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Performance and Durability of Solid Oxide Electrolysis Cells

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Solid oxide fuel cells produced at Risø National Laboratory have been tested as electrolysis cells by applying an external voltage. Results on initial performance and durability of such reversible solid oxide cells at temperatures from 750 to 950°C and current densities from ~0.25 A/cm² to ~0.50 A/cm² are reported. The full cells have an initial area specific resistance as low as 0.27 Ωcm² for electrolysis operation at 850°C. During galvanostatic long-term electrolysis tests, the cells were observed to passivate mainly during the first ~100 h of electrolysis. Cells that have been passivated during electrolysis tests can be partly activated again by operation in fuel cell mode or even at constant electrolysis mode after several hundred hours of testing.© 2006 The Electrochemical Society. DOI: 10.1149/1.2216562

In the perspective of the increasing interest in renewable energy, hydrogen economy, and CO₂ neutral energy production, reversible solid oxide cells (SOCs) are a potentially interesting technology. Using a solid oxide electrolysis cell can be a cost effective and efficient way to produce hydrogen by high-temperature electrolysis of steam (HTES).1-4 The cells can be operated as solid oxide fuel cells (SOFCs) for electricity production and as solid oxide electrolysis cells (SOEC) to produce hydrogen by high-temperature electrolysis of steam by applying an external voltage. Potentially, such reversible SOCs can be combined with already existing energy technologies. By converting surplus energy from nuclear power plants or renewable energy sources such as wind or solar, the SOCs can optimize the efficiency of such energy technologies and play an important role in the security of supply in future hydrogen-based energy systems. Some of the first results on hydrogen production by HTES using SOCs were reported more than two decades ago,5 where Dönitz presented results from the HOTELLY project for a single cell and stack including durability tests; however the project was stopped around 1990. Since then, intensive research and development in the field of SOFCs has taken place and the efforts have resulted in optimized materials giving high performing, long-term stable cells.1,2 The research within the field of HTES using SOC can easily benefit from the results obtained within the SOFC research.

For SOCs to become interesting from a commercial point of view, a low internal resistance of the cell is important, not only at start-up but also during thousands of hours of electrolysis operation as the hydrogen production price is proportional to the resistance of the cell. So far, only a few results on durability of high-performance SOECs have been reported in literature and even though the operation of the SOCs is reversible and can have comparable initial performance in electrolysis and fuel cell mode, the degree of passivation of the cells during long-term testing in fuel cell and electrolysis operation mode, respectively, can be dramatically different.3 Therefore, it is necessary not only to produce high-performance SOECs but also long-term stable electrolysis cells.

Results on performance and durability of SOECs are presented here. Polarization curves (IV curves) at various test conditions have been recorded to monitor the initial performance for both fuel cell and electrolysis operation of the SOCs produced at Risø National Laboratory. Results from galvanostatic long-term electrolysis tests for four SOCs are given and the electrolysis testing is shown to lead to a significant passivation of the cells. A partial activation of an electrolysis tested cell by fuel cell operation is reported. Furthermore, an example is given of a 776 h electrolysis test, where the passivation of the electrolysis cell was followed by a partial activation at constant electrolysis conditions.

