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Chemically Controlled Interfacial Nanoparticle Assembly into Nanoporous Gold Films for Electrochemical Applications

Mikkil U.-B. Christiansen, Nedjeljko Seselj, Christian Engelbrekt, Michal Wagner, Frederick N. Stappen, Jingdong Zhang*

Nanoporous gold (NPG) is an effective material for electrocatalysis and can be made by a dealloying method, such as etching of silver-gold alloys. Dealloyed NPG may contain residual silver that affects its catalytic performance. We report a different approach to formation of NPG at liquid/air interface, starting from gold nanoparticles (AuNPs) in aqueous solution, thus providing silver-free gold films. Chloroauric acid is reduced to AuNPs building blocks by 2-(N-morpholino)ethanesulfonic acid, which also acts as a protecting agent and a pH buffer. We reproducibly obtain continuous gold networks by the addition of potassium chloride before AuNP synthesis. Formation and hydrochloric acid to the resultant AuNP solutions, and hydrochloric acid triggers sintering of AuNPs into a long-range continuous gold network, i.e. nano film. The film formation was controlled by manipulating the trapping of nanoparticles (NPs) with temperature, ionic strength and protonation of the buffer. The morphology and properties of the NPG films can be resolved with a bottom-up approach to the synthesis of NPG via constituent nanoparticles (NPs), aiming at controlling the properties of the films. The film formation was controlled by manipulating the trapping of nanoparticles (NPs) with temperature, ionic strength and protonation of the buffer. The morphology and properties of the NPG films can be resolved with a bottom-up approach to the synthesis of NPG via constituent nanoparticles (NPs), aiming at controlling the properties of the films. The film formation was controlled by manipulating the trapping of nanoparticles (NPs) with temperature, ionic strength and protonation of the buffer. The morphology and properties of the NPG films can be resolved with a bottom-up approach to the synthesis of NPG via constituent nanoparticles (NPs), aiming at controlling the properties of the films.
Gold nanomaterials such as AuNPs have many catalytic-catalytic functions in fine chemical synthesis and electrochemical interfacial reactions, and are among the most studied reactions catalyzed by AuNPs.31-33 CO oxidation4 as well as CO2 reduction.30 Dealloyed NPG is highly active and selective for CO oxidation at low temperatures.34 In general, CO2 decomposes efficiently via formation of CO.28 Difference in stability of adsorbed intermediates in the form of CO and COOH on gold surfaces is essential for effective CO2 conversion.29 The abundance of active edge sites in NPG can enhance its catalytic efficiency. Moreover, the catalytic selectivity towards CO reduction over hydrogen evolution is crucial for CO2 reduction in aqueous media. Since hydrogen binds relatively weakly to gold surfaces, gold is very suitable for CO oxidation.30 Another benefit of utilization of NPG is its high oxidation. Such electrocatalytic performance of cNPGFs can be attributed to the reaction parameters: catalytic properties are a “reflection” of structural characteristics. This opens the possibilities to design cNPGFs for specific tasks.

Results and discussion

Synthesis and formation of cNPGFs

The synthesis of cNPGFs starts by HAuCl4 reduction with 2-(N-morpholino)ethanesulfonic acid (MES), forming AuNPs. Herein MES exhibits a three-fold role as buffering, reducing and AuNPs stabilizing agent. This procedure has been modified from our previously reported synthesis of MES-stabilized AuNPs.19,20 The cNPGF formation is summarized in Scheme 1. KCl is added before AuNPs formation to increase the ionic strength of the solution and screen the electrostatic repulsion between NPs. This procedure has been modified from our previously reported synthesis of MES-stabilized AuNPs.19,20 The scheme depicts the steps of AuNP formation, precipitation and cNPGF formation.

