalysts with each other and not the difference in preparation method.
Chapter 3 - Experimental methods and setups

Figure 3.4: Three electrode setup with a SCE reference electrode a counter electrode and a working electrode. The reference electrode is compartmentalized via a salt bridge.

The second type of electrochemical setup used is a typical three electrode electrochemical setup consisting of a working electrode(sample) a reference electrode(SCE) and a counter electrode(usually Pt mesh). The reference electrode was compartmentalized via a salt bridge to avoid contamination issues and the electrolyte was usually nitrogen purged. This setup was used to make electrochemical measurements on a variety of different samples ranging from UHV prepared and characterized nanoparticles on single crystals to sheets of different metals. Typically the measurements were conducted in a 0.5 M $\text{H}_2\text{SO}_4$ electrolyte, and the pH of the electrolyte was always measured after measurements.

### 3.2 X-ray Photoelectron Spectroscopy

X-ray Photoelectron Spectroscopy(XPS) was used to give information of the chemical state and composition of the catalysts before and after electrochemical characterization.

XPS is based on the radiation of a sample with X-Rays. If the X-Rays are absorbed by a core electron, the core electron can escape the sample as a photoelectron with a kinetic energy given by:

$$E_{\text{kin}} = h\nu - E_B - \varphi_{\text{sample}} \tag{3.2}$$
The energy of the incoming photon is given by $h\nu$, the binding energy of the electron in the sample is with respect to the Fermi energy is given by $E_B$ and the work function of the sample is given by $\varphi_{\text{sample}}$. The incoming X-rays are usually created by an Al $\text{k}\alpha$ or a Mg $\text{k}\alpha$ radiation source, with main radiation peaks centered at $h\nu = 1486.6$ eV and $h\nu = 1253.6$ eV, respectively. This means that kinetic energy of the emitted photoelectron can be up to that of the radiation source and thus the inelastic mean free path of the electrons is short, ensuring that this technique is surface sensitive. If the X-ray radiation used is well defined the kinetic energy of the escaping photoelectron will only be a function of $\varphi_{\text{sample}}$ and $E_B$. As each element has specific major and minor peaks in the XPS spectra, XPS can be used to identify the elements by comparison with a reference database\cite{28}. The intensity of the XPS peaks can be used to quantify the concentration of different elements in a given sample, and often XPS is used to check samples for unwanted impurities. The position of the peaks of the different elements can also be shifted by the local chemical environment of the elements. Finally XPS is a surface sensitive technique since the mean free path of the escaping photoelectrons is on the order of nanometers, thus only photoelectrons from the outmost layers of the sample will escape and be detected.
Chapter 4

Reducing the use of Pt

Many approaches can be used in the search for alternative catalyst and historically a trial and error approach has been successful. Haber for instance tested several thousands of catalyst in several thousands of experiments before he found the optimal catalyst for the Haber-Bosch process\cite{Haber}. This iron based catalyst is still in use today. In recent years high throughput combinatorial methods have been employed to find and improve catalysts\cite{combinatorial}. These methods of trial and error have been and are still successful but it requires significant resources to synthesize and test many different materials in this manner. By means of insight into the different descriptors that govern the activity and the stability, it should however be possible to avoid the major parts of the trial and error test. Since the descriptors can be used to exclude a majority of candidate materials. Ultimately the goal is to be able to design, synthesize, characterize and test materials for given reaction without having to due several iterations of each step.

For the hydrogen evolution reaction it is not only important to find active materials as stability is an issue especially if the catalyst is to be used in acidic environments.

4.1 The Descriptor for HER reactivity

For the hydrogen evolution reaction, the descriptor for the activity is the free energy of hydrogen adsorption $\Delta G_H$ \cite{descriptor} and it is possible to calculate the value of this descriptor by means of DFT\cite{DFT}, as described in section 2.2.4. Knowledge about the nature of the descriptor allows one to improve the a property of an already existing catalyst or to find alternative materials with a similar value of this descriptor. In the following section the focus will be on some of the different approaches that have been considered to improve HER activity of Pt and Pd.
Chapter 4 - Reducing the use of Pt

4.2 Tuning the Activity of electrocatalysts

The activity of a catalyst is usually determined by the turn over frequency per site, however in reality most of the concern is related to the activity on a per gram basis. It is well known that Pt and Pd which are members of the platinum group metals are good HER catalysts having values of \( \Delta G_H \) close to zero.

4.2.1 Nanoparticles

The use of a catalyst in the form of nanoparticles has been known for years. One of the major improvements in for instance the fuel cell industry is the use of smaller particles which give a higher surface areas per gram of catalyst. The use of nanoparticles gives an advantage in the form of surface area per gram of catalyst but in recent years it has been found that some nanoparticles are exhibiting improved activity that cannot only be explained by the high surface area. One example of this is room temperature CO oxidation on Gold nanoparticles that was discovered by Haruta et al\[31\]. This work has been followed up by several groups and in a recent study Janssens et al\[32, 33\] showed that the activity indeed was correlated with the number of the under coordinated corner sites of the Au nanoparticles. The opposite effect has been reported for the ORR on Pt nanoparticles where small particles appear to be less active on a per surface area basis than small particles\[34, 35\]. Another effect of nanoparticles that should be taken into account is wetting as some nanoparticles change shape as function of the surrounding conditions. An example of this this effect has recently been observed on Cu nanoparticles\[36\]. It is thus important to have knowledge about the number of the most active site of the nanoparticles under operating conditions.

4.2.2 Overlayers

Another way to tune the activity of a catalyst is to stress or strain the lattice. Stress or strain will according to d-band model\[37\], change the position of the d-band center and thus the binding energy of an adsorbed molecule\[38\]. Strain will cause the d-band center to be lowered with respect to the Fermi level and thus the the binding energy of adsorbates increase. The opposite is the case for the introduction of stress. This effect has been studied by Greeley et al\[25\], who studied the HER of Pd overlayers on different substrates. They found as predicted by the d-band model that it was possible to change \( \Delta G_H \) and thus the HER activity by the introduction of stress or strain.
4.2.3 Alloys

The alloying of Pt with other, possibly cheaper metals is well studied for the Oxygen reduction reaction (ORR). Pt has been alloyed with almost every transition metal and tested for ORR activity. One of the early studies where performed by Jalan et al. [39] who studied the ORR in phosphoric acid fuel cells. They found a relation between the activity and the lattice spacing of Pt, and proposed that alloys of Co or Ni and Pt should be the next promising candidates. Both of these catalysts where in fact later used in PEM fuel cell and both PtCo and PtNi have been investigated both theoretically and experimentally [40, 41]. In the case of PtCo it was, however, later found that the Co corroded away and the surface could basically be considered as a Pt skin on a PtCo alloy [42] where the lattice is slightly stressed. This should cause the oxygen binding energy to be reduced, which according to the model proposed by Nørskov et al. [8] should increase the activity.
Chapter 4 - Reducing the use of Pt
Pt based systems

Pt based systems have been extensively studied for the HER and the chance of finding an alternative Pt based system that has not previously been studied is limited. In the following chapter measurements will be presented for two different Pt based systems. The first system is a BiPt surface alloy that was identified by means of high throughput combinatorial screening. The second system is a core shell system with a core of Au and a shell of Pt.

5.1 Computational Screening and a BiPt surface alloy

The value of $\Delta G_H$ for the HER can be calculated by means of DFT. In recent years the development of computers and DFT codes have made it possible to perform large scale computational screening of different properties for many different materials. In this section the DFT screening procedures and the non intuitive result of the screening will be presented that were used in Paper 2.

5.1.1 Screening procedure

Several approaches can be undertaken when a screening procedure for alternative catalysts is employed. For the HER the main activity parameter to screen for is $\Delta G_H$. In this case $\Delta G_H$ was calculated for mono and sub-monolayers of a solute element on a host material.
Chapter 5 - Pt based systems

Figure 5.1: The calculated values for the $\Delta G_H$ of the combination of 16 metals and semimetals

The initial calculations on the 736 different systems gave approx. 105 systems with a value of $\Delta G_H$ close zero. This is illustrated as bright yellow dots in Figure 5.1. In order to narrow the amount of viable candidates further down several stability calculations were made. The stability calculations where as follows: 1: The free energy change for surface segregation, i.e. segregation of surface atoms into the bulk. 2: Free energy of infrastructure transformations, i.e. islanding and deformation. 3: Free energy of oxygen adsorption by water splitting, i.e. surface poisoning and oxidation. 4: Free energy of dissolution of the materials from the electrochemical series, i.e. corrosion effects.
Chapter 5 - Pt based systems

Figure 5.2: Pareto-optimal plot of stability and activity of surface alloys for the Hydrogen Evolution Reaction (HER). The Pareto-optimal line indicates the best possible compromise between activity and stability. The alloys in the lower left quadrant are the most promising candidates.

In Figure 5.2, the free energy of stability is plotted as a function of the absolute value of $\Delta G_{H^2}$. The Free energy of stability is the most pessimistic values from the stability calculations performed[11, 43]. The Pareto optimal candidates, the candidates which are the best comprise between stability and activity, are shown by a Pareto line. The Pareto optimal candidates are, 1/3 ML Bi on at Pt Host(i), 1 ML of Rh on a Re host, 1/3 ML Cd on Re, and 2/3 ML Ir on Fe. Of these candidates we found that 1/3 of a monolayer of Bi on top of Pt was an interesting candidate.

5.1.2 BiPt the non intuitive candidate

A Bi surface alloy on top of a Pt host is an especially intriguing candidate as it is rather counterintuitive to mix an active metal Pt and an inactive semi metal Bi. The respective current densities for Pt and Bi are illustrated on Figure 2.3. Electrocatalysts based on Pt and Bi have been extensively studied for various oxidation and reduction reactions and an overview of these studies has been given by Markovic and Ross[44]. No electrochemical studies of surface alloys of Pt and Bi were mentioned. Feliu and Aldaz did however study a similar system[45], Bi ad-atoms on Pt(111) and found that the adsorption of Bi ad-atoms poisoned the HER activity of Pt(111). Despite the previous findings on the BiPt systems it was decided to synthesize a BiPt surface alloy.
In a surface science study by Campbell et al. [46] it had previously been reported that Bi on Pt(111) single crystals formed a surface alloy when heated to 370 °C. Thus it should be possible to create a surface alloy by means of heating Bi ad-atoms on a Pt host. The first step in the preparation of the surface alloy is to find a clean Pt host material, as the high temperatures will cause any contaminants to segregate to the surface. Initially it was chosen to use electrodeposited Pt on a fluoride doped tin oxide coated glass (FTO) substrate. The as deposited Pt was calcined at 500 °C and then tested for the HER, Bi was then adsorbed by means of underpotential deposition of Bi from a 1mM Bi\textsuperscript{3+} aqueous solution. The sample then was tested for the HER again before the sample finally was heated to 500 °C. This led the creation of the BiPt surface alloy skin, and the sample was again tested for the HER.
Figure 5.3 shows the result of the test for HER activity at different stages of the synthesis on the same sample. Initially, the pure Pt sample has the expected activity for a polycrystalline Pt sample. After Bi is irreversibly adsorbed the activity is decreased just as Feliu and Aldaz had previously observed\[45\]. This sample was then heated and the activity was found to increase again and that to a point where the BiPt surface alloy was in fact more active than the initial Pt sample.

This was exactly what the DFT screening predicted. In order to make sure that this increase in activity was not due to a roughening of the surface the H\textsubscript{UPD} region was recorded, see Figure 5.4. As it can be seen the H\textsubscript{UPD} region is not increased after the different synthesis steps, this indicates that the increase in activity is not due to a simple change in surface area of the sample. A more detailed discussion of this study can be read in Paper 2\[11\].

5.1.3 Hydrogen oxidation and CO tolerance

The BiPt was also tested for the HER and the HOR with or without 100 ppm CO. We opted to test the BiPt surface alloy for CO tolerance because the hydrogen binding energy was lower on BiPt than on pure Pt. As binding energies are usually correlated\[47\], a lower hydrogen binding should give a lower CO binding energy, which should correspond to a lower CO coverage.
and thus a higher level of CO tolerance[27]. In order to test the CO tolerance
the BiPt surface alloy was deposited on a high surface area support and made
into a membrane electrode assembly.

Figure 5.5: The preparation of BiPt on ETEK and the test procedure. A
slightly modified version of the preparation procedure from[11] was used,
and as the nanoparticles had a tendency to sinter when heated the same
sample could not be tested at different stages of the synthesis.

ETEK Pt was used as the support material for the surface alloy and the
Bi$_{UPD}$ method was used to absorb Bi on the Pt nanoparticles as in[35, 45].
The heating procedure was however slightly changed as the sample was
heated to 400 °C for 12 minutes to avoid sintering of the Pt nanoparticle
and oxidation of the carbon support, See Figure 5.5. The activity of the
BiPt surface alloy on Pt nanoparticles could not be evaluated at different
stages of the synthesis as they had to be pressed into an MEA in order to be
tested. Thus the alloyed sample was always tested with respect to a sample
that had seen high anodic potentials where Bi would be oxidized and most
likely desorbed into solution[26]. It was found by means of chronoampero-
metric studies that the BiPt surface alloy was not stable at potentials more
anodic than 0.25 V vs. NHE.
Chapter 5 - Pt based systems

Figure 5.6: Left: HER on ETEK Pt nanoparticles with irreversibly adsorbed Bi. After oxidation at 1.2 V vs. NHE the activity increases, corresponding to the removal of the Bi-adatoms that poisoned the HER. Right: HER on ETEK Pt nanoparticles with a surface alloy of BiPt, after oxidation at 1.2 V vs. NHE the HER activity is decreased corresponding to the removal of the more active BiPt surface alloy. Measurements were performed in the Fuel cell setup with a flow hydrogen (20 ml/min.) over both electrodes and with a scanrate of 5 mV/s.

In Figure 5.6 the HER on two different samples are measured before and after the removal of Bi by means of oxidation at 1.2 V vs. NHE. It was found that the HER activity increased after oxidation of the sample with adsorbed Bi ad-atoms and decreased after the oxidation of the sample with the BiPt surface alloy. This finding supports the previous findings[11], where it was found that Bi ad-atoms poisoned the HER on Pt whereas a BiPt surface alloy promoted the HER on Pt.
Figure 5.7 shows an initial attempt to study the CO tolerance of the BiPt surface alloy. It can be seen that the BiPt surface alloy is slightly more tolerant to CO than the the same sample that have been subjected to a potential of 0.9 V vs. NHE (where Bi is assumed to be desorbed). From a scientific point of view this effect is interesting, but for commercial purposes the CO tolerance of the BiPt surface alloy is not significant, and as the preparation procedure was rather tedious this system was not investigated further.

5.2 PtAu core shell nanoparticles

Another Pt based system that was investigated was a core shell, composed of a Pt skin on an Au core. Such core shell systems could potentially reduce the amount of used Pt significantly and at the same time this system would not suffer from corrosion issues as both Pt and Au are quite resistant to corrosion[26]. This system will not suffer the corrosion issues found on both Pt alloys like PtNi and PtCo[42]. Is has also recently even been proposed that the addition of Au clusters to Pt will reduce corrosion [48]. Pt overlayers on Au(111) have previously been studied by Besenbacher et al [49], they find that the the reactivity of Pt is increased whit respect to pure Pt. They
explain this finding with the d-band model[37] where stress in the lattice of
Pt induces an shift in the position if the d-band and a subsequent change in
the reactivity of the surface. For the case of Pt an increase in reactivity is
optimal since Pt according to models for both the HER[9] and the ORR[8]
is to reactive. But if the amount of Pt used in a core shell system of Pt and
Au could be reduced the significantly the lower amount of Pt used would
counteract the loss in activity and the system might then be viable anyhow.
Since the system were to be tested in in the fuel cell setup described in section
3.1.2 ETEK Au nanoparticles that were supported an xc72 carbon were
used as the host material. The Au nanoparticles were mixed with Nafion
to form an ink and supported on a Toray paper. Pt was then deposited
on the Au nanoparticles in a three electrode cell(0.5 M $H_2SO_4$), where a
Pt mesh counter electrode was used as a Pt source. Consecutive cycling of
the potential of the Au working electrode from -0.2 to +1.6 V vs. NHE led
to contamination of the Au nanoparticles with Pt and thus the creation of
nanoparticles with a core of Au and a shell of Pt. This electrode was then
pressed into a MEA together with a commercially available Pt gas diffusion
electrode(serving as both the counter and the reference electrode) from IRD
fuel cells and mounted in the Fuel cell setup.

![Figure 5.8: Chronoamperommetry of the hydrogen oxidation reaction at 0.2 V vs. NHE on the PtAu core shell particles in the fuel cell setup, note how the HOR activity decreases significantly as a function of time, a hydrogen flow of 20 mL/min. was applied to both electrodes.](image)

Initially the electrode was tested for the HOR reaction in a chronoamper-
The HOR activity is initially high, but in time the activity is drastically reduced down to a level where hardly any hydrogen oxidation is occurring. The exact cause of this reduction in activity could be due to several effects like corrosion, poisoning and segregation effects. As the applied potentials are reasonably low and the gases used were pure(N60) the loss in activity is most likely not due to poisoning or corrosion. Thus the cause of the loss in activity could be a surface enrichment of Au, this notion is further supported by theoretical calculations[49, 50] where it was found that it was thermodynamically favorable to have Au on top of Pt.

Figure 5.9: Cyclic voltammogram preformed after the deactivation of at 0.2 V vs. NHE. The initial potential was 0 V followed by a sweep to -0.3 V. Then the potential was cycled between -0.3 V to 1.6 V vs. NHE. Note how the HOR initially is quenched until the potential has been cycled anodically to 1.6 V vs. NHE and an abrupt increase of the HOR can be noticed on the cathodic back sweep. The scanrate was 5 mV/s and the hydrogen flow was 20 ml/min on both electrodes.

Figure 5.9 shows CVs of the sample that deactivated during HOR in figure 5.8. On the first sweep towards the anodic potential of 1.6 V vs. NHE there is hardly any evidence of hydrogen oxidation, as it would be expected if we had a pure Au surface. But on the subsequent cathodic sweep an abrupt increase in the HOR is seen at around 0.9 V vs. NHE and the sample is suddenly behaving more like a Pt sample. This change from the sample being Au like to the sample being Pt like was intriguing and the same sample was then deactivated during HOR at 0.2 V vs. NHE and the change in activity
was investigated by means of CVs that were swept to more and more anodic potentials.

Figure 5.10: Detailed investigation of the onset of the activation of the PtAu system. Again a deactivated system was investigated, but this time several CVs were recorded. In the initial CV the potential was swept between -0.1 V and +0.1 V vs. NHE. The following CVs were swept to more and more anodic potentials, until the sample started to exhibit and increase in the HOR current. The scan rate was 5 mV/s and the hydrogen flow was 20 ml/min on both electrodes.

A more thorough investigation of the onset of the sample becoming Pt-like is shown on Figure 5.10 where the sample gradually is subjected to a more and more anodic potential. The sample behaves Au like and shows no or low HOR activity until the potential has reached 1.0 V vs. NHE where consecutive sweeps show an increase in the HOR activity. We have thus seen how the core shell sample becomes Au like a potentials between -0.2 and 0.2 V vs. NHE. While it becomes more Pt like at potentials above 0.9 V vs. NHE. The switch from being Au like to being Pt like is apparently occurring with different reaction mechanisms a fast mechanism that only required one sweep to 1.6 V vs. NHE and a slow mechanism requiring several sweeps to 1.0 V vs. NHE.

In parallel with this study, a study was also conducted in the three electrode setup, where Au foil or particles are being coated with a thin layer of Pt and this sample shows the exact same characteristic switches from the sample being Pt like to sample being Au like. This sample is currently being
investigated by means of XPS, when the sample behaves like Pt or like Au. Currently the following conclusion have been made on the mechanisms in question and the state of the surface at different potentials:

- At moderate anodic potentials up to 1.0 V vs. NHE Au seggregates to the surface covering the Pt skin.