Experimental

Nickel-yttria-stabilized zirconia (Ni/YSZ) supported DK-SOFC cells were used for the electrolysis tests. The cells are full cells produced at Risø National Laboratory.6,7 The cells have a 10–15 μm thick hydrogen electrode of Ni/YSZ cermet, a 10–15 μm thick YSZ electrolyte, a 15–20 μm thick strontium-doped lanthanum manganite (LSM-YSZ) composite oxygen electrode, and the cells are supported by a ~300 μm thick Ni/YSZ layer. The ratio between Ni and YSZ (T8Z; Tosoh Corporation, ZrO₂ stabilized with 8 mol % Y₂O₃) is 40/60 vol % both for the support layer and the active electrode layer.8 The composition of the LSM is (La₀.₇₅Sr₀.₂₅)₀.₉₅MnO₃, and the ratio between LSM and YSZ in the composite electrode is LSM/YSZ = 50/50 vol %,3,9 As illustrated in Fig. 1 (top), the SOCs are planar 5 x 5 cm cells with an active electrode area of 16 cm². One half of the setup for cell testing is illustrated in Fig. 1 (bottom); this includes the alumina housing, current collector (Ni foil), glass sealing, and Ni/YSZ-based gas distributor. The air distributor (LSM-based), current collector (gold foil), and top part of the alumina cell housing is then placed on top to give a cross flow for the gasses. A detailed description of the setup is given elsewhere:10,11 At startup,
the nickel oxide in the Ni/YSZ electrode is reduced to nickel in hydrogen at 1000°C. Before electrolysis is started, the gas passing over the oxygen electrode is changed from air to 10 L/h of O2. Thereby p(O2) is kept constant at 1 atm when switching from open-circuit voltage (OCV) to electrolysis operation of the cell. This is advantageous for subsequent analysis of the possible changes in the oxygen electrode response observed in the electrochemical impedance spectra recorded during testing. The inlet gas composition to the hydrogen electrode was p(H2O) = 0.7 atm and p(H2) = 0.3 atm during electrolysis testing for all electrolysis tests reported in this work. This gas composition to the Ni/YSZ electrode is obtained by mixing 6 L/h O2 and 17 L/h H2 in a gas mixer and leading the mixture to the cell through the inlet tube (Fig. 1). All the gas compositions reported in this work are given as initial fuel compositions at the gas inlet of the cell. The results presented are four SOC electrolysis tests. The tests were performed at various temperatures and current densities. An overview of the electrochemical durability test conditions is given in Table I. The degree of steam conversion given in Table I is calculated as the number of converted water molecules using Faraday’s law, divided by the total number of H2O molecules led to the gas inlet of the cell. For all four tests, the test conditions were kept constant during electrolysis test and the tests were run galvanostatic.

A dc characterization has been performed by recording IV curves for each of the cells before and after the long-term electrolysis tests. The polarization curves were measured using controlled current method stepping 62.5 mA (s cm2). The IV curves presented in this work are almost linear and the area specific resistances (ASR) values calculated as the chord from OCV to the voltages measured at given current densities. Where nothing else is stated, data from OCV to p = 0.75 and −0.75 atm, respectively, are included to calculate the ASR values. The ASR values therefore reflect the cell performances over a large polarization range.

Table I. Test conditions and duration for the four electrolysis tests. For all tests, oxygen was passed over the positive electrode, and the inlet gas composition to the negative electrode was p(H2O) = 0.7 atm and p(H2) = 0.3 atm. The steam conversion is the number of converted water molecules (Faraday’s law) divided by the total number of H2O molecules led to the cell.

<table>
<thead>
<tr>
<th>Test</th>
<th>Current density</th>
<th>Steam conversion</th>
<th>Temperature</th>
<th>Electrolysis test time</th>
</tr>
</thead>
<tbody>
<tr>
<td>A</td>
<td>−0.25 A/cm2</td>
<td>14%</td>
<td>750°C</td>
<td>82 h</td>
</tr>
<tr>
<td>B</td>
<td>−0.50 A/cm2</td>
<td>28%</td>
<td>950°C</td>
<td>140 h</td>
</tr>
<tr>
<td>C</td>
<td>−0.50 A/cm2</td>
<td>28%</td>
<td>850°C</td>
<td>135 h</td>
</tr>
<tr>
<td>D</td>
<td>−0.25 A/cm2</td>
<td>14%</td>
<td>850°C</td>
<td>766 h</td>
</tr>
</tbody>
</table>

The initial performance of all cells was measured by recording IV curves at various temperatures and partial pressure of steam to the Ni/YSZ electrode. Figure 2 shows a comparison of such initial IV curves for the two cells with the highest and the lowest performance, namely, the cells used for tests B and C. The IV curves shown were recorded at 850°C and p(H2O)/p(H2) = 1 and are air to the oxygen electrode.