![Scheme 1. Synthesis procedure of cNPGFs. MES buffer and KCl are preheated prior to gold precursor addition. Precursor is reduced by MES, forming AuNPs. MES reduces HAuCl4 to AuNPs. KCl increases ionic strength, decreasing both electrostatic repulsion between AuNPs and the sorption barrier for interfacial adsorption of AuNPs. Driven by the concentration gradient, bulk AuNPs diffuse towards the surface creating interfacial AuNP films. Finally, the addition of HCl induces deprotection of AuNPs by removal of MES, resulting in structural sintering and cNPGF formation.](image)

Finally, the addition of HCl effectively destabilizes the AuNPs, resulting in precipitation of particles in the bulk, and formation of sintered and robust cNPGFs. In fact, colloidal AuNPs could assemble at the air-water interface driven by vapor of formic acid, giving film-like appearance.12 Ordered AuNP monolayers have also been formed at water-hexane interface. However, films prepared with these methods consisted only of individual AuNPs and a continuous porous gold film was not formed due to lack of AuNP sintering.12,24 A range of reaction parameters have been tested to optimize conditions for obtaining desirable pore sizes and area coverage with mechanical stability and reproducibility. The observations are summarized in Table S1. An optimal concentration of KCl was identified. High concentrations caused rapid destabilization of AuNPs, precluding cNPGFs formation. cNPGF formation was enhanced when the pH of the MES buffer was increased from 7 to 8, suggesting a pH effect on cNPGF formation. It was shown previously that during formation of AuNPs with MES buffer, pH dropped drastically.26 We notice that the type of anion plays a role in cNPGF formation. For example, potassium phosphate, as one of the most common inorganic buffers, resulted in reduced film formation, whereas presence of chloride enhanced cNPGF formation. In contrast, when perchlorate (having low affinity for metal surfaces) was used to control ionic strength, little to no cNPGF formed. These observations highlight the strong and complex effects of pH, MES, anion adsorption and ionic strength on processes occurring during AuNP and cNPGF formation.
Structural characterization

The structure of the synthesized cNPGF was mapped characterized at different scales revealing a hierarchical porous structure. At macro-, micro- and nanoscales, the film appears porous with abundant interconnected networks, Figure 1.

Closer investigation of the nanostructure clearly shows the AuNP building blocks assembled into small NP agglomerates (Figure 2A) which further fused together to form extended networks of porous ligaments upon addition of HCl (Figure 2B). Similarity of nano- and microstructures are due to the 2D AuNP sintering nature at the air-liquid interface, Figure S6.

![Figure 1](image1.png)

Figure 1. Hierarchical porous structure of as-synthesized cNPGFs. (A) Photograph, (B) visible light micrograph and (C) TEM image showing porosity in three different size domains. The filled squares in (A) and (B) represent the areas of images in (B) and (C), respectively.

This is very different from the nanostructure of dealloyed NPG with continuous and smooth ligaments and pores. The concentrations of KCl and HCl influence strongly the film structure (Figure S5). This provides handles possibility for tuning the cNPGF morphology and porosity of the cNPGF. In general, increasing HCl concentrations enhances the size of the macro pores leading to web-like structures. Electron diffraction (inset of Figure 2B) confirms an fcc gold lattice with randomly oriented crystallites. The high-surface-area film displays long-range connectivity through metallic bonding with nanopores, Figure 2C-D. Since transmission electron microscopy (TEM) cannot provide direct information about film thickness, atomic-force microscopy (AFM) was employed to map the cNPGF, Figure 3. AFM displays a uniform thickness over the scanned area, Figure 3A. An average thickness of 500 ± 200 nm is found, with most of the film being between 300 and 600 nm, Figure 3B-C. The nanoporosity appears as surface roughness since the AFM tip cannot penetrate into the nanopores. The roughness of cNPGF, evidenced in the high magnification AFM image (Figure 3D), confirms the nanoscale porosity observed under TEM at a similar magnification (Figure 2B). The porosity of cNPGF is obtained on basis of measured values of areal density and thickness. The areal density of cNPGF was estimated via quartz-crystal microbalance (QCM) to be around 50 µg/cm² (Figure S1), which is consistent with 100 ± 50 µg/cm² measured by atomic absorption spectroscopy (AAS). Therefore, the apparent density is around 1.0-3.0 g/cm³, or 5-15% of bulk gold, and a porosity of 85-95%. 60-80% porosity is normally found for dealloyed NPGs using 12K and 9 K Au-Ag alloys. Close analysis of TEM images shows that the dimensions of the pores and ligaments in the cNPGF are around 25 nm (60% below 30 nm) while pores are up to several hundreds of nm are present, Figure S7 and S8.