- At 1.6 V vs. NHE Au is probably corroded\[26\] away from the surface exposing the Pt in one single sweep. Notice that Pt will form a stabilizing oxide at these potentials[26]

- At 1.9 vs. NHE Pt is apparently slowly segregating towards the surface, the driving force could for instance be OH adsorbates that adsorb stronger on Pt than on Au, thus making it more thermodynamically favourable for Pt to be at the surface.

The system is currently under further investigation[51]. The PtAu core shell system is an example of the complexity involved in the design of electrocatalysts in general, as the designed system could be a different system under operating conditions.
Chapter 6

Biomimetic systems

Nature has its own catalysts in the form of enzymes, the common denominator for these enzymes is that they in general are based on abundant materials. Some of these enzymes are indeed able to evolve hydrogen\[^52\] and the enzymes hydrogenases and nitrogenases are examples of these enzymes based on building blocks like Fe, Ni, Mo, S and N. The hydrogenase enzyme is primarily evolving hydrogen\[^53\], whereas the nitrogenases evolves hydrogen as a part of the fixation of nitrogen to form ammonia\[^54\]. Both of these have been found to evolve hydrogen effectively\[^55\] and work has also been done on electrodes where the enzymes\[^57\] or structures resembling these enzymes\[^58\], \[^56\], \[^59\] were tested for the HER.

![Figure 6.1: The active cite of nitrogenase and hydrogenase and the calculated values for $\Delta G_H$. Note how the hydrogen binding energy on the Pt and the enzymes are similar.](image)

The structure of the active site of both nitrogenase and hydrogenase has also been modelled\[^60\], \[^61\]. This allowed for the calculation of the descriptor for hydrogen evolution, $\Delta G_H$. Figure 6.1 shows the structure of the active site of hydrogenases and nitrogenases, and the corresponding values of $\Delta G_H$. 

35
Chapter 6 - Biomimetic systems

The calculated values of $\Delta G_H$ are close to zero and thus these enzymes should be active for the HER. This finding was indeed promising since the DFT calculated values of $\Delta G_H$ apparently also where a useful descriptor for HER activity of sulfides. In this chapter the HER activity of different metal sulfides will be discussed, the majority of the discussion is based on the work presented in Paper 1 and Paper 3-5.
6.1 HER on Metal Sulfides

The active site on both the nitrogenase and the hydrogenase contains both metal and sulfur atoms, and for the case of the nitrogenase the hydrogen is bound to an under coordinated sulfur atom. Such undercoordinated sites are also found on several types of inorganic metal sulfides, one of the most studied metal sulfide systems is nanoparticulate MoS\textsubscript{2}. MoS\textsubscript{2} based catalysts have been extensively studied, since they are used commercially for the hydrodesulfurization (HDS) process\cite{62}. Recent progress in both STM imaging\cite{63} and modelling\cite{64} of MoS\textsubscript{2} nanoparticles have given important insights into the morphology of supported MoS\textsubscript{2}. Bulk MoS\textsubscript{2} consists of stacked S-Mo-S layers, whereas the MoS\textsubscript{2} nanoparticles are found as single S-Mo-S layer hexagonal structures exposing two different kinds of edges, the so-called Mo-edge and the S-edge. It has been shown that the structure of nanoparticulate MoS\textsubscript{2} is a single layered truncated triangle primarily exposing the Mo-edge when supported on Au(111)\cite{63}.

![Figure 6.2: The Mo-edge of MoS\textsubscript{2} nanoparticle with different hydrogen coverages on under coordinated sulfur atoms. Note how the active site resembles the active site on the nitrogenase enzyme. The calculated value of $\Delta G_H$ on the Mo-edge of MoS\textsubscript{2} is found to be close to zero and it is thus expected that MoS\textsubscript{2} will evolve hydrogen at reasonable overpotentials.](image)

Figure 6.2 shows the active site for the HER being the Mo-edge of a MoS\textsubscript{2} nanoparticle and the DFT calculated value of $\Delta G_H$ on the Mo-edge. This finding was indeed interesting since MoS\textsubscript{2} in its bulk form has been subject to extensive photo-electrochemical studies\cite{65, 66}, and it was the general notion that MoS\textsubscript{2} was not a viable HER catalyst\cite{66}. There were however a few studies of the HER on MoS\textsubscript{2}. One study was reported by Nidola et
al[67] who found that the HER activity and the stability of a NiS coating increased when MoS$_2$ was added. The increased activity was described as being due an increase in "true surface area". Another HER study on MoS$_2$ was reported by Sobczynski[68] who studied MoS$_2$ supported on SiO$_2$ and found that hydrogen was evolved from an acidic V$^{2+}$ solution when MoS$_2$ supported on SiO$_2$ was added. The amount of hydrogen was found to be comparable with the amount of hydrogen evolved when SiO$_2$ supported Pt was added. It was concluded that MoS$_2$ and Pt had similar hydrogen evolution activity. None of these studies tried to measure an overpotential or to normalize their currents as a function of for instance surface area or number of active sites.

6.1.1 Initial Studies on MoS$_2$

To test whether or not MoS$_2$ is a viable HER catalyst, MoS$_2$ on a high surface area Dahope carbon was tested in a MEA. Where Pt was used as both the counter and the reference electrode and the MoS$_2$ was used as a working electrode. See Paper 1 for more details[10]. The MEA was then tested in the fuel cell setup and a mass spectrometer was used to verify that hydrogen indeed was evolved.

![Polarization curve of the carbon supported MoS$_2$, compared with a Polarization curve of the carbon support and Pt. The scanrate was 5 mV/s and 20mL/min. of hydrogen was passed over both electrodes.](image)

Figure 6.3: Polarization curve of the carbon supported MoS$_2$, compared with a Polarization curve of the carbon support and Pt. The scanrate was 5 mV/s and 20mL/min. of hydrogen was passed over both electrodes.

The result of the initial test is shown in Figure 6.3. The MoS$_2$ sample is able to evolve hydrogen at around -0.15 V vs NHE, at a rate which is significantly higher than the carbon support. It is also clear that MoS$_2$ is not as active
as pure Pt, this was however expected as the value of $\Delta G_{H}$ for platinum is closer to zero than for MoS$_2$. We could, based on these measurements not quantify the specific activity or for the HER on MoS$_2$ in a manner that allowed for direct comparison of the activity with other HER catalysts. We were also not able to show whether the edge of the MoS$_2$ nanoparticles indeed was the active site.

### 6.1.2 Measuring the area of MoS$_2$

The quantification of the number and type of different sites on a catalyst is one of the most important and difficult tasks when the activity of a certain catalyst is evaluated. MoS$_2$ has as mentioned before been studied for its ability to catalyze the HDS reaction, and only in recent years the role of the different cites on this catalyst has been understood\[69\].

Given that the amount of research on especially MoS$_2$ is comprehensive it should be a simple task to find a simple and suitable method that would determine the number and types of sites of the MoS$_2$ sample prior to electrochemical characterization. Several studies have been made were the HDS activity was correlated with the adsorption of different simple molecules like CO, NO and H$_2$\[70, 71, 72\]. We did some preliminary work on hydrogen deuterium exchange, where we tried to desorb preadsorped deuterium with hydrogen and thus use the amount of desorbed deuterium as a measure of the number of active sites. We had previously used the same technique in the same fuel cell setup to measure CO exchange reactions on Pt and PtRu alloys\[27\]. The MEA required humidification and thus the electrode was in an aqueous environment. The aqueous environment caused a continuous hydrogen/deuterium exchange between hydrogen from water molecules and the preadsorbed deuterium at a rate that was not zero. Thus preadsorbed deuterium would be exchanged with hydrogen from water and we were not able to use the amount of desorbed deuterium as measure of the number of active sites. We did also try to characterize the carbon supported particles with a transmission electron microscope(TEM) and realized that the contrast of the single layers of S-Mo-S was to low to be easily distinguished from the carbon support. Recent progress in high-angle annular dark-field scanning electron microscopy(HAADF-STEM) has however increased the contrast to a degree where it is possible to image the structure of single layers of MoS$_2$ on carbon\[73\]. The structure of the carbon supported MoS$_2$ had the same truncated triangular structure as previously found on Au(111) supported MoS$_2$\[63\]. But even with the ability to characterize the sample used in the fuel cell setup, the complexity of the MEA would make it hard to compare the activity of a series of samples with each other. Even small changes in for instance transport properties or MEA pressing temperature would cause change in activity that might be bigger than the changes introduced to the sample.
6.2 MoS$_2$ on Au(111)

MoS$_2$ on Au(111) has been studied extensively by Besenbacher et al.\[63, 64, 74, 75, 69\] and it has been shown that the particles have the characteristic structure of a truncated hexagon on Au(111), graphite\[73\] and HOPG\[76\]. Thus the structure of MoS$_2$ on Au(111) should be similar to that on the high surface area graphite that was recently investigated for the HER\[10\].

![STM images of MoS$_2$ on Au(111)](image)

Figure 6.4: STM images of MoS$_2$ on Au(111), left, middle: 470 Å by 470 Å of MoS$_2$ with different coverage and particle size. Right: 60 Å by 60 Å image of a single MoS$_2$ particle showing the truncated triangle with the Mo-edge being the primarily exposed edge.

Electrochemical measurements on well defined surfaces has provided important information about surfaces and surface structures like facets\[77\], steps\[78\], stress and strain effects\[79\]. The combination of STM characterized samples and electrochemical measurements is a strong tool, providing information about the relation between structure and electrochemical behavior\[80, 81, 82\]. In order to gain further insight into the relation between structure and reactivity of the MoS$_2$ nanoparticles a series of samples of MoS$_2$ on Au(111) was prepared and characterized by the STM, see Figure 6.4. The samples were deliberately synthesized with different coverages and particle sizes such that effects on the activity stemming from coverage and particles sizes could be distinguished. The UHV prepared and characterized samples were after preparation tested in a three electrode cell especially designed to measure on single crystals.
Figure 6.5: Polarization curves and tafel plots of the MoS$_2$ on Au(111) samples with different particle sizes and coverages, the measurements were done in 1 M $H_2SO_4$ with scanrate of 5 mV/s.

Figure 6.5 shows how the electrochemical measurements are able to distinguish between the HER activity for differently prepared samples. We also find that the Tafel slope consistently is around 57 mV/dec. for the MoS$_2$ samples. This allows us to compare the exchange currents densities of the different samples with each other.
The exchange current densities were then plotted as a function of two different STM recorded sample characteristics, the surface coverage and the edge length of the particles, see Figure 6.6. The exchange current density is only a function of the edge length and not the coverage of the nano particulate MoS$_2$. It was thus concluded that the edge of MoS$_2$ as predicted by Hinnemann et al.[10] indeed is the active site.

With the knowledge of the active site and the number of active sites it was possible to estimate the exchange current density pr. cm$^2$ of MoS$_2$ edge.
This allowed for a direct comparison of the HER activity of MoS$_2$ with that of other HER catalysts, see Figure 6.7. Note that the activity of MoS$_2$ is found to be just below that of the Pt family catalysts. For more details see Paper 3[83].
6.3 **MoS$_2$ on Carbon**

In parallel with the study of MoS$_2$ on Au(111), studies on carbon supported MoS$_2$ were also conducted, since the carbon supported MoS$_2$ is more commercially relevant. The carbon supported systems are as mentioned in section 6.1.1 more inhomogeneous and difficult to image. Since knowledge about the concentration of active sites is paramount in the elucidation of structure-activity relationship, we aimed to use irreversible oxidation features of MoS$_2$ as the probe for number and types of active sites.

### 6.3.1 Preparation methods

The first step in such a study is the choice of host materials since a wrong choice could give rise to complications like contamination and transport limitations. Initially the MoS$_2$ was prepared on high surface area carbon powder, by means of wet impregnation of an aqueous solution ammonia-heptamolybdate. The sample was then calcined in air and sulfided at 450 °C for 4 hours in a flow of 10 % H$_2$S in Hydrogen. The MoS$_2$ on the high surface area carbon was then bonded to an inert electrode with Nafion. Unfortunately severe transport limitations were experienced, when the sample was tested in 0.5 M H$_2$SO$_4$ in the three electrode setup. It was impossible to use our results as the Tafel slope literally were on the order of volts. We then decided to try a more planar support like Au(111), but since the sample would be heated to 450 °C the requirements to the cleanliness of the support material was high. One such clean material is highly orientated pyrolitic graphite which is known to be clean. Unfortunatelly HOPG is very inert and special treatment is required to bond MoS$_2$ to the surface[76, 10] and was thus rejected. We finally found that Toray carbon paper consisting of high temperature treated carbon fibres had the desired cleanliness and at the same time had enough defects to bond MoS$_2$. 
Figure 6.8: Tafel plot (main) and polarization curve (inset) in the cathodic potential range of MoS$_2$ supported on Toray paper. The Tafel slope in the HER region is found to be 120 mV/dec. this is in the expected range for the HER. The scanrate is 5 mV/s and the measurement was conducted in 0.5 M H$_2$SO$_4$.

Figure 6.8 shows a typical polarization curve and Tafel plot of MoS$_2$ supported on Toray paper, firstly it is noticed that hydrogen primarily is evolved at potentials lower than -0.15 V just as on the other MoS$_2$ systems and we also find that the Tafel slope is around 120 mV/dec. This value of the Tafel slope is in the expected region for the hydrogen evolution reaction, see section 2.2.2.

6.3.2 Measuring the area by oxidation

Early in our studies, we found that MoS$_2$ was irreversibly oxidized at anodic potentials higher than 0.6 vs. NHE, and we did several investigations of whether this feature could be used to determine or estimate the surface area or number of active sites on the carbon supported MoS$_2$ nanoparticles. Figure 6.9 shows the oxidation of MoS$_2$ and the subsequent decrease in activity. The oxidation peak appears to have two different components a major component at approx. 0.98 V vs. NHE and a minor component at 0.7 V vs. NHE. This indicates that some parts of the MoS$_2$ is oxidized more easy than other parts. As it is known from literature that bulk MoS$_2$ is preferably oxidized at steps/edges[84], we interpret the two peaks as oxidation of the edge(minor) or the basal plane(major) of the MoS$_2$ nanoarticles.

A back of the envelope calculation on triangles show that the relation be-
Figure 6.9: Cyclic voltammogram of the oxidation and subsequent deactivation of the MoS\(_2\) sample in 0.5 M H\(_2\)SO\(_4\). Scanrate 2mV/s. Main: The deactivation of the sample showing one sweep from -0.35 V vs. NHE to 1.05 V vs. NHE and back to -0.35 V vs. NHE. On the 1st anodic sweep an irreversible oxidation peak occurs at 0.6 V vs. NHE and is followed by a subsequent decrease in current at cathodic potentials (-0.35 V vs. NHE), indicating a deactivation of the active sites. Inset: The first and second sweep at anodic potentials showing a significant decrease in the oxidation peak.

tween major(basal plane area) and minor(edge area) peaks seen on Figure 6.9 corresponds to an approximate particle size of 25 nm. This is also in agreement with particle sizes of carbon supported MoS\(_2\) reported by Brorson et al[73]. We also checked whether the size of the major oxidation peak was correlated with the amount of Mo used during the preparation of the sample. We found a correlation where approx. 8.9 electrons (r\(^2\) = 0.55) were transferred during oxidation per Mo atom used in the synthesis.
6.3.3 XPS analysis

XPS analysis of the samples where also performed after synthesis and at different stages of electrochemical measurements. This would enable us to speculate further on the reaction mechanism in the oxidation of MoS$_2$.

Figure 6.10: XPS spectra of MoS$_2$ on Toray paper recorded at different stages of its life. 1): As prepared after sulfidation. 2a): After initial activity measurements of the HER (CVs between +0.1 and -0.4 V vs. NHE). 2b): Sample 2a after measurements of the HER and subsequent oxidation/deactivation (CVs between +1.4 and -0.4 V vs. NHE) and removal from the electrolyte at -0.32 V vs. NHE. 3): After measurements of the HER and subsequent oxidation/deactivation (CVs between +1.4 and -0.4 V vs. NHE) and removal from the electrolyte 5 at 0.4 V vs. NHE.

In Figure 6.10 XPS spectra of MoS$_2$ before(1) and after(2a) HER measurements is not changing significantly apart from the adsorption of sulfate on the surface. The spectra are also comparable with spectra recorded on MoS$_2$ on Au(111)[85]. The spectra does however change after oxidation at anodic potentials of up to 1.4 V vs. NHE, it is clear that the signals from Mo and S in MoS$_2$ are barely visible after oxidation. The exact product of the oxidation is more difficult to determine as the state of both pure Mo and S is
dependent on the electrochemical conditions[26]. A more detailed discussion of the exact products of the oxidation can be found in Paper 4[86], where it was proposed that the reaction mechanism was the following:

$$MoS_2 + 7H_2O \rightarrow MoO_3 + SO_4^{2-} + \frac{1}{2}S_2^{2-} + 14H^+ + 10e^- \quad (6.1)$$

This proposed reaction path allows us to calculate an exchange current density per area MoS$_2$ of $1.1 \times 10^{-6}$ A/cm$^2$, comparable with the $7.9 \times 10^{-6}$ A/cm$^2$ per edge area found on MoS$_2$ on Au(111). The exchange current density on the carbon supported MoS$_2$ can be seen as a lower bound as we normalize with area of MoS$_2$ and not the area of the active sites on the MoS$_2$. If we assume that the particles as previously mentioned have a particle size of around 25 nm, the edge will be approx. 8% of the area, leading to an exchange current density per cm$^2$ active site that is 12 times higher than the reported exchange current density per area of the particle.

### 6.4 The other metal sulfides

Since MoS$_2$ is not the only metal sulfide with under-coordinated sulfur atoms that might have promising HER characteristics, we opted to test a whole series of different metal sulfides. The other sulfides were tested for the HER in the fuel cell setup where the initial HER measurements on MoS$_2$ were performed.

![Figure 6.11: Polarisation on curve of the HER of different sulfides measured in the fuel cell setup with a scan rate of 5 mV/s and with a hydrogen flow of 20 mL/min. over both electrodes.](image)
The initial measurements on a choice of metal sulfides are exhibited in Figure 6.11. It can be observed that WS$_2$ apparently evolves hydrogen almost as efficient as MoS$_2$ whereas the sulfides based on Fe, Ni and Co apparently are worse HER catalysts. WS$_2$ is interesting as it has been found to have the same triangular structure as MoS$_2$ on graphite\[87\] and has previously been suggested as a catalysts for the HER\[88\]. A sample synthesized using a similar procedure as used when MoS$_2$ was synthesized with the main difference being that ammonium metatungstate was used as the precursor.

Figure 6.12 shows the electrochemical investigation of the carbon supported WS$_2$ nanoparticles. The WS$_2$ nanoparticles exhibit similar electrochemical characteristica to those previously found on MoS$_2$. The main difference is the Tafel slope that was found to be around to be 135 mV/dec. for this specific sample, but it can not be ruled out that this value will change if the synthesis procedure is further optimized.
6.5 Promotion of MoS\textsubscript{2} and WS\textsubscript{2}

Promotion of MoS\textsubscript{2} by Co is known to have an effect on the HDS activity and this system has been studied extensively for decades\textsuperscript{[89, 90, 62]}. Recently this system has also been studied using STM\textsuperscript{[75, 91]} and TEM\textsuperscript{[73]}. It is generally agreed that the promoter atoms are situated on the S-edge of MoS\textsubscript{2} \textsuperscript{[69, 62]}.