From the IV characteristic shown in Fig. 2, it is observed that no discontinuity occurs in the shift from fuel cell to electrolysis operation. The area specific resistances (ASR) at varying p(H2)/p(H2O) ratios at 850°C for the cells used for tests B and C is given in Table II. It is seen from the numbers in Table II that, even though the slopes of the IV curves in Fig. 2 look identical for positive and negative current densities for each of the two cells, the ASR is larger when running the cells in electrolysis mode than in fuel cell mode. Voltages at distinct current densities were also compared. Both cells had an open-circuit voltage of 928 mV at 850°C and p(H2O)/p(H2) = 1. Voltages of 821 and 1061 mV were measured at 0.50 A/cm2 and −0.50 A/cm2, respectively, for the cell with the highest performance (test B). The corresponding values for the cell having the lowest initial performance (test C) were 775 mV and 1091 mV at 0.50 A/cm2 and −0.50 A/cm2, respectively. The effect of temperature on the initial performance of the SOCs has also been investigated. Figure 3 shows an example of the effect of lowering the temperature from 850 to 750°C for the good performing cell used for test B. Both curves were recorded at p(H2O)/p(H2) = 1. There is still continuity across OCV for the IV curve at 750°C but the ASR has more than doubled compared with the ASR values at 850°C (Table II). For the IV curve at 750°C in fuel cell mode the ASR is 0.44 Ω cm2. For the IV curve at 750°C in electrolysis mode the ASR is 0.65 Ω cm2 if the data to −0.75 A/cm2 is included and the chord is used for the calculation of the ASR but

![Figure 2. Comparison of IV curves recorded at 850°C before the electrolysis testing for the cell used for test B (C) and test C (Δ), the cells with the highest and lowest performance. Air was passed over the oxygen electrode and the gas composition to the hydrogen electrode was p(H2O) = p(H2) = 0.5 atm.](image)

**Table II.** ASR for the cell with the highest (test B) and lowest (test C) performance at 850°C at varying steam content to the hydrogen electrode. ASR values calculated as the chord from OCV to the voltages measured at current densities of 0.75 A/cm2. For test C at p(H2O) = 0.5 atm, fuel cell mode, only data to 0.67 A/cm2 was available.

<table>
<thead>
<tr>
<th>p(H2O)</th>
<th>0.05 atm</th>
<th>0.20 atm</th>
<th>0.50 atm (fuel cell)</th>
<th>0.50 atm (electrolysis)</th>
</tr>
</thead>
<tbody>
<tr>
<td>ASRtest B</td>
<td>0.28 Ω cm2</td>
<td>0.22 Ω cm2</td>
<td>0.21 Ω cm2</td>
<td>0.27 Ω cm2</td>
</tr>
<tr>
<td>ASRtest C</td>
<td>0.40 Ω cm2</td>
<td>0.35 Ω cm2</td>
<td>0.32 Ω cm2</td>
<td>0.34 Ω cm2</td>
</tr>
</tbody>
</table>
this number hides the observed hysteresis effect. Calculating the ASR value as the chord from OCV to the voltage measured at a current density of $-0.50$ A/cm$^2$ for the start and end part of the IV curve leads to ASR values of 0.70 $\Omega$ cm$^2$ and 0.60 $\Omega$ cm$^2$, respectively. The last part of this IV curve represents the more stable system. The hysteresis effects for this IV curve in electrolysis mode at 750°C is of course also clear from the measured voltages at $-0.50$ A/cm$^2$. An open-circuit voltage of 957 mV was measured at 750°C. Voltages of 1261 and 1309 mV were measured at $-0.50$ A/cm$^2$ for the first and last part of the electrolysis IV curve in Fig. 3. The same four trends have been observed for all tests conducted. The resulting development of the cell voltages is shown in Fig. 4. For all tests, the cell voltage increased due to an increase in the internal resistance of the cell. The increase in cell voltage had a tendency to take the form of an “S”-curve and level off after $\sim$100 h of electrolysis or less. The least pronounced passivation over 135 h of electrolysis was observed for the high-temperature test B, which actually started out with a minor activation of the cell. The most significant passivation occurred for test A where the cell voltage increased from 1055 to 1275 mV within only 82 h of electrolysis. As the cell voltage seems to have stabilized at 1275 mV, electrolysis test A was stopped. The development of the polarization resistance monitored by EIS recorded during the pronounced passivation observed for test A is described and analyzed elsewhere. Another and a very simple way to monitor the passivation of the cell used for test A is by comparison of IV curves recorded before and immediately after the electrolysis test and these two IV curves are shown in Fig. 5. The passivation of the cell has led to an increased slope of the IV curve. Data from the IV curve were applied to calculate the conversion corrected ASRs as the internal resistance of the cell depends on test conditions such as the reactant utilization.