![Figure 2](image2.png)

Figure 2. TEM micrographs of (A) AuNPs synthesized in 10 mM MES (pH 8.0) with 20 mM KCl before addition of 15 mM HCl and (B) same magnification of the as-synthesized NPG film. (C-D)
Magnified view of the interconnected porous nanostructure of cNPGF.

Figure 3. AFM images of cNPGF. (A) Large area (10 x 10 µm²) 3D image showing micro scale uniform film. (B-C) Height analysis of AFM image, showing a thickness of around 400 nm. The height profile in (C) is taken along the dashed line in (B). (D) High magnification of cNPGF.

The ligament thickness has been determined as 30 ± 10 nm, Figure S7. Interestingly, porosity is mainly in the mesoporous range and...}

Electrochemical characterization

Electrochemical characterization of cNPGFs was conducted using cyclic voltammetry (CV), and compared to cyclic voltammograms of AuNPs and single-crystal Au(111). Anodic and cathodic peaks around 0.25 V on Au(111) (blue curve in Figure 4A) originates from structural transitions of reconstructed (√3 x 22) to pristine (1 × 1) phases. The sharp, symmetric pair of peaks at ~ 0.8 V is a “fingerprint” feature of Au(111) in sulfuric acid, due to phase transition between the disordered and ordered (V3 x V7) R19.1° structure of adsorbed sulfate ions. The absence of these features in the cyclic voltammogram of AuNP (black curve) and cNPGF (red curve) suggests polycrystallinity in both materials, Figure 4A. Figure 4B shows cyclic voltammograms of cNPGF, AuNPs and Au(111) in a broad potential window. The anodic peaks from 1.05 to 1.35 V denote oxide layer formation at distinct low-index coordinated Au surfaces. The cathodic peak at 0.85 V is assigned to reductive removal of these oxide layers.

Figure 4. Cyclic voltammograms of Au(111), AuNPs and cNPGF in (A) narrow and (B) broad potential window. Cyclic voltammograms were recorded at 50 mV/s in 0.10 M H2SO4. The current densities related to Au(111) were increased four-fold for better comparison.

Electrical properties of cNPGF, and thus long-range connectivity in lateral dimensions were assessed with in situ conductance measurements performed in a four-electrode system. An interdigitated microarray Au electrode acted as WE, Figure 5A. Applied potential values were selected corresponding to the potential range studied in the CV experiments (Figure 4). The resultant apparent electrical conductivity values (σ) were derived from the changes in the Ohmic current (IΩ) flowing through the examined films according to the following equations:

\[ I_{\Omega} = \frac{I_{\text{WE}(1)} - I_{\text{WE}(2)}}{2} \]

\[ \sigma = \frac{I_{\Omega} \cdot w}{n \cdot l \cdot d \cdot V} \]

Where: I_{\text{WE}(1)} and I_{\text{WE}(2)} are the measured currents at corresponding WEs, w is the length of band electrodes, n the number of band electrodes, l the length of band electrodes, d the film thickness, and V is the bias voltage applied between WE(1) and WE(2).

Almost perfect linearity and thus a clear Ohmic response was obtained up to 0.8 V, Figure 5B. Above 0.8 V, the contribution of the Faradaic current flowing between WE and CE becomes dominant and is directly related to the Au oxide formation on Au surfaces. However, the deviation from an Ohmic response at higher potentials is relatively small, suggesting an overall effectiveness of electronic transport in cNPGF.
Figure 5. (A) Schematic diagram of electrochemical apparatus for in situ conductance studies in 0.1 M H₂SO₄. cNPGFs were uniformly deposited on interdigitated electrodes. Ohmic current was probed by keeping fixed bias between electrodes (1) and (2), during application of increasing and decreasing constant potential in the three-electrode system (saturated calomel electrode (SCE) as reference electrode (RE), Pt counter electrode (CE) and as working electrode (WE), respectively). (B) Apparent electrical conductivity vs applied potential for a typical cNPGF, together with the fitting line and corresponding equation. Reversing the potential steps yielded the same conductance values as going from 0.1 to 1.6 V. The onward and reverse potential cycles were repeated numerous many times, showing no change in recorded conductance values. The latter two findings are indicative of a very stable electrochemical response of cNPGF and practically negligible hysteresis effects. Estimated values of electrical conductivity for cNPGF are in the upper semiconducting range, since the slope of approx. 120 S cm⁻¹ V⁻¹ was obtained. The resistivity values for de-alloyed NPG can be e.g. 100-fold lower. The reason for this discrepancy is unclear, since TEM images (see Figure 2) strongly support structural interconnectivity of cNPGFs. It might be speculated that the conductance of cNPGFs is affected by microscopic electron-impeding boundaries in ligament joints, but the assessed electrical properties of cNPGF are suitable for application as chemical sensors or electrocatalysts.