DFT calculations were used to estimate the effect of cobalt promoter atoms on the S-edge of MoS\textsubscript{2}. The results of this calculation are summarized in Figure 6.13. For the unpromoted system the Mo-edge of MoS\textsubscript{2} (\(\Delta G_H = 0.08\) eV) should have the highest activity, followed by the S-edge of MoS\textsubscript{2}(\(\Delta G_H = 0.18\) eV) and the pure W and S-edge of WS\textsubscript{2}(\(\Delta G_H = 0.2\) eV). The addition of Co to the S-edge of both the WS\textsubscript{2} and the MoS\textsubscript{2} shifts the value of \(\Delta G_H\) to 0.07 eV and 0.1 eV respectively. Thus we should expect that the addition of Co has a promoting effect.
The first system tested before any promotion experiments were conducted was the test of the HER on sulfided (same conditions as the MoS\(_2\)) Co acetate on Toray paper, see Figure 6.14. The initial sweep towards cathodic potentials showed a promising HER current but on subsequent sweeps the activity degraded rapidly towards the activity of the Toray paper support. Thus, it was concluded that CoS\(_x\) on its own was unstable under HER conditions.
We tested the Co promoted MoS$_2$ (Co-Mo-S) and the Co promoted WS$_2$ (Co-W-S) and it was found that the electrochemical behavior towards oxidation and hydrogen evolution was similar to that of the unpromoted catalysts, see Figure 6.15. In order to elucidate on whether there was an effect of the added Co we normalized the current with respect to the area of the respective oxidation peak (measured after HER measurements) of the specific samples.
Figure 6.16: Polarization curves where the currents of the different metal sulfides have been normalized with respect to the charge of the irreversible oxidation peak. A: Polarization curve of the HER on WS$_2$ and cobalt promoted WS$_2$(Co-W-S). B: Polarization curve of the HER on MoS$_2$ and cobalt promoted MoS$_2$(Co-Mo-S).

The polarization curves where the current had been normalized with respect to the area of the respective oxidation peaks (Figure 6.16) showed that there was an effect of Co promotion on both the WS$_2$ and MoS$_2$. If it is assumed that the particle size of the promoted and unpromoted metal sulfides are equal, we can explain the effect of the promotion in the following manner. The increase of the activity for the Co-Mo-S is due to the S-edge becoming just as active as the Mo-edge, and thus more active sites are created per particle. On the Co-W-S the S-edge is also becoming more active when Co is added and in this case a more active site is created. To sum up our findings it was found by DFT that the activity of the different edges should be WS$_2$ < MoS$_2$ = Co-Mo-S < Co-W-S whereas the electrochemical measurements on the samples indicate that WS$_2$ < MoS$_2$ < Co-W-S < Co-Mo-S. More details of the HER on metal sulfides are given in Paper 4[86].
6.6 \( \text{Mo}_3\text{S}_4 \)

\( \text{Mo}_3\text{S}_4 \) is one of the smallest molecules available where under coordinated sulfur sites are present. This system is particularly interesting as it can be considered as a basic building block for metal sulfide systems.

![Figure 6.17: The Mo cluster consisting of three molybdenum atoms and four sulfur atoms assembled in an incomplete cubane structure.](image)

The initial studies of the incomplete \( \text{Mo}_3\text{S}_4 \) cubanes were done in the fuel cell setup were the \( \text{Mo}_3\text{S}_4 \) clusters were deposited on high surface area carbon pressed into a MEA and tested for the HER.

![Figure 6.18: Electrochemical characterization of vulcan xc72 supported \( \text{Mo}_3\text{S}_4 \) in a membrane electrode assembly. The potential is measured with respect to a Pt electrode which serves as both the reference and the counter electrode. A: cyclic voltammogram in the anodic region, showing a reversible oxidation feature at ca. 0.7 V. B: Polarisation curves in the cathodic region where hydrogen is evolved. The activity of \( \text{Mo}_3\text{S}_4 \) is compared with a blank electrode only containing xc72, a sample with Pt and a sample with \( \text{MoS}_2 \). Hydrogen was passed over both electrodes and the scanrate was 5 mV/s.](image)

The result of the initial measurements show that the \( \text{Mo}_3\text{S}_4 \) clusters show HER activity that is similar to \( \text{MoS}_2 \), see Figure 6.18. The data also shows
that the HER activity is lowered for each consecutive sweep. At the same time the Mo$_3$S$_4$ cluster does not seem to be subject to irreversible oxidation at high anodic potentials. Which is the case for the MoS$_2$.

Just as in the study of Au(111) supported MoS$_2$\cite{92} we did electrochemical measurements on STM imaged samples, this time the support was activated HOPG with a submonolayer of Mo$_3$S$_4$. The measurements of the model system, see Figure 6.19, enables us calculate a turnover frequency per molecule, since the STM imaging resolved the number of Mo$_3$S$_4$ clusters on the surface of the HOPG. The turnover frequency per Mo$_3$S$_4$-molecule was found to be 0.07 s$^{-1}$. We do not know the exact number of active sites per Mo$_3$S$_4$-molecule, but this value is comparable to or even better than the turnover frequency of 0.02 s$^{-1}$ reported for a single site on the edge of MoS$_2$ nanoparticles\cite{83}. The turnover frequency is, however, smaller than for the turnover frequency per atom of the archetypical HER catalysts\cite{8} (Pt, Pd). XPS studies were also conducted on the supported clusters and it was found that the outmost layer of the Mo$_3$S$_4$ had a Mo 3d peak at a higher binding energy than the bulk of the sample. It was also found that the loss of activity was due to slow cathodic desorption of the clusters and not an oxidation of the Mo$_3$S$_4$ clusters.

Figure 6.19: Electrochemical characterization of the submonolayer of Mo$_3$S$_4$ on HOPG. Main: Polarization curve in the anodic region showing hydrogen evolution. The decrease in activity is most likely due to anodic desorption of the clusters. After a sweep to high cathodic potentials (1.2 V vs. NHE), the clusters desorb and the activity is only caused by the activated HOPG. Inset: the corresponding Tafel plot. The scanrate was 5 mV/s.
6.7 Hydrogen oxidation and metal sulfides

Some of the metal sulfides where also tested for the hydrogen oxidation reaction (HOR). It is well described that platinum effectively catalyzes the HER and the HOR and according to the basic Butler Volmer reaction, see equation 2.1 a good HER catalyst should be a good HOR catalyst, assuming that $\alpha_a = \alpha_c = 0.5$.

![Figure 6.20: Left: The HER and the HOR on the carbon supported MoS$_2$. Right: CV of carbon supported Mo$_3$S$_4$ in the region where HOR is expected to occur.](image)

In Figure 6.20 it is shown that neither MoS$_2$ nor Mo$_3$S$_4$ can oxidize hydrogen at measurable rates. The reason for this difference in the ability to catalyze the HER and the HOR could simply be that the two reactions have different free energy barriers due to different reaction paths. The reaction path for the HER could for instance be a Heyrovskí mechanism whereas the HOR could happen via the Tafel mechanism. Similar effects can also be seen on the PtAu core shell system that was able to evolve hydrogen but at the same time not able to oxidize hydrogen, see Figure 5.9. This notion about the difference in reaction paths is also supported by a recent study[93] where the hydrogen deuterium exchange rate was used to calculate the sticking probability of hydrogen on different metals. The catalyst with low values of $\Delta G_H$ apparently did not show the highest hydrogen deuterium exchange rates. Hopefully a more detailed study of the relation between the HER and the HOR could shed more light of these differences.
Chapter 7

Summary and outlook

In this thesis several alternative hydrogen evolution catalysts have been characterized and synthesized, some even to a degree where it was possible to correlate structure and reactivity of the catalysts. Based on a computational screening for HER catalyst a BiPt surface alloy was identified as a possible HER catalyst. The alloy was synthesized, characterized and found to have a higher HER activity than pure Pt. The surface alloy was also found to be more CO tolerant than pure Pt.

Investigations of a core shell system consisting of a Pt skin on Au was found to exhibit Pt like HER activity after being exposed to high anodic potentials (>1 V vs. NHE) and Au like HER activity at potentials below 1 V vs. NHE. The difference was ascribed as being a surface segregation effect.

MoS$_2$ nanoparticles were by DFT modeling found to have active sites similar to the active sites on the hydrogen evolving enzymes hydrogenase and nitrogenase. Carbon supported MoS$_2$ nanoparticles were synthesized and found to have a decent HER activity. In order to relate the structure of the MoS$_2$ to the HER activity, a series of electrochemical experiments were conducted on STM imaged MoS$_2$ supported on Au(111). These experiments corroborated the DFT calculations predicting that the edge of nanoparticulate MoS$_2$ is the active site for the HER.

Carbon supported MoS$_2$ was also synthesized and investigated for the HER. A method based on irreversible electrochemical oxidation of the MoS$_2$ was used to quantify the area of MoS$_2$, and it was found that the activity was comparable to the activity on MoS$_2$ supported on Au(111). WS$_2$ nanoparticles supported on carbon were also synthesized and characterized electrochemically. It was found that the HER activity of WS$_2$ was in the same region as the MoS$_2$. Inspired by Co promotion of WS$_2$ and MoS$_2$ in HDS catalysis, cobalt promoted carbon supported MoS$_2$ and WS$_2$ were synthesized and tested, in both cases Co was found to promote the HER.

Clusters of incomplete Mo$_3$S$_4$ cubane structures were also investigated for
their ability to catalyze the HER. Mo\(_3\)S\(_4\) was found to catalyze the HER and characterization of samples that were previously imaged by STM allowed the calculation of the turn over frequency per Mo\(_3\)S\(_4\) cluster.

Since computational screening methods have been shown as a strong tool to screen for HER catalyst, the next step would be to develop the model further in order to screen materials relevant for oxygen reduction reaction and the hydrogen oxidation reaction. The development of such a method is however not simple as for instance stability of the catalyst is far more crucial for ORR catalyst compared to the HER catalysts. An other extention of the work in thesis could be to characterize metal sulfides that were imaged by means of high-angle annular dark-field scanning electron microscopy in order to get further information of structure reactivity relations. Finally further work on the Mo\(_3\)S\(_4\) clusters could be of interest like the study of the effect of adding another metals to the incomplete cubane structure.
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61
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Included Publications

Paper 1:

“Biomimetic Hydrogen Evolution”
Journal of the American Chemical Society, 127, 5308 (2005)
Included Publications
The electrochemical hydrogen evolution process whereby protons and electrons are combined into molecular hydrogen is catalyzed most effectively by the Pt group metals.\(^1\) The interest in hydrogen evolution catalysts is currently increasing, as molecular hydrogen, \(\text{H}_2\), is being considered as an energy carrier.\(^2\) Unlike the hydrocarbon fuels used today, hydrogen produces only water during oxidation, for instance in a fuel cell. For hydrogen to be a real alternative to hydrocarbons, it must be produced in a sustainable fashion. One possibility is to use sunlight directly or indirectly (through wind power, for instance) to split water.\(^2\) This requires an efficient catalyst for hydrogen evolution, preferably based on materials that are cheap and abundant. It is therefore important to find alternatives to the Pt group metals.

Hydrogenases and nitrogenases are also effective catalysts for the hydrogen evolution process\(^3,4\) even though the catalytically active site of these enzymes contains the much less noble metals Fe, Ni, and Mo. Recently it has become possible to anchor hydrogenase to an electrode surface,\(^5\) and considerable progress has been made in the synthesis of compounds in solution resembling the hydrogenase active site and showing activity for hydrogen evolution.\(^6\)

In the present communication, we use density functional calculations to guide us to a new inorganic analogue of the other hydrogen-producing enzyme, nitrogenase. We analyze the difference between the metallic and the biological catalysts and show that in terms of being able to stabilize intermediates involving atomic hydrogen they have very similar properties. This allows us to identify a parameter that of this it is a necessary, but not sufficient, criterion for a material to be a good catalyst that the free energy of adsorbed H is close to that of the reactant or product (i.e., \(\Delta G^\circ_{\text{H}_2}\)). This principle can explain available experimental observations regarding metals as catalysts and electrode materials for hydrogen evolution.\(^7\)

Most water-splitting processes rely on electrochemical hydrogen evolution \(2\text{H}^+ + 2e^- \rightarrow \text{H}_2\) in the final step. The hydrogen evolution reaction must in the first step involve bonding of hydrogen to the catalyst \(\text{H}^+ + e^- \rightarrow \text{H}_2^\ast\), where \(\ast\) denotes a site on the surface able to bind to hydrogen. The second step is the release of molecular hydrogen through one of the two processes: \(2\text{H}^\ast \rightarrow \text{H}_2 + 2^\ast\) or \(\text{H}^\ast + e^- + \text{H}^\ast \rightarrow \text{H}_2 + ^\ast\).

Using density functional theory (DFT) calculations, we can elucidate the thermochemistry (which is independent on the precise mechanism of the second step) of the reaction; see Figure 1.\(^7\) By calculating the free energy of atomic hydrogen bonding to the catalyst, one can compare different metal surfaces as catalysts. For a chemical process to proceed at or around room temperature, no reaction step can be associated with large changes in the free energy. This immediately excludes the metals that form strong bonds to atomic hydrogen (Ni and Mo in Figure 1) as good catalysts because the hydrogen release step will be slow. Metals that do not bind to atomic hydrogen (Au in Figure 1) are also excluded because here the proton/electron-transfer step will be thermodynamically uphill and therefore slow. There could be extra energy barriers associated with the proton-transfer steps or \(\text{H}_2\) recombination, but independent of this it is a necessary, but not sufficient, criterion for a material to be a good catalyst that the free energy of adsorbed H is close to that of the reactant or product (i.e., \(\Delta G^\circ_{\text{H}_2}\)). This principle can explain available experimental observations regarding metals as catalysts and electrode materials for hydrogen evolution.\(^7\)

It is interesting to apply the same analysis to the active sites in nitrogenases and hydrogenases. For nitrogenase we have considered the model of the active site, the FeMo cofactor (FeMoco) with three hydrogen atoms bound at the equatorial \(\mu_3\text{S}\) sulfur atoms. (Middle) Hydrogenase active site with one hydrogen atom bound. The structure is taken from ref 11. (Right) MoS\(_2\) slab with sulfur monomers present at the Mo edge. The coverage is 50%, i.e., hydrogen is bound at every second sulfur atom. The lines mark the dimension of the unit cell in the \(x\)-direction.

and therefore slow. There could be extra energy barriers associated with the proton-transfer steps or \(\text{H}_2\) recombination, but independent of this it is a necessary, but not sufficient, criterion for a material to be a good catalyst that the free energy of adsorbed H is close to that of the reactant or product (i.e., \(\Delta G^\circ_{\text{H}_2}\)). This principle can explain available experimental observations regarding metals as catalysts and electrode materials for hydrogen evolution.\(^7\)
results in a binding energy close to that of Pt. The FeMoco thus complies with the $\Delta G^\circ = 0$ requirement.\textsuperscript{9} A number of compounds in crude oil\textsuperscript{13} and methods for their preparation can be understood from our findings. Furthermore, MoS\textsubscript{2} has been tested for photocatalytic hydrogen evolution and shows activity but with significantly lower currents.\textsuperscript{17}

Our findings suggest that we can begin searching for new catalytic materials using quantum chemical methods. The MoS\textsubscript{2} nanoparticles supported on graphite may be an example of a new class of electrode materials.

Acknowledgment. M. Brorson is gratefully thanked for providing us the samples. We thank P. E. M. Siegbahn for providing us results and structures prior to publication. We acknowledge support from the Danish Center of Scientific Computing through Grant No. HDW-1101-05.

Supporting Information Available: Details of the DFT calculations, experimental setup, and obtained data. This material is available free of charge via the Internet at http://pubs.acs.org.

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14. The difference in pH and chemical potential for electron transfer is of the same order and cancel each other; see Supporting Information.

Figure 3. (Left) Polarization curve for hydrogen evolution on Pt, daphoe C-support, and MoS\textsubscript{2} cathodes. The polarization curves for Pt and C support are made at 25 °C. The potentials are measured with respect to a carbon-supported Pt anode in a proton exchange membrane electrode assembly. (Right) STM images of MoS\textsubscript{2} nanoparticles on modified graphite.
Paper 2:

“Computational high-throughput screening of electrocatalytic materials for hydrogen evolution”

The pace of materials discovery for heterogeneous catalysts and electrocatalysts could, in principle, be accelerated by the development of efficient computational screening methods. This would require an integrated approach, where the catalytic activity and stability of new materials are evaluated and where predictions are benchmarked by careful synthesis and experimental tests.

In this contribution, we present a density functional theory-based, high-throughput screening scheme that successfully uses these strategies to identify a new electrocatalyst for the hydrogen evolution reaction (HER). The activity of over 700 binary surface alloys is evaluated theoretically; the stability of each alloy in electrochemical environments is also estimated. BiPt is found to have a predicted activity comparable to, or even better than, pure Pt, the archetypical HER catalyst. This alloy is synthesized and tested experimentally and shows improved HER performance compared with pure Pt, in agreement with the computational screening results.
principle, a general explanatory paradigm in heterogeneous catalysis and electrocatalysis that states that optimal catalytic activity can be achieved on a catalytic surface with intermediate binding energies (or free energies of adsorption) for reactive intermediates. If the intermediates bind too weakly, it is difficult for the surface to activate them, but if they bind too strongly, they will occupy all available surface sites and poison the reaction; intermediate binding energies permit a compromise between these extremes. In the particular case of hydrogen evolution, it turns out that these general principles can be quantified by analysing the free energy of hydrogen adsorption: this quantity is a reasonable descriptor of hydrogen evolution activity for a wide variety of metals and alloys.

As first suggested by Parson’s, the optimum value should be around $\Delta G_{i} = 0$. In Fig. 1, we have plotted experimentally measured HER exchange current densities against the calculated $\Delta G_{i}$ values calculated using DFT (see below). Although there is some scatter in the data (the experimental values are from several different authors), the figure seems to suggest that an optimum in the measured HER exchange current densities is found for DFT-derived $\Delta G_{i}$ values very close to (if not identically equal to) zero. We have therefore adopted the approach of using calculated values of $|\Delta G_{i}|$ to search for new HER catalysts, and we will assume in the following that the closer $|\Delta G_{i}|$ is to zero, the better the catalyst.

Using periodic $(\sqrt{3} \times \sqrt{3}) R30^\circ$ unit cell, self-consistent, DFT calculations (see the Methods section), we evaluate the value of $\Delta G_{i}$ on the 736 distinct binary transition-metal surface alloys that can be formed from the 16 metals Fe, Co, Ni, Cu, As, Ru, Rh, Pd, Ag, Cd, Sb, Re, Ir, Pt, Au and Bi (these elements are simply chosen to give a very broad pool of metallic and semimetallic elements, most of which are thermodynamically resistant to bulk oxide formation in water at 298 K and U = 0 V versus the standard hydrogen electrode; SHE). Such alloys, which are composed of a pure metal substrate with a solute element alloyed into the surface layer (Fig. 2), can exhibit surface properties that are vastly different from the properties of the bulk alloy. In Fig. 3, we show schematically the calculated free energies of hydrogen adsorption on a subset of these surface alloys—those with a 1/3 monolayer (ML) of solute in the surface layer (also see Fig. 2b). The figure clearly demonstrates that a number of binary surface alloys have high predicted activity for the HER.

Our purely computational screening procedure thus identifies a number of interesting candidates for HER catalysts. However, the analysis has so far neglected any consideration of whether or not the indicated alloys will be stable in real electrochemical environments. To estimate the stability of the surface alloys for the HER, we carry out four simple tests for each alloy (see the Supplementary Information). First, we estimate the free-energy change associated with surface segregation events; such events can cause surface solute atoms to segregate into the bulk. Second, we determine the free-energy change associated with intrasurface transformations such as island formation and surface de-alloying. Third, we evaluate the free energy of oxygen adsorption, beginning with splitting of liquid water; facile oxygen adsorption can lead to surface poisoning and/or oxide formation. Finally, we estimate the likelihood that the surface alloys of interest will corrode in acidic environments ($pH = 0$). For this test, we simply take the free energies of dissolution as reported in the electrochemical series.