The over voltage will not be equal at the gas-inlet and gas-outlet and therefore a conversion correction has been made for the ASR using an iterative calculation method as discussed elsewhere. The conversion corrected ASRs are included in Fig. 5. A significant hysteresis effect is observed for the IV curve recorded immediately after the electrolysis test A. This hysteresis effect corresponds to a partial activation of the cell obtained during the recording of the IV curve in fuel cell mode after electrolysis test A. In Fig. 5, the direction of time is indicated by arrows. Qualitatively, the course of the cell voltage for test C seems to be similar to that of test A (Fig. 4) but test C was run for a longer time than test A. Figure 6 shows a

Figure 3. Initial IV curves at 850°C (□) and 750°C (△) for the cell used in test B. Arrows indicate direction of time when recording the IV curve at 750°C (△).

Figure 4. Cell voltage as a function of time at constant electrolysis conditions (Table I). Pure O$_2$ was passed over the positive electrode and the gas composition to the negative electrode was $p$(H$_2$O) = 0.7 atm and $p$(H$_2$) = 0.3 atm.

Figure 5. IV curve recorded before (△) and immediately after (□) electrolysis test A. The IV curves were recorded at 750°C and with gas composition of $p$(H$_2$) = 0.95 atm and $p$(H$_2$O) = 0.05 atm to the hydrogen electrode. Arrows indicate direction of time. Open symbols show the cell voltages and closed symbols the conversion corrected ASR values.

Figure 6. Cell voltage during electrolysis for test C (850°C, $-0.5$ A/cm$^2$) emphasizing on the last 80 h of electrolysis test period where a tendency of decreasing cell voltage corresponding to a decrease in resistance of the cell was observed.
from 0.16 to 0.46

operation of the cell and this correspond to a partial activation of
voltage increased by 49 mV during the 97 h of constant fuel cell
test period is shown in Fig. 8. The cell voltage increased from
test was run for a total electrolysis time of 766 h at constant condi-
tions, and the course of the cell voltage over the entire electrolysis
test is shown in Fig. 9. The EIS in Fig. 9 were recorded after 1, 45, 116,
and O2 was passed over the oxygen electrode. Closed symbols for Rs corre-
spond to the impedance spectra shown in Fig. 9.

Nyquist and the Bode plots show that Rs stays constant during the
entire electrolysis test D. From the imaginary part of the Bode plot in
Fig. 9, it is observed that numeric maximum for Zimag decreases from a frequency of ca. 7 kHz after 1 h of electrolysis to a frequency of ~2 kHz after 45 h of electrolysis and down to ~400 Hz at the most passivated state after 116 h of electrolysis testing. Furthermore, the imaginary part of the Bode plot shows minor increase/decrease in Zimag for the impedance at a frequency of 1–5 Hz, which is ascribed to gas conversion impedance.15