**Electrocatalysis**

The catalytic properties of cNPGF towards CO/CO₂ oxidation/reduction were assessed by immobilization of the studied films on glassy carbon electrodes (GCEs) in a three-electrode system.

CO₂ reduction was conducted in CO₂-saturated 0.10 M NaHCO₃ electrolyte, Figure 6A. The onset potential for the reaction was −1.0 V vs SCE representing a 100 mV reduction in overpotential compared to reports for polycrystalline Au. The increased cathodic current densities for cNPGF in CO₂ relative to Au, indicate a high-yielding reaction with limited H₂ evolution and 88% CO₂ reduction at −1.2 V. Generally, gold nanostructures in 0.10 M NaHCO₃ convert CO₂ mainly to CO with current efficiency around 87%. Because of the significant porosity of cNPGFs and the proximity of a range of different active sites within the pores, prompted the study of electrochemical CO to CO₂ oxidation of CO on cNPGF. This was conducted via rotating disk electrode (RDE) experiments in CO-saturated 0.10 M KOH electrolyte, Figure 6B. cNPGFs exhibit characteristic features in cyclic voltammograms in Ar-saturated 0.10 M KOH electrolyte without any current increase at the rotation rate of 1600 rpm. In CO-saturated electrolyte, anodic current densities increase with rotation rate. The plateauing anodic currents in the potential region from −0.5 to 0.35 V are limited by CO(aq) diffusion to the Au active sites, which is controlled by the rotation rate. Previously reported onset values for CO oxidation, conducted on single-crystal Au(111) electrode are at −0.61 V vs SCE. Our cNPGFs catalyzed CO oxidation at the onset potential of −0.71 V, thus resulting in 100 mV lower overpotential. This agrees well with the results obtained for CO₂ reduction, confirming the high catalytic activity of cNPGFs.
activity for both CO₂ reduction and CO oxidation through the abundance of active sites at the cNPGF.

Conclusions and perspectives

Gold nanostructures such as AuNPs, gold nanorods, core-shell NPs have been extensively developed extensively over the last two decades. In contrast to bulk gold, its nanostructures are active and their applications in electronic, chemical and optical properties offer potential applications in molecular electronics, bio-imaging, nanomedicine, sensors, catalysts, and nanotechnology. As a perspective, cNPGFs can be used as gold impact on bioelectrocatalysis and biofuel cells. The large NPG does not need any additional solid support and can be directly used as nanocatalysts. Development of de-alloying strategy can be employed to develop highly efficient bioanodes for immobilization of electrochemically active bacteria for microbial fuel cells due to the highly porous structures. In contrast to bulk gold, its nanostructures are active and their outstanding electronic, chemical and optical properties offer potential applications in molecular electronics, bio-imaging, nanomedicine, sensors, catalysts, and nanotechnology. Gold nanostructures such as AuNPs, gold nanorods, core-shell NPs have been extensively developed extensively over the last two decades. In contrast to bulk gold, its nanostructures are active and their applications in electronic, chemical and optical properties offer potential applications in molecular electronics, bio-imaging, nanomedicine, sensors, catalysts, and nanotechnology. As a perspective, cNPGFs can be used as gold impact on bioelectrocatalysis and biofuel cells. The large NPG does not need any additional solid support and can be directly used as nanocatalysts. Development of de-alloying strategy can be employed to develop highly efficient bioanodes for immobilization of electrochemically active bacteria for microbial fuel cells due to the highly porous structures. In contrast to bulk gold, its nanostructures are active and their outstanding electronic, chemical and optical properties offer potential applications in molecular electronics, bio-imaging, nanomedicine, sensors, catalysts, and nanotechnology.