In Fig. 4, we plot the most pessimistic of the free-energy transformation values determined for each alloy against the absolute magnitude of $|\Delta G_{i}|$. The stability considerations immediately eliminate a large number of alloys from consideration; although many alloys have high predicted HER activity, only a small fraction are predicted to be both active and stable in acidic HER environments. Including, among others, surface alloys of BiPt, Pt,$\text{Rh}$, Rh,$\text{Ir}$, Pt$\text{Ru}$, As,$\text{Au}$, Sb,$\text{Pt}$, Bi,$\text{Rh}$, Rh,$\text{Re}$, Pt,$\text{Ru}$, As,$\text{Rh}$, Rh,$\text{Ru}$, Ir,$\text{Ru}$ and Pt$\text{Rh}$ (note that the last seven of these have solute coverages different from 1/3 ML and therefore are not found in Fig. 3). Thus, these results demonstrate that stability considerations are essential for finding realistic candidate catalysts for hydrogen evolution. Furthermore, the results suggest that more detailed computational analyses and, ultimately, experimental testing of the promising alloys are in order.
The stabilization free-energy can be thought of as a free energy of formation for the surface alloys; the stability of the alloys with respect to various reconstructive/deactivating processes (including surface segregation, island formation, water splitting/oxygen adsorption and metal dissolution) is evaluated for each alloy, and the most pessimistic such energy (that is, the energy that would give the maximum probability that the alloy would destabilize) is plotted. The Pareto-optimal line indicates the best possible compromise between activity and stability, but given the simplicity of our model, other alloys could certainly be worth considering for use as HER catalysts; the alloys in the lower left quadrant, in particular, are promising. For the labelled points, a single element indicates a pure metal. For the bimetallic alloys, the solute is listed first; the solute coverages are 1/3 ML (BiPt), 1 ML (RhRe), 1/3 ML (CdRh) and 2/3 ML (IrFe).

For the present study, we have selected BiPt for further analysis. This surface alloy is particularly interesting for the HER because of the stark contrast between its two constituent elements; pure Pt exhibits high catalytic activity, whereas pure Bi is not active at all (see Fig. 1). A surface alloy formed from these two elements, however, yields a material predicted by the calculations to have an activity comparable to, or even better than, pure Pt. We note, in passing, that this ‘counterintuitive’ surface alloy might not have been investigated without the aid of a combinatorial screening approach to guide the search for new catalysts.

To study this material in more detail, we first carried out further calculations on more detailed structural models to assess the effect of Bi coverage on our results. These calculations were done at a Bi coverage of 1/4 ML (2 × 2 unit cells—see the Methods section); we note that Bi coverage equal to or greater than 1/2 ML caused very large distortions of the alloy surface and, thus, were not considered in detail. The coverage of adsorbed hydrogen was also 1/4 ML in these analyses. Although the preferred site preference for H on the BiPt surface alloy was different for the 2 × 2 versus (\(\sqrt{3}\) × \(\sqrt{3}\))R30° unit cells (face-centred cubic and near-top sites, respectively), the trends in \(\Delta G_{\text{H}1}\) values for the BiPt versus Pt slabs (−0.03 versus −0.07 eV on the (\(\sqrt{3}\) × \(\sqrt{3}\))R30° slab and 0.00 versus −0.04 eV on the 2 × 2 slab) were unchanged. For this more detailed analysis, we also carried out additional stability tests. We estimated the potential versus SHE at which hydroxyl will adsorb on the surface after water splitting; a value of ∼0.4 V was obtained by following the procedure described previously. Thus, although hydroxyl (which preferentially adsorbs on a Bi top site) adsorbs at a potential several tenths of a volt lower than atomic oxygen, even this more reactive species will not be present under typical HER conditions. We also evaluated the adatom formation energy for a BiPt surface alloy in the presence of 1/4 ML of adsorbed hydrogen. Adatom formation was found to be endothermic (∼0.07 eV) in the presence of adsorbed hydrogen, suggesting that under typical HER conditions, the stable form is that of the surface alloy, and there will only be a relatively small concentration of Bi adatoms present at equilibrium. These results, when combined with the activity and stability calculations described above, paint a picture of BiPt surface alloys as structurally stable systems that exhibit moderately improved hydrogen evolution kinetics compared with pure Pt. It is important to emphasize, however, that the HER model used is quite simple and that the calculated difference in \(\Delta G_{\text{H1}}\) for BiPt and Pt (∼0.04 eV) is small, given the typical accuracy associated with DFT calculations. Thus, it is only possible to conclude from the computational results that BiPt surface alloys might be interesting for the HER. To validate the results of our computational screening, careful experimental tests are clearly necessary.

Mixed Pt and Bi catalysts have been studied experimentally for decades. The two most commonly reported forms are: (1) a Pt surface modified by irreversibly adsorbed Bi (Pt–Bi\(_s\)) and (2) bulk intermetallic PtBi or PtBi\(_2\). Interest in bulk intermetallics has recently increased; they are typically investigated for electro-oxidation reactions (see, for example, ref. 23). Literature on the Pt–Bi\(_s\) system—the most commonly studied form of mixed Pt and Bi—extends back to the 1970s; many studies focus on the oxidation of \(\text{H}_2\), \(\text{CO}\) and formic acid, in addition to the reduction of \(\text{O}_2\) (refs 24–33). Perhaps the most pertinent Pt–Bi\(_s\) work is that of Gómez et al. who investigated the impact of Bi\(_s\) on hydrogen evolution for Pt(100) and Pt(111). They showed that Bi\(_s\) severely poisons both Pt surfaces for HER, as evidenced by a significant decrease in current as a function of Bi\(_s\) coverage. This result seems, at first glance, to contradict our theoretical prediction that mixed Pt and Bi alloys show improved activity for the HER. However, it should be recalled that our calculated BiPt
surface structure is a surface alloy and not Bi\textsubscript{upd} on Pt. By carefully controlling our electrochemical synthesis procedure (see discussion below), we create a surface alloy, and we show that this specific structure does, in fact, show improved activity for the HER. We note that, to our knowledge, no previous electrochemical studies on the Pt–Bi\textsubscript{upd} system have attempted to form a surface alloy from the Bi\textsubscript{upd} submonolayer.

To create the BiPt surface alloy, we used a three-step approach: (1) electrodeposition of an initial Pt film onto an inert support, (2) spontaneous deposition of a submonolayer of Bi\textsubscript{upd} by Bi underpotential deposition (UPD) and (3) annealing of the Pt–Bi\textsubscript{upd} precursor to form a BiPt surface alloy. For a given sample, we measured hydrogen evolution after each step of the synthesis. Following the work of Gómez et al., we used the cathodic current density at 0.0 V versus SHE (i\textsubscript{0V}) as the primary figure of merit for comparison—approximately the exchange current density\textsuperscript{26,27,34}. Current–voltage measurements of a representative sample are plotted in Fig. 5. After heating the initial polycrystalline Pt film in a furnace (see the Methods section), the sample demonstrated a current density of i\textsubscript{0V} = 5.6(10) \textsuperscript{−1} A cm\textsuperscript{−2} (−log i\textsubscript{0V} = 3.25), consistent with previous literature values for Pt and indicating that the starting material was prepared adequately with negligible contamination\textsuperscript{35}. We then adsorbed a submonolayer of Bi\textsubscript{upd} onto the Pt film by the same Bi UPD method used in previous reports (10 min, 1 mM Bi\textsubscript{3+}/0.5 M H\textsubscript{2}SO\textsubscript{4}), resulting in an estimated Bi\textsubscript{upd} coverage of \sim 10\% by X-ray photoelectron spectroscopy (data not shown). Immediately after Bi UPD, the measured activity of the Pt–Bi\textsubscript{upd} sample was considerably less than that of the initial Pt film—as expected, the Bi\textsubscript{upd} blocked Pt sites and poisoned the surface for hydrogen evolution\textsuperscript{24}. To complete the synthesis of a BiPt surface alloy, we annealed the Pt–Bi\textsubscript{upd} precursor to 500 °C in a tube furnace.

Heating the Pt–Bi\textsubscript{upd} system is the crucial step, as it provides the temperature necessary for Bi to overcome the kinetic barriers associated with alloying into the surface layer of Pt. Evidence of this process was reported by Paffett, Campbell and Taylor who conducted extensive studies of Bi adsorbed onto a Pt(111) crystal under ultrahigh vacuum conditions\textsuperscript{36,37}. They used low-energy electron diffraction and Auger electron spectroscopy to track morphological changes in a Bi adlayer as a function of coverage and temperature. One of the features they observed was that of a multilayered surface alloy—nominal composition Pt\textsubscript{1}Bi\textsubscript{2}—that formed at moderate temperatures (\sim 370 °C) after starting with a multilayer Bi film. These surface-science experiments clearly indicate that Bi can and will form a surface alloy with Pt at elevated temperatures. The synthesis procedure that we have developed shows that surface alloys can also be produced from electrochemically adsorbed species; this is a simple, low-cost and general method that can be applied to numerous electrochemical systems.

Having completed the annealing procedure, we again measured the hydrogen evolution rate. Figure 5 shows significantly improved current–voltage characteristics; quantitatively speaking, i\textsubscript{0V} is double that of the Pt–Bi\textsubscript{upd} precursor and, more importantly, \sim 50\% greater than i\textsubscript{0V} of the initial Pt film. The improved performance of this film cannot be attributed to a simple change in morphology or to increased surface area because cyclic voltammetry in the H UPD region after each stage of synthesis indicates, if anything, a decrease in Pt surface area (see the Supplementary Information). To verify that the final anneal did not influence the Pt itself, we conducted several control experiments on similarly prepared pure Pt samples through multiple heat treatments. These data are shown as an inset in Fig. 5 and demonstrate a negligible effect.

The three-step procedure used to synthesize the annealed sample is highly reproducible, and the results provide strong evidence that the final, annealed sample is indeed a BiPt surface alloy, consistent with the ultrahigh vacuum studies of Paffett et al.\textsuperscript{36,37}. It is worth noting that Bi itself is a notoriously poor electrocatalyst for the HER\textsuperscript{32}; so it is difficult to imagine a morphology involving Bi—other than a surface alloy—that could possibly improve the electrocatalytic activity of Pt for hydrogen evolution. Hence, we have shown that a combination of \textit{in silico} screening and careful experimental synthesis of promising candidates can lead to improved electrocatalysts, even when the identified metal combinations involve an element that is known to be inert for the HER, or when the metal combinations have previously been studied in different structural forms and have been found to be inactive.

We have thus confirmed experimentally the prediction that the HER activity of BiPt surface alloys should be comparable to that of pure Pt, or perhaps slightly better. This result suggests that our computational screening procedure is a promising technique for use in catalyst searches. The screening procedure can be viewed as a general, systematic, DFT-based method of incorporating both activity and stability criteria into the search for new metal alloy catalysts. As the accuracy and quality of kinetic models and DFT calculations improve, such \textit{in silico} combinatorial screening procedures should become broadly useful for catalytic materials discovery.

**METHODS**

The computational analysis is carried out using DACapo\textsuperscript{37}, a total energy calculation code. For the high-throughput computational screening, a three-layer slab, periodically repeated in a super-cell geometry with five equivalent layers of vacuum between any two successive metal slabs, is used to determine hydrogen binding energies. Close-packed surfaces are considered in all cases. A (√3 × √3)R30° unit cell, corresponding to a hydrogen coverage of 1/3 ML, is used. With this unit cell, and using the 16 elements described in the text, it is possible to create 736 symmetrically distinct pure metal and binary surface alloy slabs—16 pure metals, 240 surface alloys with a pure overlayer of solute metal, 240 with only two solute atoms in the surface layer and an extra 240 with one solute atom in the surface layer. For all slabs, the metal atoms are kept fixed in their bulk-truncated positions, and the hydrogen atoms are
allowed to relax until the total force is less than 0.5 eV Å⁻¹. The total energy is then further refined by using the residual forces and estimated harmonic vibrational frequencies for hydrogen (∼1.05 cm⁻¹ for three-fold sites and ∼2.02 cm⁻¹ for top sites) to extrapolate to the bottom of a parabola in energy space. 0.24 eV is added to the calculated binding energies (with respect to gaseous H₂) to give adsorption free energies. The oxygen binding energies (used for the oxygen adsorption stability criterion) are computed in the same way, except that a slightly higher force cutoff (0.6 eV Å⁻¹) is used and appropriate vibrational frequencies (∼500 cm⁻¹ for three-fold sites and ∼775 cm⁻¹ for top sites) are used for the energy extrapolations.

For all DFT calculations, adsorption is allowed on only one of the two exposed surfaces of the metal slabs, and the electrostatic potential is adjusted accordingly. Ionic cores are described by ultrasoft pseudopotentials, and the Kohn–Sham one-electron valence states are expanded in a basis of plane waves with kinetic energy below 340 eV; a density cutoff of 500 eV is used. The surface Brillouin zone is sampled with an 18(√3 × √3) p(Chadi–Cohen) k point grid. In all cases, convergence of the total energy with respect to the cutoff energies and the p point set is confirmed. The exchange–correlation energy and potential are described by the generalized gradient approximation (GGA–RPBE98). The self-consistent RPBE98 density is determined by iterative diagonalization of the Kohn–Sham Hamiltonian, Fermi population of the Kohn–Sham states (k_T = 0.1 eV) and Puley mixing of the resulting electronic density. All total energies have been extrapolated to k_T = 0.0 eV. Zero-point energy effects are assumed to be approximately constant for all metals and alloys considered. Spin–polarization effects are included in the reported results for alloys in which naturally magnetized metals (Ni, Co, Fe) are present. Graphical inserts are produced using visual molecular dynamics.

For the more detailed calculations on the BiPt surface alloy, a 2 × 2 surface unit cell is used, corresponding to a hydrogen coverage of 1/4 ML. A four-layer slab is used for these calculations, and the top two layers of the slab are allowed to relax. The maximum force permitted for any vector component is 0.05 eV Å⁻¹. The k point grid in this case, involves 18(1 × 1)

Chadi–Cohen points.

Pt was electrodeposited from a 1 mM H₂PtCl₆ (99.995%, Aldrich) solution onto an inert substrate—fluorine-doped tin-oxide coated onto glass (Hartmann Glass). A square wave pulse deposition was used, referenced to a Ag quasi-reference electrode (AgQRE). The three-step deposition sequence consisted of a 0.5 s pulse at −0.9 V versus AgQRE, followed by a 0.5 s pulse at −0.2 V versus AgQRE, followed by a 1 s rest interval. The sequence was repeated 120 times, totalling 4 min. The AgQRE was produced by oxidizing a 1-mm Ag rod and recording the current density versus potential. For these experiments, a Pt(110)-(1×2): The influence of bismuth on hydrogen and oxygen adsorption, and the electro-oxidation of carbon monoxide. J. Electroanal. Chem. 224, 226−232 (1987).


Acknowledgements JG and T.E.J. acknowledge H.C. Ørsted Postdoctoral Fellowships from the Technical University of Denmark. J.B. acknowledges support from the Danish Strategic Research Council. The Center for Individual Nanoparticle Functionality is supported by the Danish National Research Foundation. The Center for Atomic-scale Materials Design is supported by the Lundbeck Foundation. We thank the Danish Center for Scientific Computing for computer time. We also thank K. P. Jorgensen and J. Larsen for technical assistance. Correspondence and requests for materials should be addressed to J.K.N.

Supplementary Information accompanies this paper on www.nature.com/naturerelials.

Competing financial interests

The authors declare that they have no competing financial interest.

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Paper 3:

“Identification of Active Edge Sites for Electrochemical $H_2$ Evolution from MoS$_2$ Nanocatalysts”

Science, 316, 100 (2007)
Identification of Active Edge Sites for Electrochemical H₂ Evolution from MoS₂ Nanocatalysts

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The identification of the active sites in heterogeneous catalysis requires a combination of surface sensitive methods and reactivity studies. We determined the active site for hydrogen evolution, a reaction catalyzed by precious metals, on nanoparticulate molybdenum disulfide (MoS₂) by atomically resolving the surface of this catalyst before measuring electrochemical activity in solution. By preparing MoS₂ nanoparticles of different sizes, we systematically varied the distribution of surface sites on MoS₂ nanoparticles on Au(111), which we quantified with scanning tunneling microscopy. Electrocatalytic activity measurements for hydrogen evolution correlate linearly with the number of edge sites on the MoS₂ catalyst.

Progress in the field of heterogeneous catalysis is often hampered by the difficulty of identifying the active site on a catalyst surface (1, 2). In homogeneous catalysis, the active center is generally more clearly defined and quantified, with spectroscopic and mechanistic studies providing direct insight into reactive intermediates. Solid-state catalysts, however, commonly exhibit a variety of different surface sites that are difficult to identify and quantify; the scenario is further complicated when multiple sites work together in turning over a reaction. Identifying the most active site(s) is critical to designing and developing improved catalytic materials. Many useful in situ and ex situ experimental techniques, as well as computational methods, have been developed (3–5) to address this problem, but identifying the active site remains a challenging task.

In this study we used such methods to determine the active site of nanoparticulate MoS₂ for the hydrogen evolution reaction (HER), 2H₂ + 2e⁻ → H₂ (6, 7), which is fundamentally important for a variety of electrochemical processes, fuel cells (as the reverse reaction), and solar H₂ evolution (8). Despite numerous studies on this material, there is a call for studies that uniquely link activity to specific surface sites, in that MoS₂ is used industrially as a hydrodesulfurization (HDS) catalyst (10, 11). Detailed insight has been gained from studies on simplified model systems in ultra-high vacuum (UHV) and by using computational methods (12–15), as well as from combining reactivity measurements and ex situ characterization of industrial catalyst samples (10, 11, 16). Structural studies on the MoS₂ catalyst have shown that it is composed almost entirely of flat polygons of S-Mo-S trilayers (10); depending on the synthesis conditions, these trilayers may stack in a graphite-like manner or remain as single trilayers. For single trilayers, two general kinds of surface sites exist—terrace sites, which are those on the basal plane, and edge sites, which lie at the edge of the nanoparticles. DFT studies suggest that the active site for HDS is on the edge of the MoS₂ nanoparticles. This result is supported by adsorption studies of thiophene using scanning tunneling microscopy (STM) (17).

Immediately after deposition, each sample was vacuum transferred to a second UHV chamber for STM imaging (Fig. 1). The crystallized, single-layered MoS₂ nanoparticles can be described as flat polygons with a conducting perimeter. Comparison of representative images of samples annealed at 400°C (Fig. 1A) and 550°C (Fig. 1B) shows how particle size increased after sintering at the higher temperature. The particles annealed to 400°C are consistent with similarly prepared MoS₂ nanoparticles on Au(111) (13). Besenbacher et al. have shown that the dominant edge structure of MoS₂ nanoparticles is that of a sulfided Mo edge and that this edge is particularly favored by larger-sized particles (12, 18). We also observe the predominance of the sulfided Mo-edge in our samples, regardless of annealing temperature. Thus, controlled sintering allows us to change the ratio of basal plane sites to edge sites without changing the nature of this edge. This sulfided (1 0 –1 0)

Fig. 1. A series of STM images of MoS₂ nanoparticles on Au(111). The particles exhibit the typical polygon morphology with conducting edge states and are dispersed on the Au surface irrespective of coverage and annealing temperature (400°C or 550°C). (A) Low coverage (0.06 nm² MoS₂/ nm² Au(111)), annealed to 400°C (470 Å by 470 Å, 1.2 nA, 4 mV). (B) High coverage (0.23 nm² MoS₂/ nm² Au(111)), annealed to 550°C (470 Å by 470 Å, 1.2 nA, 1.9 V). (C) Atomically resolved MoS₂ particle, from a sample annealed to 550°C, showing the predominance of the sulfided Mo-edge (19, 20) (60 Å by 60 Å, 1.0 nA, 300 mV).
Mo-edge is the same structure predicted by DFT calculations to be the active site for H2 evolution (8).