In Fig. 10, the IV curves (850°C, p(H2O) = 0.46 atm and
p(H2) = 0.53 atm, air to the oxygen electrode) in fuel cell and elec-
A1744
tricity operation before and after electrolysis test D is shown. Only
a limited increase in the ASR is observed for the IV curve after
elctrolysis test D. The difference in the slope of the IV curves is
slightly larger for electrolysis mode than for fuel cell mode. Figure
11 shows two fuel cell mode IV curves recorded just before elec-
ytrolysis test D and immediately after finishing the electrolysis opera-
tion of the cell. These IV curves are almost identical. The partial
activation of the cell that was caused by recording a fuel cell IV
curve immediately after test A (Fig. 5), was not observed for this test
D. IV curves for test D were also recorded at the same conditions as
those for test A in Fig. 5 (fuel cell operation, 750°C, p(H2O)
= 0.05 atm and p(H2) = 0.95 atm). For those IV curves, the ASR
increased from 0.54 Ω cm2 before test to 0.61 Ω cm2 after the
elctrolysis test but no noticeable activation effect was observed for the
fuel cell mode IV curve recorded after electrolysis test D.

Discussion

Initial performance of the SOECs.— The continuity of the IV
curves (Fig. 2 and 3) close to OCV verifies that even though these
cells were produced and optimized for fuel cell use, they can work as reversible SOCs. In general the initial ASR obtained from IV
curves was lower in fuel cell mode than in electrolysis mode (Table II). Table IV lists some initial performances obtained from IV curves in electrolysis mode for SOECs reported in literature. As discussed by Mogensen et al.,14 the concept of area specific resistance for

Table III. Rs, Rp and increase in Rp relative to Rs after 1 h of electrolysis for test D. Rs is taken as the real part of the impedance at 82 kHz and
Rp is taken as the real part of the impedance at 0.82 Hz minus Rs.

<table>
<thead>
<tr>
<th>Time</th>
<th>1 h</th>
<th>45 h</th>
<th>116 h</th>
<th>317 h</th>
<th>767 h</th>
</tr>
</thead>
<tbody>
<tr>
<td>Rs</td>
<td>0.131 Ω cm²</td>
<td>0.135 Ω cm²</td>
<td>0.133 Ω cm²</td>
<td>0.132 Ω cm²</td>
<td>0.126 Ω cm²</td>
</tr>
<tr>
<td>Rp</td>
<td>0.163 Ω cm²</td>
<td>0.321 Ω cm²</td>
<td>0.455 Ω cm²</td>
<td>0.322 Ω cm²</td>
<td>0.224 Ω cm²</td>
</tr>
<tr>
<td>Rp/Rs (1 h)</td>
<td>1</td>
<td>2.0</td>
<td>2.7</td>
<td>2.0</td>
<td>1.4</td>
</tr>
</tbody>
</table>
SOFCs is often used, though no generally accepted definition seems to exist. Because the ASR depends on fuel utilization, a more direct description of the cell performance is given by the conversion corrected ASRs (Fig. 5 and 11). Unfortunately, conversion corrected ASRs or information enabling the calculation of it, is not always reported in literature. Therefore, the listing of ASRs given in Table IV is simply obtained by taking the slopes of the reported IV curves in the linear regions. The references in Table IV have been selected as they represent results for cells and test conditions close to those reported in literature. Such a build-up of impurities could lead to an increase in cell voltage as observed in Fig. 4, where the cell voltage curve levels off when the impurity build-up at the Ni-YSZ TPB stops.

Previously, Jensen found evidence for a build-up of impurities containing silicon at the three-phase-boundary (TPB) for a Ni-YSZ model system and Liu found segregated silicon containing impurities in a tested half cell by scanning and transmission electron microscopy. Such a build-up of impurities could lead to an increase in cell voltage as observed in Fig. 4, where the cell voltage curve levels off when the impurity build-up at the Ni-YSZ TPB stops (all impurities collected). A subsequent decrease in the cell voltage could be due to a conversion of the impurity phase, e.g., crystallization of the glass, see below. The tendency for the course of the cell voltage for all four tests to take the form of “S-curves” supports the explanation for the passivation of the electrolysis cells given by Jensen et al. By deconvolution of the EIS recorded during electrolysis, it was shown that the rate limiting step for the steam electrolysis using SOCs was due to a process at the TPB in the Ni/YSZ electrode and it was argued that it is related to an increase in the diffusion path length during the passivation of the electrolysis cell. This phenomenon was explained as impurities building up at the TPB in the Ni/YSZ.