Experimental section

Chemicals

Gold(III) chloride trihydrate (HAuCl₄·3H₂O, 99.5%, Aldrich), potassium chloride (KCl, 99.9%, Aldrich), potassium phosphate dibasic (K₂HPO₄, ≥95%, Aldrich), potassium hydroxide (KOH, ≥95%, Aldrich), potassium hydroxide (KOH, ≥95%, Aldrich), potassium phosphate monobasic (KH₂PO₄, 99.995%, TraceSELECT®, Fluka), potassium phosphate dibasic (K₂HPO₄, 99.999%, TraceSELECT®, Fluka), potassium hydroxide (KOH, ≥99.9%, Aldrich), potassium hydroxide (KOH, ≥99.9%, Aldrich), hydrochloric acid (HCl, 35-37%, Aldrich), potassium phosphate monobasic (KH₂PO₄, 99.995%, TraceSELECT®, Fluka), potassium phosphate monobasic (KH₂PO₄, 99.995%, TraceSELECT®, Fluka), nitric acid (HNO₃, ≥65%, puriss. p.a., Fluka), potassium hydroxide (KOH, ≥99.9%, Aldrich), potassium hydroxide (KOH, ≥99.9%, Aldrich), perchloric acid (HClO₄, 70% in water, 99.999% trace metal basis, Aldrich), sodium bicarbonate (NaHCO₃, ≥99.7%, Fluka), and sulfuric acid (H₂SO₄, ≥95%, TraceSELECT®, Fluka) were used without further purification. Aqua regia was prepared by mixing nitric and hydrochloric acids in 1:3 ratio. Aqueous solutions were prepared with MilliQ purified water (18.2 MΩcm, Sartorius). Ar (5.0 N, AGA A/S) and H₂ (5.0 N, AGA GAS AB) were used for electrochemical experiments. Electrocatalytic experiments were performed with CO (4.7 N, Linde AG) and CO₂ (5.2 N, AGA GAS AB).

Synthesis of cNPGFs

In a standard synthesis, 2 mL of ultrapure water, 4 mL 100 mM KCl and 2 mL 100 mM MES (pH 8.0) were mixed in a 50 mm ϕ beaker, which was placed in an oil bath for pre-heating at 80 °C for 5 min. Then, 2 mL 20 mM HAuCl₄ was added while the beaker was kept at 80 ° C. AuNPs were formed almost instantly and the reaction was left to proceed for 30 min. At this point, an oily "pre-film" consisting of individual microscopic gold assemblies at the air-liquid interface had formed. Upon addition of 75 µL of 4.0 M HCl, the cNPGF morphology started to change for 10 min at 80 ° C after which the beaker was removed from the oil bath. After the synthesis, it is important to separate the film from the AuNP solution to stop the process completely. The reaction mixture was partly removed from underneath the cNPGF morphology started to change for 10 min at 80 ° C after which the beaker was removed from the oil bath. After the synthesis, it is important to separate the film from the AuNP solution to stop the process completely. The reaction mixture was partly removed from underneath the cNPGF. The large surface area of cNPGFs will increase loading of microorganisms such as metalloproteins and enzymes and therefore increase their signals. The three-dimensional nanoropes inside cNPGFs are expected to improve stability of the proteins and enhance their electrochemical activities. Immobilization of the protein molecules on cNPGFs can be achieved through surface modification or reactions as that those have been well-developed on gold surfaces. Such strategy can be employed to develop highly efficient bioanodes and biocathodes for enzymatic biofuel cells. Thanks to the large surface area, cNPGFs provide a class of electrode materials for immobilization of electrochemically active bacteria for microbial fuel cells due to the highly porous structures. In addition, cNPGFs can assemble on a pair of electrodes as transparent organic/inorganic membrane or substrates, such as glass slides. cNPGFs offer an alternative of to conducting glass indium tin oxide glass for spectroelectrometry and photochemistry.