After imaging, we transfer the samples from UHV into an electrochemical cell to measure HER activity (21). Polarization curves (i – E) within a cathodic potential window, and corresponding Tafel plots (log i – E), are shown in Fig. 2. Current densities are normalized to the geometric area of the exposed face of all samples. The most inherent measure of activity for the HER is the exchange current density, i0 (6, 7, 22, 23), which is determined by fitting i – E data to the Tafel equation (6), yielding Tafel slopes of 55 to 60 mV/decade and exchange current densities in the range of 1.3 × 10–7 to 3.1 × 10–7 A/cm2, for all MoS2 samples (table S1). In Fig. 3, we plot the exchange current density for each sample versus two sample parameters, the MoS2 area coverage (Fig. 3A), and the MoS2 edge state length (Fig. 3B). The data points fall on a straight line only when plotted versus edge length. Although the points show some scatter around this trend, they are described by a best-fit linear relation with a slope of 1.67 × 10–20 A/(nm × MoS2-edge).

Because the rate of reaction is directly proportional to the number of edge sites for all samples, regardless of particle size, we conclude that the edge site is indeed the active site (24). Bearing this in mind, we note in Fig. 3A that the exchange current densities of the samples sintered at 550°C are significantly lower than those prepared at 400°C, per MoS2 coverage, exactly as one would expect considering that the sintered samples have less edge length per area of MoS2.

We also compared nanoscale MoS2 to other materials that catalyze the HER on a per active site basis (f). For this direct site-to-site comparison, we used the 1.5 × 1015 sites/cm2 for the Pt(111) face as the basis for comparison as Pt is the archetypical HER catalyst (25). An exchange current density of 4.5 × 10–4 A/cm2 for this face (26) yields a turnover frequency (TOF) of 0.9 s–1 (table S2). In general, TOFs of transition metals range over 10 orders of magnitude (Hg, for instance, has a TOF as low as ~10–15 s–1) (22). Given the slope in Fig. 3B, we have calculated the TOF of the MoS2 edge to be 0.02 s–1, indeed in the high range of TOFs for metals.

For further insight into the catalytic nature of the MoS2 edge, we have added our data for nanoparticulate MoS2 to a recent version of the volcano-type relations observed for HER catalysts (Fig. 4), in this case for the Gibbs free energy for atomic hydrogen adsorption (ΔGθH) (22, 23). These volcano relations ultimately reflect the Sabatier principle, which accounts for optimal surfaces as ones that exhibit moderate binding energies of reaction intermediates, hydrogen adsorption in the case of the HER. In Fig. 4, the exchange current density is shown as a function of the DFT-calculated free energy of adsorption of hydrogen, which was recently determined to be +0.08 eV for the MoS2 edge (8). To add MoS2 to this figure, we converted the TOF of nanoparticulate MoS2 to its exchange current density per 1.5 × 1015 sites/cm2, which yields 7.9 × 10–6 A/cm2 (table S2). This value surpasses those of the common metals and lies just below those of the precious Pt-group metals. When plotting this experimentally determined activity of the edge site versus its DFT-calculated ΔGθH (8), we see that it follows the volcano trend (23). This agreement validates the predictive capability of this DFT model as well as its applicability beyond metal catalysts.

After identifying the active site and comparing it with typical metal catalysts, we may...
consider how to improve its activity. The DFT-calculated ΔG_H of the MoS_2 edge site is slightly positive at ~0.08 eV, with calculations suggesting an H coverage of only one-quarter on the edge under operating conditions (8). Thus, only 1 in 4 edge atoms evolves molecular H_2 at a given time, unlike Pt(111) which operates at a H-coverage of ~1 ML (7, 26, 27). If all MoS_2 edge sites could be made to adsorb H, activity could be increased by a factor of 4. This might be accomplished by appropriately tuning the electronic structure of the edge to increase the bond strength of the adsorbed H (23). Such a modification could simultaneously improve the inherent turnover of each edge site, further improving the overall activity of the material toward that of Pt-group metals.

References and Notes
19. Materials and methods are available as supporting material on Science Online.
21. The cell, specifically designed for studies on UHV-transferred samples, was sealed under the ion gun (11).
22. An activated H sample, with a viton o-ring, exposing ~0.10 cm² to the H_2SO_4 electrolyte (pH 0.24, 23 °C) and cyclic voltammograms were recorded. This procedure ensures a one-to-one correlation between the imaged Mo,S nanoparticles and the measured activity for hydrogen evolution.
25. We note that Au atoms along the particle edge also scale with the edge length; however, previous experimental and computational studies have shown negligible interaction between the MoS_2 and the support (34, 15), leading us to conclude that such sites would be as inactive as those of the blank samples, prepared without MoS_2 deposition.
26. In our comparison with Pt(111), we assumed that the observed HER on this surface is dominated by terrace atoms, not defects. This is the consensus in the literature (7).
28. This project was supported by the Danish Strategic Research Council. T.F.J. acknowledges an H. C. Ørsted Postdoctoral Fellowship from the Technical University of Denmark. The Center for Individual Nanoparticle Functionality is supported by the Danish National Research Foundation. The Center for Atomic-scale Materials Design is supported by the Lundbeck Foundation.

Supporting Online Material
www.sciencemag.org/cgi/content/full/317/5834/102/DC1
Materials and Methods
SOM Text
Fig. S1
Tables S1 and S2
References
20 February 2007; accepted 16 May 2007
10.1126/science.1141483

Understanding Reactivity at Very Low Temperatures: The Reactions of Oxygen Atoms with Alkenes

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A remarkable number of reactions between neutral free radicals and neutral molecules have been shown to remain rapid down to temperatures as low as 20 kelvin. The rate coefficients generally increase as the temperature is lowered. We examined the reasons for this temperature dependence through a combined experimental and theoretical study of the reactions of O(1D) atoms with a range of alkenes. The factors that control the rate coefficients were shown to be rather subtle, but excellent agreement was obtained between the experimental results and microcanonical transition state theory calculations based on ab initio representations of the potential energy surfaces describing the interaction between the reactants.

Application of the CRESU technique (1) has shown that a surprising number of biomolecular reactions between neutral gas-phase species are rapid at very low temperatures. To date, rate coefficients for some 45 neutral-neutral reactions have been measured (2, 3), in some cases at temperatures as low as 13 K (4).

All have rate coefficients at 298 K [k(298 K)] that are equal to or exceed ~10^{-11} cm³ molecule⁻¹ s⁻¹. Moreover, the general trend with temperature is for the rate coefficients to increase as the temperature is lowered.

These observations have led to a reevaluation of the chemistry that occurs in the cold cores (10 to 20 K) of dense, dark interstellar clouds (ISCs), where the majority of interstellar molecules have been identified. Although sequences of ion-molecule reactions initiated by the cosmic ray-induced ionization of H_2 clearly play a central role in this chemistry, neutral-neutral reactions are now expected to be more important than previously thought (5).

Unfortunately, the kinetic database required for detailed astrochemical modeling (6) is still far from complete. There are many reactions that may occur in ISCs, such as those between pairs of unstable species, for which low-temperature kinetic data neither exist nor are likely to be obtained in the foreseeable future. Theoretical or semi-empirical methods of estimating these rate coefficients are therefore desirable.

Several complementary theoretical treatments (7–9) have been advanced to explain the observed negative temperature dependences of rate coefficients for radical-radical reactions, principally on the basis of the notion of adiabatic capture of the reactants via long-range attractive forces. Although they differ in their details, these treatments all predict large rate coefficients at very low temperatures for radical-radical reactions, where there is generally no barrier on the minimum energy reaction path. In the case of radical-molecule reactions, a key issue is whether a potential energy barrier exists along the minimum energy path from reactants to products: either a real barrier (i.e., a maximum above the energy of the separated reactants) or a “submerged” barrier corresponding to a maximum along the minimum energy path between the shallow minimum associated with a prereaction complex and the products (see below). Further theoretical work, particularly by Georgievskii and Klippenstein (10, 11), has shown that a submerged barrier can serve as a second inner transition state (or bottleneck), because the internal states at this smaller interreactant separation are more widely spaced than at the outer, capture transition state. In these circumstances, the rate of reaction falls below that predicted by capture theories, and a version of microcanonical transi-
Paper 4:

“Hydrogen Evolution on Nano-particulate Transition Metal Sulfides”
J. Bonde, P. G. Moses, T. F. Jaramillo, J. K. Nørskov and I. Chorkendorff,
Accepted Faraday Discussions, (2008)
Hydrogen Evolution on Nano-particulate Transition Metal Sulfides

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Received (in XXX, XXX) 1st January 2007, Accepted 1st January 2007
First published on the web 1st January 2007
DOI: 10.1039/b000000x

The Hydrogen Evolution Reaction (HER) on carbon supported MoS\textsubscript{2} nanoparticles is investigated and compared to findings with previously published work on Au(111) supported MoS\textsubscript{2}. An investigation into MoS\textsubscript{2} oxidation is presented and used to quantify the surface concentration of MoS\textsubscript{2}. Other metal sulfides with morphologies similar to MoS\textsubscript{2} such as WS\textsubscript{2}, cobalt-promoted WS\textsubscript{2}, and cobalt-promoted MoS\textsubscript{2} were also investigated in the search for improved HER activity. Experimental findings are compared to Density Functional Theory (DFT) calculated values for the hydrogen binding energies (\(\Delta G_{\text{H}}\)) on each system.

Introduction

Research efforts to develop electrocatalysts for energy conversion reactions have increased substantially in recent years. Platinum, the ubiquitous electrocatalyst used in PEM fuel cells, is both expensive and scarce, prompting widespread efforts to discover cost-effective materials to replace Pt. In this work we focus on non-noble metal sulfide catalysts for the Hydrogen Evolution Reaction (HER) under acidic conditions, a reaction catalyzed most effectively by Pt-based materials\textsuperscript{1}. Previously, MoS\textsubscript{2} has been studied as a catalyst in hydrodesulfurisation\textsuperscript{2} and in the photo-oxidation of organics\textsuperscript{3,4}. In electrocatalysis, it has recently been shown that the edge structure of nanoparticulate MoS\textsubscript{2} is active for the HER, mimicking the active sites/co-factor of the hydrogen evolving enzymes nitrogenase and hydrogenase\textsuperscript{5,6}. This work aims to extend the investigation on carbon-supported nanoparticulate MoS\textsubscript{2} for the HER. Unlike the case of Au(111) supported MoS\textsubscript{2} studied by STM in previous work\textsuperscript{6}, the catalysts probed herein are more commercially relevant, which also implies that they are less homogeneous and more difficult to image on the atomic scale. As knowledge of the concentration of active sites on a catalyst surface is paramount to elucidating structure-composition-activity relationships, the first aim of this work is to utilize electrochemical oxidation to probe MoS\textsubscript{2} surface area, distinguishing between basal plane and edge sites. In developing this methodology to quantify active sites on a macroscopic scale, we then direct our attention to related catalyst systems, namely WS\textsubscript{2}, cobalt-promoted WS\textsubscript{2}, and cobalt-promoted MoS\textsubscript{2}. We end by comparing experimentally determined activity data to predictions made by Density Functional Theory (DFT) models of these systems in order to gain insight into trends in catalyst activity.

It has been found that \(\Delta G_{\text{H}}\), the hydrogen binding energy to a given surface, is a good descriptor for identifying electrocatalyst materials with high exchange current densities\textsuperscript{1,7,8}. A recent study\textsuperscript{9} using DFT showed that the active sites on nitrogenase and hydrogenase bind hydrogen weakly, similar to Pt. It was also found that the overpotential of carbon supported MoS\textsubscript{2} is comparable to the DFT calculated hydrogen binding energy on the edge of the nanoparticles. In another study MoS\textsubscript{2} nanoparticles on Au(111) were synthesized under UHV conditions, characterized by STM and examined for HER activity\textsuperscript{9}. This study showed direct evidence that the active site of the MoS\textsubscript{2} nanoparticles is indeed the edge. The exchange current density was also found to be in agreement with the volcano relation between the HER exchange current density and the DFT calculated values for \(\Delta G_{\text{H}}\) proposed by Nørskov et al\textsuperscript{10}. By having identified the active site of MoS\textsubscript{2} particles, the next step is to modify that edge such that its \(\Delta G_{\text{H}}\) approaches even closer to zero where the HER volcano curve has its maximum, and this is a major aim of the work presented herein.

Bulk MoS\textsubscript{2} consists of stacked S-Mo-S layers, and MoS\textsubscript{2} nanoparticles can be synthesized as single layer hexagonal structures exposing two different kinds of edges, the so-called Mo-edge and the S-edge\textsuperscript{2}. It has been shown that the structure of nanoparticulate MoS\textsubscript{2} is a single layered truncated triangle primarily exposing the Mo-edge when supported on Au(111)\textsuperscript{5}. Highly Ordered Pyrolytic Graphite (HOPG)\textsuperscript{10,11} or graphitic carbon\textsuperscript{11,12} also found truncated triangles by means of HAADF-STEM (High-Annular Dark-Field-Scanning Transmission Electron Microscopy) in their investigation of MoS\textsubscript{2}, WS\textsubscript{2} and cobalt-promoted MoS\textsubscript{2}.

Estimating the number of active sites on a nanoparticulate catalyst is not trivial. One approach is to measure activity on well defined model systems characterized by STM, for example UHV-deposited nanoparticles\textsuperscript{6,13} or chemisorbed molecular clusters\textsuperscript{14}. Another option is to use a well established method to measure electrochemically active surface area such as that used with Pt based on the adsorption-desorption behavior of underpotentially deposited hydrogen,
We note, however, that this method still relies upon the assumption that the sites active for Hupd are the same as those active for the HER. As we are studying metal sulfides where no such method exists, the irreversible oxidation of metal sulfides will be investigated as a measure of their surface area and edge sites.

In the following we will show our investigation of MoS$_2$, WS$_2$ and cobalt promoted WS$_2$ (Co-W-S) and MoS$_2$ (Co-Mo-S), prepared similarly to the ones imaged by Brorson et al. and supported on Toray carbon paper. The electrochemical measurements will be discussed in relation to DFT calculations of $\Delta G_{H}$ for each of the metal sulfates investigated in order to identify structure-composition-activity relationships for these systems.

**Results and discussion:**

**Synthesis and electrochemical characterization of MoS$_2$.**

MoS$_2$ particles on Toray carbon paper were prepared by dropping 25 µL of an aqueous ammonia heptamolybdate (1mM Mo) solution onto 1 cm$^2$ of Toray paper. The sample was dried in air at 140°C followed by sulfidation in 10% H$_2$S in H$_2$ at 450°C for 4 hours, and subsequently cooled in that same gas stream. This preparation method would typically give the highest current on a per gram basis; higher loadings usually led to lower currents.

**Fig. 1** Tafel plot (main) and polarization curve (inset) in the cathodic potential range of MoS$_2$ supported on Toray paper. The scan rate is 5 mV/s and the Tafel slope in the HER region is found to be 120 mV/dec. HER activity was measured (See experimental details) and the results are plotted as Tafel (log i - E) and polarization curves (i-E) in Figure 1. The Tafel plot exhibits a slope of 120 mV/dec and an exchange current density of 4.6 × 10$^{-6}$ A/cm$^2$ geometric. Samples prepared by different methods have often yielded different Tafel slopes, ranging between 110 mV/dec, and several hundred mV/dec. We attribute this to transport limitations through the fibrous, porous network characteristic of Toray carbon paper. Although sample/substrate preparation could potentially be optimized further, the consistent results achieved using the preparation method described above allows for accurate cross-comparisons among different catalyst materials. It should be noted that hydrogen evolution is taking off at around -0.2 V vs. NHE just as we have previously seen on MoS$_2$.

The current measured from approx. +0.1 V vs. NHE to -0.15 V is most likely not due to the HER but rather oxygen reduction at the interface between the electrolyte and the electrode. Finally it should be noted that sweeps between -0.35 and +0.1 V vs. NHE showed negligible change over time, apart from the effects of bubble formation on the electrode.

**Fig. 2** Cyclic voltammogram of the oxidation and subsequent deactivation of the MoS$_2$ sample. Scanrate 2mV/s. Main: The deactivation of the sample showing one sweep from -0.35 V vs. NHE to 1.05 V vs NHE and back to -0.35 V vs. NHE. On the 1st anodic sweep an irreversible oxidation peak occurs at 0.6 V vs. NHE and is followed by a subsequent decrease in current at cathodic potentials (-0.35 V vs. NHE), indicating a deactivation of the active sites. Inset: The first and second sweep at anodic potentials showing a significant decrease in the oxidation peak.

Figure 2 shows a cyclic voltammogram of MoS$_2$/C where the potential is cycled between -0.3 and +1.05 V vs. NHE. At approx. +0.6 V vs. NHE an irreversible oxidation begins to occur with a maximum at +0.98 V vs. NHE. On the subsequent cathodic sweep a significant drop in HER activity is noticed. On the ensuing anodic sweep seen in the inset of Figure 2, the oxidation peak is no longer present. Thus, the loss of HER activity is attributed to irreversible MoS$_2$ oxidation. In subsequent studies, fresh samples were subjected to CVs in which an initially narrow potential window was widened gradually to more positive (anodic) potentials. It was found that the HER activity of MoS$_2$/C remained stable with every sweep as long as the anodic potential was limited to ≤ +0.6 V vs. NHE.
**MoS$_2$ electro-oxidation**

To our knowledge, the electrochemical oxidation of nanoparticulate MoS$_2$ is not covered in the literature, which instead focuses on the corrosion of bulk MoS$_2$. Kautek$^{16}$ found that the bulk system preferentially oxidized at the $\{10\bar{1}0\}$ face and that it did not corrode at the $(0001)$ basal plane. On the nanoparticles this would correspond to corrosion of the particle edges. Closer examination of the insert of Figure 2 reveals two distinct oxidation peaks. The major peak has its maximum at approx. +0.98 V vs. NHE whereas the minor peak has its maximum at approx. +0.7 V vs. NHE. As the edges of MoS$_2$ nanoparticles are expected to be more readily oxidized than the basal plane$^{16}$, we interpret the two distinct oxidation peaks to correspond to the edges (minor peak, +0.7 V vs. NHE) and the basal planes (major peak, +0.98 V vs. NHE) of the particles. While only one cycle to +1.05 V vs. NHE will completely deactivate the sample for the HER, it takes several cycles to +0.7 V vs. NHE to achieve the same effect. This implies that not all edge sites are oxidized with a single sweep to +0.7 V vs. NHE. Had the sample been deactivated for the HER after a single sweep to +0.7 V vs. NHE, we could definitely have used this peak as a measure of the concentration of edge sites. But as this is not the case we will use the major peak at +0.98 V vs. NHE to determine the total surface area of MoS$_2$/C. We have however attempted to use the weak feature at +0.7 V vs. NHE to get an estimate of our particle size. At low sweep rates (2 mV/s) the feature is typically not dominated by the major feature at +0.98 V vs. NHE. The area of the edge feature is approx. 8% of the major peak. If the particles are triangular this corresponds to an edge length of around 25 nm, consistent with the particle sizes observed by Brorson et al$^{11}$.

XPS was also employed in this investigation to study the MoS$_2$/C at three stages of its life: freshly prepared, after HER in H$_2$SO$_4$ and after oxidation in H$_2$SO$_4$ at high anodic potentials (see experimental details). To obtain a reasonable signal to noise ratio for the XPS studies the Toray paper was dip coated in a 0.14 M Mo solution instead of dropping a known amount of solution on the surface, resulting in a higher loading of Mo than previously described (a factor of 5-10 according to the charge of the oxidation peak). The survey spectra of the different samples showed no contaminants on the freshly prepared samples. On the samples that had been submerged in H$_2$SO$_4$ peaks corresponding to sulfate were seen and a peak corresponding to N $1s$ was also seen. The N $1s$ peak is most likely caused by trace amounts of NH$_3$ present in air absorbed by H$_2$SO$_4$ as (NH$_4$)$_2$SO$_4$ with a N $1s$ binding energy of 401.3 eV$^{17}$. 