### Activation of an electrolysis cell

Activation of an electrolysis cell.—To the best of our knowledge, the passivation of an electrolysis cell followed by a partial...
activation at constant electrolysis conditions, as obtained for electrolysis test D, has not been reported for solid oxide electrolysis cells previously. This phenomenon has been observed for several of the electrolysis tested cells. Furthermore, the EIS recorded during this long-term electrolysis test D did not only lead to polarization resistances being equal during passivation and the subsequent activation of the cell; the EIS with the same $R_p$ recorded during passivation and activation are identical at each measured frequency (Fig. 9). This strongly suggests that it is the same processes that are the rate limiting steps both during the passivation and the following activation of the electrolysis cell. The analysis of the EIS during electrolysis for the first 116 h presents towards that the rate limiting step responsible for the passivation of the cell can be diffusion at the TPB caused by an increased diffusion path length. If the passivation of the cell used for test D is due to a build-up of glassy phase impurities at the TPB of the hydrogen electrode, then a plausible, but not yet experimentally verified, explanation for the subsequent partial activation of the cell could be a break-up of the glass caused by crystallization of these glassy phases. Such break-up would lead to a decrease in the diffusion path length and enable the complete overlap of the EIS recorded during passivation and activation of the cell as observed in Fig. 9. Impurities containing silica have been observed by scanning electron microscopy and detected by energy dispersive spectroscopy in the hydrogen electrode of the electrolysis tested cells. Further microscopy work is in progress. The partial activation of the cell by running an IV curve in fuel cell mode after the electrolysis test has not only been observed for test A but also for other tests, where a fuel cell IV curve was recorded immediately after a significant passivation in electrolysis operation. For the IV curve recorded after electrolysis test A (Fig. 5), the conversion corrected ASR decreased by 20% at a current density of $-62.5 \text{ mA/cm}^2$ and this activation occurred within the 22 min it took to record the entire IV curve. This activation of the cell is possibly of a nature different from that of the activation observed for test D. The activation due to FC operation is governed by the change in current direction, e.g., a change in the direction of $O^2-$ ions conducted in the electrolyte and a change in the $p(\text{H}_2\text{O})$ gradient in the Ni/YSZ electrode. Impurities can be removed from the TPB and transported towards the bulk of the composite Ni-YSZ electrode, i.e., away from the interface at the electrolyte where the electrochemistry happens. This fuel-cell-activation of the cell takes place on a much shorter time scale than the activation of the cell in test D, where a 51% reduction in $R_p$ was obtained during 651 h of constant electrolysis conditions. Considering the partial activation of cell A by running an IV-curve in fuel cell mode after electrolysis passivation, it is not surprising that a partial activation of a SOC after electrolysis testing can also be obtained by operating the cell at constant fuel cell conditions as observed for cell C (Fig. 7). For the activation of the cell at constant fuel cell operation, the time scale is again shorter than for the activation observed for test D at constant electrolysis conditions. For test D the $R_p$ had already dropped to half its maximum value when the long-term test was stopped after 766 h of electrolysis, and no further activation due to recording an IV curve in fuel cell mode could be obtained (Fig. 11).

**Conclusion**

From the results presented here using DK-SOFCs for high temperature electrolysis of steam it can be concluded that:

The cells produced at Risø National Laboratory can be operated both as fuel cells and electrolysis cells.

The area specific resistance obtained from the IV curves run in electrolysis mode was higher than for fuel cell IV curves for the same cells.