Microscopy techniques

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Commented [NS5]: I think it could be quite confusing to the new reader. Of if we really want, then at least put the quotation marks ("oily") since there was no oil.
Visible light micrographs were recorded on a Leica M205C from Leica Microsystems (Wetzlar, Germany). Transmission electron micrographs were obtained on a Tecnai G2 T20 from FEI Company (Hillsboro OR, USA) at 200 kV accelerating voltage. TEM samples were prepared by placing a grid on top of a drop of a suspension of the particles in water. The glassy carbon (GCE, ϕ = 4.0 mm, A = 0.1256 cm²) and rotating disk electrode (RDE, ϕ = 5.61 mm, A = 0.2472 cm²) were wet-polished by sand paper (grit roughness 2000, followed by 4000) for 10 min by hand. Further polishing with alumina slurry (particle sizes of 1.0, 0.3 and 0.05 µm in diameter), was done on a polishing machine (300 rpm, 5 min per cycle) providing a mirror-like electrode surface finish. The electrodes were sonicated for 30 min in total with intermittent water exchange. The cNPFG immobilization was performed by submersion of the electrode in a 2 L beaker containing ultrapure water and floating cNPFGs, followed by a rapid upward movement of the electrode to effectively entrap the film on the electrode surface. The covered electrodes were dried at room temperature in the fume hood. Afterwards, excess film was removed by a micropipette tip to match the electrode surface. Finally, the immobilized cNPFGs was protected by drop casting 5 µl of 0.05 wt% Naflon solution (dissolved in ethanol) and drying at 60 °C for 20 minutes. As-prepared electrodes were imaged with an optical microscope, and cNPFGs areas integrated to precisely report the material activity, Figure S3. All electrochemical measurements were performed in Faraday cage at room temperature (20 ± 2°C) using Autolab PGSTAT 12 (Metrohm, The Netherlands) controlled by the general purpose electrochemical system (GPES) or NOVA 1.11 software. The distance between each band electrodes (1.1 mm long, 40 in total) was 20 µm. Deposited cNPFGs were rinsed with water and dried at ca. 60 °C for 1 h. Monitoring the changes in conductance upon increasing potential was realized by application of constant bias of 0.5 V between electrodes. The measurements in 0.10 M H₂SO₄ and Au electrode were conducted utilizing Autolab PGSTAT 12. The control of applied potentials and recording of obtained currents were done in bipotentiostat mode. Further details on the in situ conductance measurement set-up and calculations have been reported previously.56

### Electrical experiments

The stationary glassy carbon electrode (GCE, ϕ = 4.0 mm, A = 0.1256 cm²) and rotating disk electrode (RDE, ϕ = 5.61 mm, A = 0.2472 cm²) were wet-polished by sand paper (grit roughness 2000, followed by 4000) for 10 min by hand. Further polishing with alumina slurry (particle sizes of 1.0, 0.3 and 0.05 µm in diameter), was done on a polishing machine (300 rpm, 5 min per cycle) providing a mirror-like electrode surface finish. The electrodes were sonicated for 30 min in total with intermittent water exchange. The cNPFG immobilization was performed by submersion of the electrode in a 2 L beaker containing ultrapure water and floating cNPFGs, followed by a rapid upward movement of the electrode to effectively entrap the film on the electrode surface. The covered electrodes were dried at room temperature in the fume hood. Afterwards, excess film was removed by a micropipette tip to match the electrode surface. Finally, the immobilized cNPFGs was protected by drop casting 5 µl of 0.05 wt% Naflon solution (dissolved in ethanol) and drying at 60 °C for 20 minutes. As-prepared electrodes were imaged with an optical microscope, and cNPFGs areas integrated to precisely report the material activity, Figure S3. All electrochemical measurements were performed in Faraday cage at room temperature (20 ± 2°C) using Autolab PGSTAT 12 (Metrohm, The Netherlands) controlled by the general purpose electrochemical system (GPES) or NOVA 1.11 software. The distance between each band electrodes (1.1 mm long, 40 in total) was 20 µm. Deposited cNPFGs were rinsed with water and dried at ca. 60 °C for 1 h. Monitoring the changes in conductance upon increasing potential was realized by application of constant bias of 0.5 V between electrodes. The measurements in 0.10 M H₂SO₄ and Au electrode were conducted utilizing Autolab PGSTAT 12. The control of applied potentials and recording of obtained currents were done in bipotentiostat mode. Further details on the in situ conductance measurement set-up and calculations have been reported previously.56