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16 Kautek, 17 Brorson et al.
Fig. 3 XPS spectra of MoS₂ on Toray paper recorded at different stages of its life. 1): As prepared after sulfidation. 2a): After initial activity measurements of the HER (CVs between +0.1 and -0.4 V vs. NHE). 2b): Sample 2a after measurements of the HER and subsequent oxidation/deactivation (CVs between +1.4 and -0.4 V vs. NHE) and removal from the electrolyte at -0.32 V vs. NHE. 3): After measurements of the HER and subsequent oxidation/deactivation (CVs between +1.4 and -0.4 V vs. NHE) and removal from the electrolyte at 0.4 V vs. NHE.

The XPS data, see Figure 3, reveals that the freshly prepared sample (no. 1) of MoS₂ is similar to previously reported spectra. The XPS data from a similarly prepared sample that was tested for the HER (sample no. 2a) by sweeping the potential between +0.1 V and -0.45 V vs. NHE, showed an increase in the SO₄²⁻ peak which is to be expected as the sample had been submerged in H₂SO₄. Apart from the increase in the SO₄²⁻ peak no significant changes were found compared to the freshly prepared sample, indicating that MoS₂/C does not change significantly during the HER. After XPS analysis of sample no. 2a was examined for the HER again, then cycled between -0.4 V vs. NHE and +1.4 V vs. NHE and removed from the solution at -0.32 V vs. NHE (sample no. 2b in the XPS spectra). A significant decrease of the Mo 3d, Mo 3p, S 2s and S 2p peaks was observed and there was no XPS signal corresponding to MoO₃. Thus, although the amount of surface Mo decreased significantly it still maintained its Mo⁴⁺ character (as in MoS₂). There are several possible explanations for the lack of Mo on the surface: (1) The MoS₂ desorbs from the surface at high anodic potentials, (2) The oxidation product of MoS₂, MoO₃, dissolves, (3) That MoO₃ is reduced to Mo³⁺ at -0.32 V vs. NHE and subsequently dissolves. To answer this question, sample no. 3 was subjected to the same oxidation treatment as sample no. 2b but in this case the sample was pulled out of solution at a higher potential (+0.4 V vs. NHE) where MoO₃ is thermodynamically stable according to the Pourbaix diagrams. XPS reveals a shift of the Mo 3d and Mo 3p towards higher binding energies just as expected for MoO₃. Thus it is unlikely that MoS₂ dissolves at anodic potentials. We note that the Mo peaks of the MoO₃ were significantly greater than the Mo peaks observed on the other samples and at the same time the C 1s peak was significantly smaller. The increase in intensity could be due to a higher loading on this specific sample but we only found a factor of 2 larger oxidation peak on sample 3 than on sample 2a/b. This leads us...
to believe that there could be surface enrichment of Mo species on the outermost exposed surface the Toray paper after repeated dissolution-redeposition cycles during each potential sweep.

![Fig. 4](image-url)  
**Fig. 4** The charge of the irreversible oxidation peak as a function of the amount of Mo used during the synthesis of MoS₂.

Having established that the MoS₂ is in fact being oxidized at high anodic potentials we will now elaborate on possible reaction mechanisms. The reaction mechanism will enable us to use the irreversible oxidation peaks to determine the amount of MoS₂ present on the surface. A plot of the correlation between the amount of Mo used during synthesis and the charge of the irreversible oxidation peak is shown on figure 4. In a corrosion study by Jaegermann, bulk MoS₂ was electrochemically oxidized in KNO₃ and examined by XPS. A shift toward higher binding energies was observed for the S 2p and Mo 3d peaks and a broadening was observed in the S 2p line. This was interpreted as MoS₂ degradation to SO₄²⁻, S₂⁻ and MoO₂. We can with this knowledge consider how many electrons we expect to use to oxidize one Mo atom. If we consider one extreme where the carbon supported MoS₂ is decomposed into MoO₃ and SO₄²⁻ the following reactions would take place, where 18 electrons are transferred per Mo atom

\[
\text{MoS₂} + 11 \text{H₂O} \rightarrow \text{MoO₃} + 2 \text{SO₄}²⁻ + 22 \text{H}⁺ + 18\text{e}⁻ \quad (1)
\]

The other extreme would be that MoS₂ is decomposed into MoO₃ and S₂⁻ where 4 electrons are needed

\[
\text{MoS₂} + 3 \text{H₂O} \rightarrow \text{MoO₃} + \text{S₂}⁻ + 6 \text{H}⁺ + 4\text{e}⁻ \quad (2)
\]

According to figure 4 the correlation between the oxidation peak and the deposited amount of Mo yields 8.9 (r² = 0.55) electrons per Mo atom used in the deposition. This number is in between the two extremes mentioned above. Revisiting the XPS data we can not see whether we have produced excess SO₄²⁻ due to the background of H₂SO₄. We are however also not seeing any significant amounts S₂⁻ after cyclic voltammetry. While the samples have been subject to a high anodic (1.4 vs. NHE) potential where S₂⁻ can be oxidized to SO₄²⁻ the subsequent high cathodic (-0.4 vs. NHE) potential can reduce the S₂⁻ to H₂S.

\[
\text{MoS₂} + 7 \text{H₂O} \rightarrow \text{MoO₃} + \text{SO₄}²⁻ + \frac{1}{2} \text{S₂}⁻ + 14 \text{H}⁺ + 10\text{e}⁻ \quad (3)
\]

**HER activity of MoS₂/C**

In order to determine the activity of the MoS₂/C system per active site, we start with the irreversible oxidation to estimate the total surface area of MoS₂ on the Toray paper. The irreversible oxidation peak of the sample shown in Figure 1 and Figure 2 has a charge of 0.014 C. If we assume that 10 electrons are involved in the oxidation of MoS₂, as presented in the previous section, the surface area will be 4.2 cm² of single layered MoS₂ giving an exchange current density(i₀) of 1.1*10⁶ A/cm² (and a Tafel slope of 120mV/dec). We have previously shown that the active sites of Au(111) supported MoS₂ nanoparticles are situated on the edge (i₀ = 7.9*10⁶ A/cm²).

The exchange current density on a per active site basis will clearly be higher than the exchange current densities reported above since few of the MoS₂ sites are on the edge. Thus the values above constitute a lower bound for activity. If we incorporate the fact that the MoS₂ nanoparticles are triangular with an edge length of 25 nm approx. 8 % of the atoms will be situated at the edge of the particle. This would lead to a 12-fold increase in exchange current density per active site.

**Electrochemical characterization of WS₂**

WS₂ exhibits a layered structure similar to MoS₂, forming the same triangular shape as MoS₂ when prepared under similar conditions. WS₂ supported on SiO₂ has previously been proposed as a catalyst for the hydrogen evolution reaction.

![Diagram](image-url)
Fig. 5 A: Tafel and polarization curve (inset) of WS$_2$/C, scan rate 5 mV/s both the initial and the final stable scan is shown. B: CV of WS$_2$/C showing the deactivation of WS$_2$.

The WS$_2$ was studied on Toray paper and the preparation method was similar to that of the MoS$_2$ samples (See experimental details). In Figure 5 the results of the electrochemical measurements are shown. Figure 5 A) shows a Tafel plot ($\log i$ vs $E$) and a polarization curve ($i$ vs $E$) in the region where we have previously observed hydrogen evolution on MoS$_2$ samples. On the polarization curve (inset of Figure 5 A)) the cathodic current increases at potentials more negative than -0.2 V vs. NHE ascribed to HER activity. The Tafel slope on this sample is found to be 135 mV/dec. indicating that the current could be transport limited. Figure 5 b) shows the deactivation of the sample at positive potentials. As with the MoS$_2$/C sample, an oxidation feature is observed with a peak at approx. 1 V vs. NHE, and on consecutive sweeps the peak disappears concurrent with a drop in the HER current. This is the same behavior as we have seen on the MoS$_2$/C sample except that the WS$_2$ sample required two sweeps towards highly anodic potentials before the HER current was affected. This behavior was also observed with high loadings of MoS$_2$/C samples that surely had formed multilayer’s. In this case, the outer layer could be passivated by a sulfur/oxide layer, thus requiring several oxidation/reduction steps to completely dissolve the metal sulfide. Apart from the potential formation of multilayer’s the oxidation of the WS$_2$ is similar to that of MoS$_2$/C and it is assumed that the oxidation process is similar to that of MoS$_2$/C.

Cobalt promoted MoS$_2$ and WS$_2$

Cobalt is often used to promote WS$_2$ and MoS$_2$ in catalyzing the hydrodesulfurization reaction. Both the structural and the catalytic effect of adding cobalt has been extensively studied$^2$. It is widely accepted that the cobalt is located at the edge of MoS$_2$ more specifically the so called S-edge(-1010). Cobalt promotion of MoS$_2$ has also been shown to change the morphology significantly$^{11}$. Cobalt promoted MoS$_2$ is usually found as truncated triangles exposing the S-edge(1010) predominantly, unlike the unpromoted MoS$_2$, in which the triangles are less truncated and primarily expose their Mo-edge(10-10)$^{23}$. In the following we will show data for sulfided Co and Co promoted WS$_2$ and MoS$_2$.

Electrochemical characterization of cobalt sulfide CoS$_x$.

The first step in testing the promotion by cobalt is the test of sulfided cobalt itself. We have used Co(acetate) as the Co precursor as described in$^{11}$. The precursor was sulfided under the same conditions as the MoS$_2$ and the WS$_2$ samples (see experimental details). The Co is expected to be in the form of CoS$_9$ immediately after sulfidation, but as this form is not stable in air$^{24}$ our Co sulfide is most likely partially sulfided (CoS$_x$).

Fig. 6 A: Tafel and polarization curve (inset) of CoS$_x$/C, both the initial and the final stable scan is shown. B: CV of CoS$_x$/C showing the oxidation/deactivation of CoS$_x$.

Figure 6 A) shows the initial and the stable Tafel ($\log i$ vs $E$) and polarization curves ($i$ vs $E$) within a narrow potential window (maximum +0.1 V vs. NHE). Initially the activity is high, but unlike MoS$_2$ and WS$_2$ subsequent sweeps within this potential window show a significant decrease in activity. The decrease is most likely due to the CoS$_x$ instability in acidic conditions, introducing ambiguity into the interpretation of the current at cathodic potentials as the HER competes with cathodic desorption or dissolution of CoS$_x$. In Figure 6 B) a wide sweep is exhibited. The CoS$_x$ exhibits similar oxidation features as we have seen on the MoS$_2$ and WS$_2$, but in this case the variation peak is shifted towards a higher potential(1.14 V vs. NHE). After oxidation the HER activity drops just as with MoS$_2$ and WS$_2$, again indicating oxidation of the material.

CoS$_x$/C, Co-Mo-S and Co-W-S

CoS$_x$/C, Co-Mo-S and Co-W-S were prepared by co impregnation of the Mo/W and the Co precursor (see experimental details). Figure 7 A) and C) shows the Tafel (log $i$ vs $E$) and the polarization (inset) curve ($i$ vs $E$) within a narrow potential window (maximum +0.1 V vs. NHE). The
HER current diminishes just as on the pure CoS$_x$ sample: it is initially high and after subsequent sweeps the current decreases noticeably, but unlike the case of pure CoS$_x$, remains stable at a fairly high level. This indicates that some of the Co promoter is in the state of CoS$_x$, but as the current stabilizes at a higher level than pure CoS$_x$, MoS$_2$ or WS$_2$ the remaining Co must have a promotion effect. The Tafel slopes are also in the expected region (Co-W-S 132 mV/dec., Co-Mo-S 101 mV/dec.). In Figure 7 B) and D) the CVs within a wide potential window are shown and just as on the other metal sulfides we observe an irreversible oxidation peak followed by a significant decrease in HER activity. The peak maximum, however, seems to be shifted to a more anodic potential (approx. 1.1 V vs. NHE) than those corresponding to the unpromoted MoS$_2$ and WS$_2$.

**DFT calculations on WS$_2$, MoS$_2$, Co-Mo-S and Co-W-S**

We have calculated $\Delta G_H$ at the S edge (1̅01̅0) and the Mo/W edge (1̅01̅0) of WS$_2$ and MoS$_2$ and on the Co promoted S edge (1̅01̅0) edge of WS$_2$ and MoS$_2$ over a wide range of S coverage and H coverage. The choice of the relevant edge configurations have been based on the chemical potential of hydrogen and sulfur at the experimental sulfiding conditions using a thermodynamic model similar to the one presented in$^{25}$ . The structure and the differential free energies of H adsorption for these structures can be seen in Figure 8. The results indicate that non promoted WS$_2$ and MoS$_2$ nanoparticles should be reasonably good hydrogen evolution catalysts since both edges on both systems have free energies of adsorption close to zero. Hydrogen evolution on MoS$_2$ is expected to take place predominantly at the Mo edge ($\Delta G_H = 0.08$ eV) rather than the S edge ($\Delta G_H = 0.18$ eV), while for WS$_2$ both edges are equally good ($\Delta G_H = 0.22$ eV). Given these values for $\Delta G_H$, non-promoted MoS$_2$ is predicted to be a better hydrogen evolution catalyst than WS$_2$.

The incorporation of cobalt into the edge structures of both WS$_2$ and MoS$_2$ is expected to have a promotion effect. The cobalt only incorporates itself into the S edge of both cases, so $\Delta G_H$ values at the Mo/W edge remain unaffected. At the S edge, however, $\Delta G_H$ is reduced to 0.10 eV and 0.07 eV for MoS$_2$ and WS$_2$, respectively (down from 0.18 eV and 0.22 eV). We note that the free energy of hydrogen adsorption at the cobalt-promoted S edge of MoS$_2$ is very similar to the free energy of hydrogen adsorption on the Mo edge of MoS$_2$. Therefore, for MoS$_2$ the effect of promotion is the increase in the number of sites with high activity. On WS$_2$ the effect of cobalt promotion is the creation of new sites with higher activity than that prior to promotion.

In comparing all catalyst systems, DFT calculations suggest that cobalt-promoted MoS$_2$ (Co-Mo-S) should be a better catalyst than Co promoted WS$_x$ (Co-W-S) because it would have active sites on both edges and therefore a higher total number of active sites.

**Linking catalyst structure and composition to HER activity**

Calculated DFT values are best compared to experimental data where the activity has been normalized with respect to the number of active sites on the catalyst, in this case the number and type of edge sites on the different metal sulfides. We accomplish this normalization by using the irreversible oxidation features of each sulfide.

Figure 9 exhibits normalized polarization curves (E- i) pertaining to each of the different samples. There is an apparent promotion effect of Co on both the MoS$_2$ and the WS$_2$ samples. The promotion effect on the WS$_2$ sample can be explained by the DFT calculations predicting that the Co promotion should decrease the free energy of hydrogen adsorption from 0.22 eV to 0.07 eV on the S-edge and thus effectively increase the activity of the active site. MoS$_2$ is a slightly different case. It has previously been found that MoS$_2$’s Mo-edge, which has a $\Delta G_H$ of 0.08 eV, is the major edge exposed, and that this edge does not adsorb cobalt$^{23, 26}$. However, the inhomogeneous nature of these nanoparticulate catalysts suggests that both the Mo-edge and the S-edge will be present in significant fractions. Thus, the cobalt on MoS$_2$’s S-edge promotes the HER as its free energy of hydrogen adsorption is decreased from 0.18 eV to 0.10 eV. In other words, the number of active sites is increased since the normally less active S-edges becomes more active in the
presence of cobalt.

**Experimental/Calculation details**

Toray carbon paper was used as support material because it is inert, of high purity, has high conductivity and because it has adsorption sites/defects that will anchor the metal sulfide particles. The Toray paper was cut into strips that were 1 cm wide and 5 cm long. The Toray paper was loaded with catalyst by wetness impregnation with an aqueous solution of (NH₄)₂MoO₄·4H₂O in the case of MoS₂ and an aqueous solution of H₂N₂O₄·4H₂O in the case of WS₂. In the case of the sulfided Co, C₆H₄CoO₄·4H₂O in an aqueous solution was used. The promoted WS₂ and MoS₂ were made by co-impregnation of Co and Mo/W. The impregnation of the pure sulfides was done by dropping a 25 µL aliquot (0.3-1 mM for Mo, 0.8 mM for W, 4mM for Co). The co-impregnation of the promoted sulfides was done by adding a 25 µL aliquot of Mo(0.7mM) or W(0.8 mM) solution followed by a 25 µL aliquot of Co(4mM) solution. A different sample preparation was used for the MoS₂ sample for XPS analysis where a the Toray paper was dip coated in the Mo solution(0.14 M) to obtain a more uniform impregnation.

The samples were dried at 140 °C and afterwards sulfided in a tube furnace under 10% H₂S in H₂ at 450 °C for 4 hours. The samples were cooled down in the same gas stream. The electrochemical measurements were performed in N₂ purged 0.5 M H₂SO₄ (pH 0.4). To avoid contamination from the SCE reference electrode a salt bridge was used. A Pt mesh was used as the counter electrode.

The XPS data was recorded using a Perkin-Elmer surface analysis system (Physical Electronics Industries Inc., USA) with a chamber base pressure of 10⁻¹⁰ Torr. Al-Kα radiation (1486.6 eV) was used for excitation. The XPS scans on Figure 4 were measured with a pass energy of 100 eV, a step size of 1 eV, and 250 ms/step.

**DFT calculations**

An infinite stripe model, which has previously been proven successful to investigate MoS₂ based systems, is used to investigate the edges of MoS₂. The infinite stripe exposes both the {1010} Mo edge and the {1010} S edge. The supercell has 4 Mo atoms in the x-direction and 4 Mo atoms in the y-direction, in order to allow for important reconstructions with a period of 2 in the x direction and to allow decoupling of the Mo edge and the S edge in the y-direction. The stripes are separated by 14.8 Å in the z-direction and 9 Å in the y-direction.

The plane wave density functional theory code DACAP0 is used to perform the DFT calculations. The Brillouin zone is sampled using a Monkhorst-Pack k-point set containing 4 k-points in the x-direction and 1 k-point in the y- and z-direction. The calculated equilibrium lattice constant is 3.235 Å and 3.214 Å for MoS₂ and WS₂ respectively. A plane-wave cutoff of 30 Rydberg and a density wave cutoff of 45 Rydberg are employed using the double-grid technique. Ultrasoft pseudopotentials are used except for sulfur, where a soft pseudopotential is employed. A Fermi temperature of k_bT=0.1eV is used for all calculations and energies are extrapolated to zero electronic temperature. The exchange correlation functional RPBE is used. The convergence criterion for the atomic relaxation is that the norm of the total force should be smaller than 0.15eV/Å, which corresponds approximately to a max force on one atom below 0.05eV/Å . Figures of atomic structures have been made using VMD. The differential free energies are calculated as described in where 0.29eV is added to the pure DFT energy of adsorption in order to take zero point energy and entropy into account.

**Conclusions**

We have studied the hydrogen evolution on Co promoted and unpromoted nanoparticulate MoS₂ and WS₂ structures. Cyclic voltammetry revealed that they are irreversibly oxidized at high anodic potentials. We have used the irreversible oxidation features to determine the surface area of MoS₂ and proposed a possible oxidation mechanism of MoS₂. XPS analysis showed no change in the oxidation state of MoS₂ after HER measurements, but after oxidation at potentials above 0.6 V vs. NHE MoS₂ was oxidized. We found that the activity of the carbon supported MoS₂ is comparable to that of our previously published results on Au(111) supported MoS₂. WS₂ has a similar structure and was also investigated in this study. It was found to irreversibly oxidize at high anodic potentials, just like MoS₂ and was found to be almost as active. Tests of Cobalt promoted MoS₂ and WS₂ samples were also performed and Co is indeed promoting the HER in both cases. The findings are corroborated by DFT calculations showing that the activity of the different samples should be WS₂<Co-MoS<Co-Mo-S<Co-W-S.