The IV curves show that the SOECs tested in this work performs very well compared with similar SOECs reported in literature.

At constant galvanostatic electrolysis conditions, the internal resistance of the cells increased significantly during the first ~100 h, after which the cell voltage stabilized or even decreased.

For the most passivated cell (test A, $-0.25 \text{ A/cm}^2$, 750°C, 82 h of electrolysis), a partial activation of the cell was obtained by running an IV curve in fuel cell mode immediately after electrolysis life-time test.

A cell that has been passivated during electrolysis can at least be partly activated by operating the cell at constant fuel cell conditions.

For test D ($-0.25 \text{ A/cm}^2$, 850°C, 766 h of electrolysis), a partial activation of the cell following the initial passivation was observed when operated at constant galvanostatic electrolysis conditions.

**Acknowledgment**

This work was supported by the EC via the project “Hi2H2” contract no. FP6-503765. Thanks to the Fuel Cell group and to senior scientist J. Bilde-Strengren, Materials Research Department, at Risø National Laboratory for their help and assistance.

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**References**


**Table IV. Some reported initial performances of electrolysis cells. Comparison of ASRs obtained from IV curves. The ASRs are taken as the slopes in the linear regions of the electrolysis IV curves presented in the references cited. For each reference the ASR on full cells for systems and with the experimental conditions closest to the ones applied in this work is given.**

<table>
<thead>
<tr>
<th>Ref.</th>
<th>$T$ [°C]</th>
<th>$p(\text{H}_2\text{O})$ [atm]</th>
<th>$p(\text{H}_2)$ [atm]</th>
<th>ASR [$\Omega \text{ cm}^2$]</th>
<th>Specifications</th>
</tr>
</thead>
<tbody>
<tr>
<td>This work</td>
<td>850</td>
<td>0.50</td>
<td>0.50</td>
<td>0.27</td>
<td>Ni/YSZ-YSZ-LSM (planar 2G DK-SOFC), test B</td>
</tr>
<tr>
<td>This work</td>
<td>950</td>
<td>0.50</td>
<td>0.50</td>
<td>0.15</td>
<td>Ni/YSZ-YSZ-LSM (planar 2G DK-SOFC), test B</td>
</tr>
<tr>
<td>16</td>
<td>1000</td>
<td>0.67</td>
<td>0.33</td>
<td>1.17</td>
<td>Ni/YSZ-YSZ-LSM (tubular cell)</td>
</tr>
<tr>
<td>19</td>
<td>908</td>
<td>0.67</td>
<td>0.33</td>
<td>2.7</td>
<td>Ni/YSZ-YSZ-LSM</td>
</tr>
<tr>
<td>20</td>
<td>1000</td>
<td>0.91</td>
<td>0.09</td>
<td>2</td>
<td>Ni/YSZ-YSZ-LSM</td>
</tr>
<tr>
<td>21</td>
<td>1000</td>
<td>0.50</td>
<td>0.50</td>
<td>0.7</td>
<td>Ni/YSZ-YSZ-LSM</td>
</tr>
<tr>
<td>22</td>
<td>850</td>
<td>0.50</td>
<td>0.50</td>
<td>0.45</td>
<td>Ni/YSZ-ScSZ (175 µm)-LSM</td>
</tr>
<tr>
<td>23</td>
<td>900</td>
<td>0.50</td>
<td>0.50</td>
<td>1.8</td>
<td>Figure 2 in Ref. 23 was used to estimate the ASR for the full cell.</td>
</tr>
<tr>
<td>24</td>
<td>850</td>
<td>0.11</td>
<td>0.89</td>
<td>0.35</td>
<td>Ni/YSZ-ScSZ (125 µm)-LSM</td>
</tr>
</tbody>
</table>

*a The ASR given is the average per cell in the stacks. The numbers calculated from the IV curves have been reduced by 31% to correct for the resistance due to interconnect. The reduction by 31% was estimated by comparison of single cell and stack test at the same conditions as they were reported in Ref. 16.