### Areal density estimation

The areal density of cNPFGs was investigated by quartz-crystal microbalance (QCM) measurements on a QCM200 instrument (SRS Stanford Research Instruments, USA) utilizing the Sauerbrey equation.57 The used QCM sensors were coated with Au (surface area = 1.327 cm²), with sensitivity of 56.6 Hz μg⁻¹ cm², and oscillating at around 5 MHz. Atomic adsorption spectroscopy (AAS) was used as a secondary method to determine the areal density of cNPFGs, and carried out on a Perkin Elmer Atomic Absorption Spectrometer 4100 and a light source from S&I Juniper & Co (Au, 242.8 nm). The samples were quantified by recording a standard curve alongside the samples. Triple determination was performed for all data points. Samples were prepared by transfer of floating cNPFG from water bath onto a glass slide or GCE. The GCE area is well-defined and the Au covered glass slide area was defined by analyzing micrographs of the slides with ImageJ software. The cNPFG was dissolved by keeping the specimens in aqua regia for overnight. The glass slide or electrode was removed from the vial and the solution was heated to dryness. The solid remnants were dissolved in 5 mL of 10 vol% HCl as recommended by Perkin Elmer to use for Au AAS on their instrument.14

### Conflicts of interest

There are no conflicts to declare. The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

### Acknowledgements

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References


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Supporting information
Figure S1. Typical massograms obtained during recording of QCM signals of bare (blue line) and eNPGF-coated (red line) sensors in air. The difference in frequency of oscillating sensors yields apparent areal density of deposited film.

Figure S2. AAS measurements showing the standard curves and the measured samples.
Figure S3. Micrograph images of GC electrodes under an optical microscopy after electrochemical and catalysis testing (A-D) Images processed in software to become black and white, for area determination treatment (E-H). cNPGF coverages of GC area are shown in images.

Figure S4. Cyclic voltammograms of cNPGF film deposited on RDE (A = ca. 0.18 cm$^2$) in Ar-purged (A) 0.10 M H$_2$SO$_4$ solution recorded at different scan rates, (B) 0.10 M KOH solution recorded at 20 mV/s at stationary conditions (black line) and at electrode rotation rate of 1600 rpm (red line). Current densities are normalized per geometrical area of the electrode.
**Figure S5.** TEM images showing changes of cNPGF in microstructure following changes in both KCl and HCl concentrations. Images are all taken with identical magnification, for size comparison.
The cNPGF structure has similar nano- and microstructure features (Figures 1B and 1C). This originates from the nature of Au NPs aggregation during the entire cNPGF formation. The cNPGF assembles at the air-liquid interface. The predominantly spherical Au NPs undergo interfacial sorption. When at the interface, Au NP moves in two dimensions and sinters to individual Au NP or cluster of NPs. Due to surface tension forces, synthesis solution evaporation and size of NPs, the sintered nanoscale structure contains fine pores (nm range). As the cNPGF synthesis continues, the nanoscale structures continuously expand by sintering to other nanoscale units. By sintering into larger 2D networks, the microscale features are formed. Due to the nature of material sintering, the nanoscale features are reproduced and maintained within the newly formed microscale structure, resulting in similar morphology observed at two different scales, Figure S6.
Figure S7. (A) TEM image of cNPGF film with highlighted AuNP boundaries (blue). (B) A AuNP grain boundary size distribution histogram. (C) Scheme of a ligament attachment of individual AuNPs into cNPGF films.

Figure S8. (A) TEM image of cNPGF with highlighted pores (purple) and (B) the corresponding pore size distribution histogram.
Table S1. The summary of cNPG synthesis parameters, chemical concentrations and experimental observations.

<table>
<thead>
<tr>
<th>Type of additive</th>
<th>Concentration range [mM]</th>
<th>Observed effect</th>
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<tr>
<td>KCl</td>
<td>10-50</td>
<td>Increased amount of film formation, and increased ligament thickness</td>
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<tr>
<td>HCl</td>
<td>10-20</td>
<td>Larger macropore-to-mesopore ratio</td>
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