**Acknowledgments**

J.B. acknowledges support from the Danish Strategic Research Council. T.F.J. acknowledge H.C. Ørsted Postdoctoral Fellowships from the Technical University of Denmark The Center for Atomic-scale Materials Design is supported by the Lundbeck Foundation. We thank the Danish Center for Scientific Computing for computer time. The Center for Individual Nanoparticle Functionality is supported by the Danish National Research Foundation.

**Notes and references**

31 J.B. acknowledges support from the Danish Strategic Research Council. T.F.J. acknowledge H.C. Ørsted Postdoctoral Fellowships from the Technical University of Denmark The Center for Atomic-scale Materials Design is supported by the Lundbeck Foundation. We thank the Danish Center for Scientific Computing for computer time. The Center for Individual Nanoparticle Functionality is supported by the Danish National Research Foundation.

8 Journal Name, [year], [vol], 00–00 This journal is © The Royal Society of Chemistry [year]
Paper 5:

“Hydrogen Evolution on supported incomplete cubane-type $[Mo_3S_4]^{4+}$ electrocatalyst”
T. F. Jaramillo, J. Bonde, J. Zang, B. L. Ooi, J. Ulstrup and I. Chorkendorff,
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Hydrogen Evolution on supported incomplete cubane-type $[\text{Mo}_3\text{S}_4]^{4+}$ electrocatalysts

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Abstract:

Electrocatalytic properties of biomimetic supported incomplete cubane-type [Mo₃S₄]⁴⁺ clusters are investigated. The activity towards the hydrogen evolution reaction (HER) is evaluated on both a high surface area gas diffusion electrode in a membrane electrode assembly and on highly orientated pyrolytic graphite (HOPG) supports. Submonolayers of the clusters were imaged by means of scanning tunnelling microscopy (STM) prior to electrochemical characterization. This enabled the quantification of the activity on a per cluster basis for the HER and the comparison of the activity with other HER catalysts. We find that the HER activity of the [Mo₃S₄]⁴⁺ is comparable with that of the edge sites of MoS₂. On the basis of these studies, supported [Mo₃S₄]⁴⁺-molecules were investigated by X-ray Photoelectron Spectroscopy (XPS) and the observed deterioration in electrocatalytic activity with time was assigned to slow [Mo₃S₄]⁴⁺ cathodic desorption from the catalyst support.

Keywords: Hydrogen evolution, Electrocatalysis, [Mo₃S₄]⁴⁺, XPS, STM, Biomimetics.

1. Introduction.

Interfacial electrocatalytic charge transfer processes are at the heart of future energy conversion. Decades of research into conventional metals and alloys has paved the way for the development of the electrocatalytic materials used in today’s commercial devices including fuel cells and electrolyzers¹. A number of challenges, however, remain in developing active, stable, and low-cost catalysts. Sulfur-rich transition metal clusters have attracted much attention due to their inherently interesting chemistry and by their structural and functional relevance to redox metalloproteins and metalloenzymes²⁻¹¹. The latter observations offer particular perspectives for the potential use of such clusters in heterogeneous catalysis and interfacial electrocatalysis.

Monolayers of a variety of redox metalloenzymes in protein film voltammetry can be brought to retain electrocatalytic electron transfer and redox enzyme function in the immobilized state on, usually modified, electrode surfaces. These enzymes include iron-sulfur cluster-based enzymes as well as Mo-
enzymes in processes such as fumarate reduction and succinate oxidation, nitrate reduction, DMSO reduction and particularly, hydrogen evolution and proton reduction\textsuperscript{12-14}. Although of great importance, redox metalloenzyme function is, however, unlikely to carry over as large-scale industrial electrocatalysts. This is due to their high cost and fragile nature in non-native surroundings outside their natural biological environments. Even though the individual enzyme molecule may be catalytically highly active, the large size of the molecules also means that the number of enzyme molecular sites on the surface in a full monolayer is much smaller than the number of active sites on metallic catalytic surfaces.

Biomimetic electrocatalysts based on chemically synthesized low-cost, earth-abundant metal-containing groups that resemble the active core groups in redox metalloenzymes may, on the other hand be able to retain electrocatalytic activity at a small size. The heme group as the core group in a broad variety of heme proteins is an example of such a robust and catalytically specific molecular electrocatalyst\textsuperscript{15-19} Synthetic iron-sulfur cubane-like structures resembling the core metal centres in iron-sulfur proteins and metalloenzymes have also been studied extensively by Holm and associates\textsuperscript{2-5}. These are, however, fragile and difficult to handle, and they have not been addressed with a view on interfacial electrocatalysis.

Polynuclear sulfide-containing Mo-complexes offer instead a class of robust and relatively easier-to-handle homologues to the iron-sulfur cubane-type clusters The trimeric, incomplete cubane-type cluster aqua ion, \([\text{Mo}_3\text{S}_4(\text{H}_2\text{O})_6]^{4+}\) (\([\text{Mo}_3\text{S}_4]^{4+}\)), in particular has been extensively investigated\textsuperscript{5}, see Figure 1. \([\text{Mo}_3\text{S}_4]^{4+}\) is extremely stable, but electrochemically reactive, and when coordinated to ligands such as iminodiacetate, evidence for three different oxidation states has been obtained\textsuperscript{6}. This suggests that its electrocatalytic activity may be promising. \([\text{Mo}_3\text{S}_4]^{4+}\) reacts readily with a range of metals or metal ions to form heterometallic complete cubane-type clusters, \([\text{Mo}_3\text{MS}_4]^n+\) (\(M = \text{Pd, Pt, Ru, Ni, Fe, etc.}\)), thus offering possibilities for tuning of the redox potential and other catalytically important properties. Several of these heterometallic cubane-type \([\text{Mo}_3\text{MS}_4]^n+\) clusters appeared to be catalytically active, for example, in hydrazine N-N cleavage reactions\textsuperscript{7}, lactonization of alkynoic acids\textsuperscript{8}, hydrodesulfurization\textsuperscript{9},
and other organic chemical reactions\textsuperscript{10}. The recently reported spontaneous adsorption of $[\text{Mo}_3\text{S}_4]^{4+}$ on Au(111) surfaces resulting in $[\text{Mo}_3\text{S}_4]^{4+}$-monolayer formation\textsuperscript{11} is also noteworthy. It is the first example of monolayer formation \textit{in situ} by a polynuclear metal-sulfide cluster directly anchored onto the gold surface and structurally mapped to single-molecule resolution by scanning tunnelling microscopy directly in aqueous electrolyte solution (\textit{in situ} STM). Coupled with the different accessible redox-states for polynuclear metal-sulfide clusters, this has sparked off the idea of exploiting the $[\text{Mo}_3\text{S}_4]^{4+}$ for functionalization of other materials.

We report here comprehensive data for supported incomplete cubane-type $[\text{Mo}_3\text{S}_4]^{4+}$ as an intriguing new class of materials that holds promise for efficient electrocatalysis of one of the altogether most important electrochemical processes, the electrochemical hydrogen evolution reaction (HER). Flexibility in cluster synthesis, composition, and structure may allow for new electrode materials to be tailored based on these relatively low-cost, earth abundant elements. We provide voltammetric current-overpotential relations on carbon-supported (high surface area carbon and highly oriented pyrolytic graphite, HOPG) multi- and monolayers of the $[\text{Mo}_3\text{S}_4]^{4+}$ molecules. STM of the cluster molecules on activated HOPG that “resemble” the high surface area carbon, resolved to the level of the single molecule, enables further direct comparison between the macroscopic electrochemical exchange currents and the density of molecules on the surface. Catalytic activity per molecule can be estimated in this way. Comprehensive electronic characterization of different states of the $[\text{Mo}_3\text{S}_4]^{4+}$ molecules using X-ray Photo-electron Spectroscopy (XPS) has finally been achieved.

Models of catalysis that allow us to understand catalytic reactions at a molecular level have emerged in recent years\textsuperscript{20, 21}. By carefully examining the active site of working systems, we can identify the material (or site) properties that principally control electrocatalytic activity and then search for other, less expensive, systems that emulate these properties. As noted, the HER is an important interfacial electrochemical process that could have a crucial role in a future hydrogen economy\textsuperscript{22}. In 1958, Parsons\textsuperscript{23} first proposed the descriptor for facile hydrogen evolution, i.e. thermo-neutral H-adsorption from electrochemical reduction of the solute proton. Optimal catalytic surfaces are those which exhibit
a Gibbs free energy of electrochemical H-adsorption ($\Delta G_{\text{H}}$) close to zero, a notion which has recently gained much more attention. Recent DFT calculations have shown that there is indeed a correlation between calculated values of $\Delta G_{\text{H}}$ and experimentally measured exchange current densities on a wide range of materials\textsuperscript{24-27}.

In approaching catalytic materials discovery, we have previously aimed to emulate enzymatic catalysis using solid-state materials\textsuperscript{27, 28}. The edge sites of nanoparticulate MoS$_2$ mimics, in many ways the metallic active sites in the nitrogenase and hydrogenase enzymes, both of which evolve hydrogen from protons in a facile manner without the use of precious metals\textsuperscript{27, 28}. Structurally speaking, the primary commonality between the enzymatic active center and the MoS$_2$ edge is under-coordinated sulfur, which in both cases was shown to be the active site for the HER\textsuperscript{27} as determined by DFT calculations\textsuperscript{28}. The [Mo$_3$S$_4$]$^{4+}$ studied herein also exhibits this structural commonality, inspiring further study of electrocatalytic reactions, and the HER in particular.

The incomplete cubane-type [Mo$_3$S$_4$]$^{4+}$ was first synthesized nearly two decades ago\textsuperscript{6}, but studies of electrocatalytic activity have not been reported before. In this report we present data for the electrocatalytic activity of [Mo$_3$S$_4$]$^{4+}$ on a high surface area carbon support (Vulcan xc72) followed by a more detailed study of the adsorption and electro-catalytic activity of individually separated [Mo$_3$S$_4$]$^{4+}$ molecules supported onto HOPG surfaces. By combining electrochemistry with surface science techniques, we have investigated and assessed this new class of bio-mimetic, solid-state electrocatalysts.

2. Experimental Methods


[Mo$_3$S$_4$]$^{4+}$ was synthesized by reduction of either ammonium tetrathiomolybdate or the Mo(V)$_2$ aqua ion, [Mo$_2$O$_2$S$_2$(H$_2$O)$_6$]$^{2+}$, with sodium borohydride, as reported in the literature\textsuperscript{29, 30}. Solutions of [Mo$_3$S$_4$]$^{4+}$ in H$_2$SO$_4$ were obtained by repeated cation-exchange chromatography (Dowex 50W-X2, 400 mesh) and checked by UV-Vis spectroscopy (Hewlett Packard 8453).
2.2. [Mo₃S₄]⁴⁺-electrode preparation

Phosphate buffer (PB) solutions were prepared from Suprapur® chemicals, di-potassium hydrogen phosphate (K₂HPO₄, 99.99%) and potassium dihydrogen phosphate (KH₂PO₄, 99.995%). The [Mo₃S₄]⁴⁺/HOPG samples were prepared in two different ways. One is to directly cast a drop (50 µl) of [Mo₃S₄]⁴⁺ solution (1.8 mM [Mo₃S₄]⁴⁺ in 0.4 M H₂SO₄) onto a freshly cleaved HOPG surface. Physical adsorption is the main interaction between the adsorbate and HOPG surface in this procedure. The drop was allowed to dry in a desiccator and the sample was used as such after a single rinse with a drop of water. This simple method works efficiently with hydrophobic molecules for which hydrophobic interactions are the driving force in stable adlayer formation. However, a stable adlayer for hydrophilic molecules cannot easily be maintained in this way. The HOPG surface must first be activated into a hydrophilic form compatible with a hydrophilic adsorbate such as [Mo₃S₄]⁴⁺. In the second method, the HOPG was thus activated prior to drop-casting. The activation of HOPG was carried out in pH 7.0 PB solution by the electrochemical method in reference 31. Immediately after activation and rinsing, a freshly prepared 50 µL aliquot of 1.8 mM [Mo₃S₄]⁴⁺ in 0.4 M H₂SO₄ (aq.) was drop-casted onto the activated HOPG surface and soaked for 2 hours. Excess solution was washed away by thorough rinsing with Millipore water (18.2 MΩ-cm).

2.3 Electrochemical measurements in a three-electrode cell

Cyclic voltammograms of the HER in de-aerated 0.5 M H₂SO₄ (pH 0.40) were conducted. An O-ring seal was used to expose only the STM-imaged part of the sample to the electrolyte. The cyclic voltammogram sweep rate was 5 mV/s. A SCE and a Pt mesh were used as reference and counter electrode, respectively.

2.4. Electrochemical measurements in a membrane electrode assembly (MEA)

2.4.1. MEA preparation
1 mL [Mo$_3$S$_4$]$^{4+}$ solution (40mM [Mo$_3$S$_4$]$^{4+}$ in 4M H$_2$SO$_4$) was wet-impregnated on 160 mg of Carbot xc72 carbon black and dried. An ink was then prepared with 44 mg of the carbon supported [Mo$_3$S$_4$]$^{4+}$, 0.1 mL wt % Nafion® solution (5 % wt. in a mixture of lower aliphatic alcohols and water) and 0.2 mL methanol (99.99 %). The ink was sonicated for 40 min, drop-cast on a 3.14 cm$^2$ graphite paper disk (Toray) and dried to create a gas diffusion electrode (GDE) with [Mo$_3$S$_4$]$^{4+}$.

The GDE with [Mo$_3$S$_4$]$^{4+}$ was HOT pressed (130 °C and 32 bar for 5 min.) together with a 117 Nafion® membrane and a commercial Pt GDE (IRD Fuel Cell Technology), creating an MEA with a Pt GDE on one side and a [Mo$_3$S$_4$]$^{4+}$ GDE on the other side.

2.4.2. Electrochemical measurements of the MEA

The MEA was tested in a fuel cell test setup, where pre-humidified H$_2$ (99.9999 %) was passed over the GDE’s. All measurements were performed at room temperature, with a gas flow of 20 mL/min over both electrodes and at a scan rate of 5 mV/s. The Pt GDE was used as both reference (effectively a NHE electrode) and counter electrode.

2.5. Scanning Tunneling Microscopy (STM)

STM was performed using a Pico SPM instrument (Molecular Imaging Co., USA). Pt/Ir tips (80/20, Ø 0.25 mm) were prepared by electrochemical etching$^{11}$. The sharpness of the tips was checked by recording HOPG images with atomic resolution. [Mo$_3$S$_4$]$^{4+}$/HOPG samples were characterized to single-molecule resolution by STM in air.

2.6. X-ray Photoelectron Spectroscopy (XPS)

XPS was performed using a Perkin-Elmer surface analysis system (Physical Electronics Industries Inc., USA). Al-K radiation (1486.6 eV) was used for excitation. High-resolution scans were recorded with a pass energy of 25 eV, a step size of 0.2 eV, and 500 ms/step. Base pressure of the chamber was $10^{-10}$ Torr. Spectra were calibrated to C 1$s$ at 284.5 eV.
3. Results and Discussion

3.1. Electrochemical characterization of \([\text{Mo}_3\text{S}_4]^{4+}\) in a fuel cell setup

3.1.1. Electrochemical fuel cell behaviour of carbon-supported \([\text{Mo}_3\text{S}_4]^{4+}\) multi-layers

The \([\text{Mo}_3\text{S}_4]^{4+}\) were initially deposited on high surface area carbon (xc72), which allowed us to test
the electrocatalytic activity in a Membrane Electrode Assembly (MEA). The xc72 is a highly defected
high surface area carbon and was chosen to anchor the \([\text{Mo}_3\text{S}_4]^{4+}\) to defects, eventually creating an MEA
with a high dispersion of \([\text{Mo}_3\text{S}_4]^{4+}\). The MEA was tested in a fuel cell test setup enabling testing for the
activity of both the HER and the HOR (hydrogen oxidation reaction). One particular concern of a metal
sulfide in an acidic environment (such as Nafion®) is its irreversible oxidation at anodic potentials.
However, \([\text{Mo}_3\text{S}_4]^{4+}\) showed no sign of irreversible oxidation at anodic potentials up to 1 V vs. NHE in
the initial electrochemical measurements, Figure 2A. This is in contrast to MoS₂ which is known to
oxidize irreversibly at these potentials\(^{32,33}\). We did, however, notice a reversible oxidation feature at
0.73 V vs. NHE that we assign to desorption and re-adsorption of the \([\text{Mo}_3\text{S}_4]^{4+}\). Cathodic desorption of
the \([\text{Mo}_3\text{S}_4]^{4+}\) has been observed previously in an aqueous solution by Kristensen et al\(^{11}\). The fact that
we observe a re-adsorption peak, located at 0.65 V vs. NHE in Figure 2A, is most likely due to the slow
diffusion of \([\text{Mo}_3\text{S}_4]^{4+}\) into the electrolyte (Nafion) where desorbed \([\text{Mo}_3\text{S}_4]^{4+}\) is maintained in close
proximity to the electrode unlike the case reported by Kristensen et al. on a sub-monolayer of \([\text{Mo}_3\text{S}_4]^{4+}\).
The desorption/re-adsorption ([Mo₃S₄]⁴⁺ mobility) scenario is furthermore consistent with our
observations that the Nafion membrane used for this study turned green, the same colour as [Mo₃S₄]⁴⁺ in
the H₂SO₄ solution. At anodic potentials we did not measure any evidence of the hydrogen oxidation
reaction, see Figure 2A. We did however find that hydrogen was evolved at low cathodic overpotentials
(ass. -150 mV). A comparison of the overpotentials is shown in Figure 2B. It is noted that the
[Mo₃S₄]⁴⁺ overpotential is similar to that of nanoparticulate MoS₂\(^{27,28}\). It should also be noted that the
current at high cathodic potentials decreases for every cathodic sweep while the overpotential remains constant. This could be caused by dissolution or agglomeration of the clusters.

Based on the similarity in overpotentials for the HER, the [Mo$_3$S$_4$]$^{4+}$ appear to have an activation barrier for the HER comparable to that of the previously studied MoS$_2$ system. [Mo$_3$S$_4$]$^{4+}$ could thus be a promising electrocatalyst for the HER. However, in order to compare the two systems more directly, studies of well defined [Mo$_3$S$_4$]$^{4+}$-samples with known quantity and dispersion are essential, as reported in the following section.

3.1.2. Electrochemical preparation and STM imaging of carbon-supported [Mo$_3$S$_4$]$^{4+}$ sub-monolayers on HOPG

The preparation of well-dispersed molecular scale catalytic units with minimal agglomeration is critical to all nano-scale materials studied. A significant challenge in the present work was the preparation of an electrode with a high loading of [Mo$_3$S$_4$]$^{4+}$-molecules while maintaining independent moieties. HOPG was chosen as the support material as (1) it is an atomically flat substrate suitable for high-resolution (i.e. atomic) mapping by STM and (2), it is electrochemically inert within the potential window under investigation. The strong two-dimensional bonding within the sheets of graphite and inherently low defect density results, however, in rather weak bonding with surface adsorbates, which are typically physisorbed through van der Waals forces only. STM imaging of drop-casted [Mo$_3$S$_4$]$^{4+}$ on freshly cleaved HOPG surface thus revealed significant agglomeration with large mounds of [Mo$_3$S$_4$]$^{4+}$ (images not shown). High [Mo$_3$S$_4$]$^{4+}$ surface mobility presumably results from weak interactions between the adsorbate and the pristine HOPG surface.

In order to prevent surface diffusion of the clusters, we activated a freshly cleaved HOPG surface electrochemically by an anodic pulse procedure developed previously$^{31}$. Surface functional groups such as –OH, –C=O, –COOH which serve to anchor the [Mo$_3$S$_4$]$^{4+}$ (by electrostatic interaction or ligand substitution) as well as to improve surface hydrophilicity could be introduced in this way.
Figure 3 shows STM images of this surface after drop-casting an aliquot of the aqueous \([\text{Mo}_3\text{S}_4]^{4+}/\text{H}_2\text{SO}_4\) solution. Only insignificant agglomeration is seen. Verifying the predominance of independently adsorbed \([\text{Mo}_3\text{S}_4]^{4+}\) (approx. 2 nm diameter which includes the \(\text{H}_2\text{O}/\text{SO}_4^{2-}\) coordination shell), was a crucial step needed to directly correlate the structure and function of these molecular entities at the level of the single \([\text{Mo}_3\text{S}_4]^{4+}\)-molecule catalyst. Furthermore, STM offered a quantitative measure of the surface coverage, which was determined to be \(1.6 \pm 0.2 \times 10^{11}\) mol \([\text{Mo}_3\text{S}_4]^{4+}\) / cm\(^2\) HOPG, corresponding to a surface density of \(1.0 \pm 0.1 \times 10^{13}\) \([\text{Mo}_3\text{S}_4]^{4+}\)-molecules / cm\(^2\).

### 3.1.3. Electrocatalysis of the HER by a \([\text{Mo}_3\text{S}_4]^{4+}\) sub-monolayer in a three-electrode cell:

**Electrocatalytic efficiency per single-molecule \([\text{Mo}_3\text{S}_4]^{4+}\)**

After STM imaging of the sub-monolayer of \([\text{Mo}_3\text{S}_4]^{4+}\) on the activated HOPG had established that the sample indeed consisted of a sub-monolayer of independent molecules, exactly the same sample was characterized electrochemically in a three-electrode cell. Figure 4 shows a series of CVs. Hydrogen is evolved in the initial sweeps at more cathodic potentials than -0.15 V vs NHE. This is the same region where the HER on the high surface area carbon supported supported \([\text{Mo}_3\text{S}_4]^{4+}\) was observed. The Tafel plot in the inset of Figure 4 yields an exchange current density of \(2.2 \times 10^{-7}\) A/cm\(^2\) geometric with a Tafel slope of 120 mV/dec.

The combination of electrochemistry and STM resolved to the level of the single molecule, enables direct conversion to the turnover frequency per molecule. We have previously exploited this combination to assess the number of redox metalloproteins on (bio)electrochemical surfaces\(^{27,28,34,35}\). The turnover frequency per \([\text{Mo}_3\text{S}_4]^{4+}\)-molecule was found to be 0.07 s\(^{-1}\). We do not know the exact number of active sites per \([\text{Mo}_3\text{S}_4]^{4+}\)-molecule, but this value is comparable to or even better than the turnover frequency of 0.02 \(^{-1}\) reported for a single site on the edge of MoS\(_2\) nanoparticles\(^{27}\). The turnover frequency is, however, smaller than for the turnover frequency per atom of the archetypical HER catalyst\(^{24}\) (Pt, Pd). In addition there are many more active atoms on the Pt metal surface than \([\text{Mo}_3\text{S}_4]^{4+}\)
even in a densely packed monolayer. The per molecule activity is, on the other hand larger compared with commonly considered non-precious metal HER catalysts\textsuperscript{24}, say Ni, Cu or W.

A reduction peak is seen around -0.12 V vs. NHE in the first sweep. This peak could be due to cathodic desorption of [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+} as previously proposed\textsuperscript{11}, but no conclusive evidence exist on the exact nature of this peak.

### 3.2 XPS data

Spectroscopic evidence of [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+} on the surface is crucial, as this ensures that catalytic activity of the [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+}-molecules is addressed and not that of a contaminant or an oxidized Mo species. XPS holds clues to such assessments. XPS data from both a bulk sample with multilayers of [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+} to verify that the [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+}-molecules have the expected electronic structure, and from sub-monolayer [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+} evaluated for catalytic activity are compared below.

#### 3.2.1. Bulk [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+}

XPS data of [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+} immobilized on “activated” HOPG offers important insight in the electronic states of the Mo- and S-atoms of the active catalyst molecule. As a reference, XPS of bulk [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+}/SO\textsubscript{4}\textsuperscript{2-} was first recorded, Figure 5. The deconvoluted Mo 3d, S 2s, and S 2p regions of the XPS spectra of this film are shown in Figures 5A and 5B. Niemantsverdriet et. al\textsuperscript{36}, reported XPS of K\textsubscript{5}[Mo\textsubscript{3}S\textsubscript{4}(CN)\textsubscript{9}]-3KCN-4H\textsubscript{2}O, which contains the same Mo\textsubscript{3}S\textsubscript{4}\textsuperscript{4+}-core, albeit with cyano ligands and potassium cations rather than aqua ligands and sulfate anions. The sulfate anion present in our system gives additional sulfur lines. At first glance these appear to distinguish our spectra from those of\textsuperscript{36}, but peak deconvolution discloses that the two [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+}-films are in fact spectroscopically similar (see Table 1).

The energy window shown in Figure 5B, containing overlapping Mo 3d and S 2s transitions, has been deconvoluted into six distinct peaks: S 2s of [Mo\textsubscript{3}S\textsubscript{4}]\textsuperscript{4+}, S 2s of SO\textsubscript{4}\textsuperscript{2-}, and two Mo-components leading to two sets of Mo 3d doublets, with a major (84\%) and a minor component (16\%, shifted 1.75 eV
positive). The S 2p lines of [Mo₃S₄]⁴⁺ and SO₄²⁻, shown in Figure 5A, are readily distinguished from one another. The lower intensity broad peak at 233.1 eV was assigned to the S 2s line in SO₄²⁻, consistent with reference 37 and verified by XPS experiments in which H₂SO₄ was drop-cast alone onto an HOPG surface (data not shown). This line, not present in the system studied by Niemantsverdriet et. al 36 coincides with the Mo 3d₃/2 line that they assign to [Mo₃S₄]⁴⁺. Deconvolution of the peaks was undertaken using Gaussian peak shapes and the following data fitting constraints: Splitting: S 2p = 1.19 eV, Mo 3d = 3.15 eV; Intensity ratios: Mo 3d₃/2 = 2/3 × Mo 3d₅/2, S 2p ([Mo₃S₄]⁴⁺)/S 2p (SO₄²⁻), and finally the Gaussian peak full width half maximum (FWHM) was fixed for the same species in the same energy window (e.g. FWHM Mo 3d₅/2 = FWHM Mo 3d₃/2). Apart from the S 2s and S 2p lines assigned to SO₄²⁻, and the Mo 3d doublet of the minor Mo-component, our data and that of reference 36 are quite similar. The Mo 3d₅/2 line at 229.70 eV indicates that molybdenum is in the formal 4+ state 32, while the S²⁻ ligands of [Mo₃S₄]⁴⁺, observed as the S 2p line at 161.76 eV, is close to the S²⁻ ligands in MoS₂. Only small differences between the spectra reported with thick [Mo₃S₄]⁴⁺ film and those in reference 36 are thus found, presumably caused by the different ligands and counter ions. We shall discuss the nature of the minor Mo-component in Section 3.2.2.

3.2.2. XPS: bulk vs. sub-monolayer [Mo₃S₄]⁴⁺

XPS of the sub-monolayer [Mo₃S₄]⁴⁺ on “activated” HOPG was recorded, the survey spectra (not shown) showed no evidence of contamination. Figure 6 shows the deconvoluted S 2p spectrum (Figure 6A) which corresponds well with that of bulk [Mo₃S₄]⁴⁺. The S 2p peak of SO₄²⁻ is of lower intensity (relative to the 2p of [Mo₃S₄]⁴⁺) than for bulk [Mo₃S₄]⁴⁺, this being due to the difference in film preparation. Excess drop-cast reagent (1.8mM [Mo₃S₄]⁴⁺ in 0.4 M H₂SO₄) was not thoroughly rinsed off in bulk [Mo₃S₄]⁴⁺-film preparation whereas this was the case for sub-monolayer [Mo₃S₄]⁴⁺. However, a more notable difference in the XPS is observed for the deconvoluted Mo 3d and S 2s regions. Compared to the bulk case, two sets of Mo 3d doublets are now easily distinguished; the major
(59%) doublet at 230.14 eV and 233.29 eV line up with the Mo 3d_{5/2} and Mo 3d_{3/2} of bulk [Mo_3S_4]^{4+}, while the minor (41%) set of peaks at 232.41 eV and 235.56 eV are shifted by 2.27 eV towards higher binding energies. The intensity of the major set of peaks relative to the minor set of peaks is significantly lower in this sample compared to the bulk sample. We find that the intensity ratio between the set Mo 3d peaks (major + minor component) and the S 2p peak at 162.44 eV, all related to [Mo_3S_4]^{4+}, is very similar to the bulk case (Figure 5). This shows that there is no loss of sulfur in the sub-monolayer [Mo_3S_4]^{4+}.

The minor, high binding energy Mo 3d-component we observe in both the sub-monolayer and multilayer sample is not likely due to oxidation of Mo ^{4+} to Mo ^{6+}(say,MoO_3)_{32,33}, or due to replacement of S^2- ligands in the cluster by O^2-, studied by Shibahara et al. for [Mo_3O_{4-n}S_n]^{4+} (n = 0-4) clusters^{38}. This is because we obtain near identical sulfur-to-Mo ratio in [Mo_3S_4]^{4+} for the sub-monolayer and multilayer sample although the ratio between major and minor Mo 3d component is drastically different.

We instead assign the minor Mo 3d component to distinct differences in local environment of Mo atoms in the adsorbed [Mo_3S_4]^{4+} structures. Such differences can arise from: (1) lower coordination of the clusters giving rise to a lower final state screening in XPS, (2) a difference in number (concentration) of SO_4^{2-} counter ions, or (3) O adsorption on the clusters. The lower concentration of the minor Mo 3d high binding energy component in the multilayer [Mo_3S_4]^{4+}-sample compared to the sub-monolayer sample is likely due to the fact that such local environment differences only are found at the outmost surface layer(s) in the multilayer case.

3.2.3. XPS: sub-monolayer [Mo_3S_4]^{4+} after cyclic voltammetry

The loss in catalytic activity of [Mo_3S_4]^{4+} on HOPG after voltammetric sweeps, was also investigated by XPS, Figure 7. The S 2p and S 2s peaks that correspond to SO_4^{2-} were still strongly visible in the XPS spectra recorded subsequent to cyclic voltammetric. The Mo 3d and S 2p peaks assigned to [Mo_3S_4]^{4+}, however, decreased significantly. Little or no Mo is left on the surface.
4. Conclusions and Perspectives

[Mo₅S₄]⁴⁺-type clusters are long known as a crucial part of polynuclear Mo-S chemistry. [Mo₅S₄]⁴⁺ is, one example, a prototype cluster in a broader class of [Mo₃MS₄]ⁿ⁺ cluster types with a variety of hetero-metal ions. Catalytic [Mo₃MS₄]⁴⁺-based activity in hydrodesulfurization, N-N cleavage of hydrazine and other organic chemical reactions are notable.

We have introduced [Mo₅S₄]⁴⁺ as a novel prototype catalyst in heterogeneous catalysis, specifically in the electrocatalytic evolution of hydrogen on different graphite supports of direct relevance to fuel cell electrode surfaces. [Mo₅S₄]⁴⁺ immobilized on electrode surfaces by electrochemical pulse pre-treatment to create molecular surface functionalities were found, firstly to act as an efficient electrocatalyst for the HER in the electrochemical fuel cell-like environment. The [Mo₅S₄]⁴⁺-based catalytic efficiency, compares well with the recently reported high electrocatalytic efficiency of identified MoS₂ edge surface sites. The observed catalysis displayed some instability on successive potential scanning but this was largely due to desorption of the catalyst molecules from the surface rather than decomposition, and can be improved on further optimization of the immobilization procedure. Secondly, the unique combination of electrochemistry and STM resolved to the level of the single molecule, could directly convert the catalytic efficiency to the single-molecule scale. In this way the electrocatalytic activity per molecule was found to be comparable with that of the edge sites of MoS₂. Through the multifarious and tunable [Mo₅S₄]⁴⁺-unit extendable to the whole class of heterometallic cluster compounds, these observations have opened novel routes to a new and potentially powerful class of interfacial electrocatalysts useable for one of the altogether most important electrochemical processes, the electrochemical HER reaction.

Our study rests on a comprehensive approach. In addition to cyclic voltammetry, chemical surface immobilization and electrochemical in situ STM, we have characterized the electronic properties of the [Mo₅S₄]⁴⁺ by XPS, used to investigate both bulk and sub-monolayers of [Mo₅S₄]⁴⁺. XPS of macroscopic [Mo₅S₄]⁴⁺/SO₄²⁻ film samples showed the expected majority of 3d⁵/₂ and 3d³/₂ spectral peak features for Mo in the formal 4+ oxidation state and a minority of 3d⁵/₂ and 3d³/₂ spectral peak features for Mo at higher binding energies as well as the 2s and 2p⁹/₂ and 2p⁷/₂ S signals. In addition corresponding
signals of the SO$_4^{2-}$ anion were observed. These data showed, interestingly that the [Mo$_3$S$_4$]$^{4+}$-core structure is similar to that for a related but different Mo-complex studied by Niemantsverdriet and associates$^{36}$.

The sub-monolayers of [Mo$_3$S$_4$]$^{4+}$ were found to differ from [Mo$_3$S$_4$]$^{4+}$ in the bulk state. The Mo 3$d$ peaks at higher binding energies were more pronounced than in the bulk sample indicating that part of the Mo-atoms in the [Mo$_3$S$_4$]$^{4+}$ are present in a different local environment. We cannot presently offer a precise rationale for this difference, but the mixed-valence nature of [Mo$_3$S$_4$]$^{4+}$ in the (strongly) adsorbed state can be a rationale for the observed electrocatalytic activity. Such “tunable” electronic properties are often prerogatives for efficient electronic transmission in chemical and biological charge transfer processes.

XPS finally substantiated that [Mo$_3$S$_4$]$^{4+}$-desorption is the primary cause of the apparent gradual catalytic deactivation on successive voltammetric scanning.

In conclusion, a new class of polynuclear Mo-S cluster-based catalyst has been found to show very high per-molecule catalytic efficiency in the electrochemical HER on carbon-based electrode surfaces. The prototype representative, i.e. the incomplete cubane-type cluster, [Mo$_3$S$_4$]$^{4+}$, has been in focus but is otherwise representative of the broader class of homologous heterometallic cubane-type clusters, [Mo$_3$MS$_4$]$^{4+}$. Once the observed electrocatalytic stability issue has been rectified, this class may offer new prospects as a broad and electronically versatile class of electrocatalysts in practical working environments.
ACKNOWLEDGMENTS

This work was supported by the National Research Foundation (CINF), the Danish Research Council through FTP and NABIIT. T. F. J. acknowledges the H.C. Ørsted Postdoctoral Fellowship from the Technical University of Denmark, K. A. acknowledges financial support from The Wenner-Gren Foundations and J. B. acknowledges support from the Danish Strategic Research Council.

Figure 1: Molecular structure of $[\text{Mo}_3\text{S}_4]^{4+}$. Blue (Mo), yellow (S) and red (O from water ligands).
Figure 2: Electrochemical characterization of vulcan xc72 supported $[\text{Mo}_3\text{S}_4]^{4+}$ in a membrane electrode assembly. The potential is measured with respect to a Pt electrode which serves as both the reference and the counter electrode. A: cyclic voltammogram in the anodic region, showing a reversible oxidation feature at ca. 0.7 V. B: Polarisation curves in the cathodic region where hydrogen is evolved. The activity of $[\text{Mo}_3\text{S}_4]^{4+}$ is compared with a blank electrode only containing xc72, a sample with Pt and a sample with MoS$_2$. Hydrogen were passed over both electrodes and the scanrate was 5 mV/s.
Figure 3: STM image of HOPG supported [Mo$_3$S$_4$]$^{4+}$. The coverage of the [Mo$_3$S$_4$]$^{4+}$ molecule (most likely including ligands) on HOPG is about 1.0 (±0.1)$\times10^{13}$ molecules/cm$^2$, corresponding to 1.6 (±0.2) x $10^{11}$ mol/cm$^2$. Each molecule is ~2.0(±0.5) nm in diameter from STM measurement, and an ideal close packed surface contains ~3.2 (±0.8)$\times10^{13}$ molecules/cm$^2$.

![Figure 3: STM image of HOPG supported [Mo$_3$S$_4$]$^{4+}$](image)

Figure 4: Electrochemical characterization of the submonolayer of [Mo$_3$S$_4$]$^{4+}$ on HOPG. Main: Polarization curve in the anodic region showing hydrogen evolution. The decrease in activity is most likely due to anodic desorption of the clusters. After a sweep to high cathodic potentials (1.2 V vs. NHE), the clusters desorb and the activity is only caused by the activated HOPG. Inset: the corresponding Tafel plot. The scanrate was 5 mV/s.

![Figure 4: Electrochemical characterization of the submonolayer of [Mo$_3$S$_4$]$^{4+}$ on HOPG](image)

Figure 7: XPS spectra of a submonolayer of \([\text{Mo}_3\text{S}_4]^{4+}\) on HOPG before (Gray) and after electrochemical measurements (Black). A: The S 2p region of \(\text{SO}_4^{2-}\) and \([\text{Mo}_3\text{S}_4]^{4+}\). B: The S 2s of \(\text{SO}_4^{2-}\) and \([\text{Mo}_3\text{S}_4]^{4+}\) and the Mo 3d region of \([\text{Mo}_3\text{S}_4]^{4+}\).

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<th>FWHM (eV)</th>
<th>Binding Energy (eV)</th>
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<td>S 2p(_{3/2})</td>
<td>2.36</td>
<td>162.44 ± 0.1</td>
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<td>161.76 ± 0.1</td>
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<tr>
<td>([\text{Mo}_3\text{S}_4]^{4+})</td>
<td>S 2s</td>
<td>1.14</td>
<td>226.68 ± 0.1</td>
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<tr>
<td>([\text{Mo}_3\text{S}_4]^{4+})</td>
<td>Mo 3d(_{5/2})</td>
<td>2.09</td>
<td>230.14 ± 0.1</td>
<td>1.46</td>
<td>229.70 ± 0.1</td>
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<td>([\text{Mo}_3\text{S}_4]^{4+})</td>
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<td>2.09</td>
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<td>1.74</td>
<td>168.71 ± 0.1</td>
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</table>

Niemantsverdriet\(^{36}\)(values read off raw spectra)
|       | S 2p$_{1/2}$ |            |               |               |               |
|-------|--------------|------------|---------------|---------------|
| SO$_4^{2-}$ | 2.36        | 169.64 ± 0.55 | 1.74          | 169.90 ± 0.1  |
| SO$_4^{2-}$ | 1.14        | 234.80 ± 0.37 | 2.13          | 232.99 ± 0.1  |
| [Mo$_3$S$_4$]$^{4+}$ | 2.09  | 232.41 ± 0.1   | 1.46          | 231.44 ± 0.1  |
| [Mo$_3$S$_4$]$^{4+}$ | 2.09  | 235.56 ± 0.1   | 1.46          | 234.59 ± 0.1  |

**Table 1:** XPS data for a sub-monolayer of [Mo$_3$S$_4$]$^{4+}$ and a “thick”, multi-layer [Mo$_3$S$_4$]$^{4+}$ film deposited on HOPG.

**References:**


