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Definition of a scoring parameter to identify low-dimensional materials components

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The last decade has seen intense research in materials with reduced dimensionality. The low dimensionality leads to interesting electronic behavior due to electronic confinement and reduced screening. The investigations have to a large extent focused on 2D materials both in their bulk form, as individual layers a few atoms thick, and through stacking of 2D layers into heterostructures. The identification of low-dimensional compounds is therefore of key interest. Here, we perform a geometric analysis of material structures, demonstrating a strong clustering of materials depending on their dimensionalities. Based on the geometric analysis, we propose a simple scoring parameter to identify materials of a particular dimension or of mixed dimensionality. The method identifies spatially connected components of the materials and gives a measure of the degree of “1D-ness,” “2D-ness,” etc., for each component. The scoring parameter is applied to the Inorganic Crystal Structure Database and the Crystallography Open Database, ranking the materials according to their degree of dimensionality. In the case of 2D materials the scoring parameter is seen to clearly separate 2D from non-2D materials and the parameter correlates well with the bonding strength in the layered materials. About 3000 materials are identified as one-dimensional, while more than 9000 are mixed-dimensionality materials containing a molecular (0D) component. The charge states of the components in selected highly ranked materials are investigated using density functional theory and Bader analysis showing that the spatially separated components have either zero charge, corresponding to weak interactions, or integer charge, indicating ionic bonding.

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I. INTRODUCTION

Low-dimensional materials with one or more characteristic lengths of the materials limited to the atomic scale have received significant attention recently. Since the discovery of graphene the world has seen intense research in 2D materials involving synthesis and investigation of mechanical, electronic, magnetic, and catalytic properties of new materials [1–4]. Also a number of computational efforts have been dedicated to the identification of new 2D materials and to the construction of computational databases with information about their stability and (photo)electronic properties [5–7].

One of the driving forces behind this research has been an interest in ultrasmall electronic components and this has also led to studies of 1D or quasi-1D materials as possible interconnects [8,9]. Furthermore, the possibility of combining materials of different dimensionality into new van der Waals bonded mixed-dimensional heterostructures has recently been discussed [10]. The realization of such structures relies on the identification of appropriate weakly interacting material components of different dimensionalities.

In the following we shall define a simple geometrical scoring parameter to identify low-dimensional components in existing materials. The scoring parameter is easy to compute and can be applied to large materials databases. We illustrate this by mining the Inorganic Crystal Structure Database [11] (ICSD) and the Crystallography Open Database [12] (COD) to find materials with clearly identifiable low-dimensional atomic structures. The identified materials consist of weakly interacting components as we demonstrate for 2D materials by comparison with previously calculated exfoliation energies. Apart from being interesting in their own right, the materials components may also form templates for substitution of similar chemical elements to form new materials of different dimensions [5,7].

II. RESULTS AND DISCUSSION

A. Bond-length interval analysis

The definition of the scoring parameter requires, first, that we can identify the dimension(s) of a periodic solid. Given an atom in a bonded cluster, the cluster dimension is given by the rank of the subspace spanned by the atom and its periodically connected neighbors. We refer to this method as the rank determination algorithm (RDA), which is described in detail in the Methods section.

An accurate identification of bonded clusters requires a full electronic structure calculation, where the bond strength and character can be addressed. However, for purposes of screening large materials databases this approach is computationally infeasible. Instead, we use a simple geometric criterion for bonding. We describe two atoms, i and j, as bonded if the distance between them is less than a specified multiple of their covalent radius sum:

$$d_{ij} < k(r_i^{cov} + r_j^{cov}).$$

(1)
Here, \(d_{ij}\) is the distance between atoms \(i\) and \(j\), \(r_{ij}^{\text{cov}}\) and \(r_{ij}^{\text{cov}}\) are the corresponding covalent radii \([13]\), and \(k\) is a variable to be investigated. The latter choice is motivated by the strong dependence of the classification of the dimensionality of a material upon the \(k\) value; as illustrated for the boron-nitride structure in Fig. 1(a), too small a \(k\) value will underestimate the dimensionality, whereas too large a \(k\) value will overestimate it. Rather than attempt to identify a good value of \(k\), we observe that, for any given structure, there exists a finite number of relevant \(k\) intervals to investigate.

We start by considering the set of interatomic distances in a material, sorted by increasing \(k\) value, where \(k = (\sqrt{r_i^{\text{cov}} + r_j^{\text{cov}}})/d_{ij}\). Each interatomic distance corresponds to a possible bond; as shown in Fig. 1(a), bonds can be physical or not. Bonds are inserted one at a time, and at each step the RDA process finds all \(k\)-value intervals in which the dimensionality classification is constant, of which there are a finite number. The process terminates when a single 3D component is left; i.e., all atoms are contained in the same bonded cluster. This process finds all \(k\)-value intervals in which the dimensionality classification is constant, of which there are a finite number. The interval identification method is described in more detail in the Methods section.

Figure 1(b) shows the application of this method to two different layered structures. It can be seen that different dimensionality classifications exist at different \(k\) values. Furthermore, the intervals have very different widths; the first interval is of the form \([0, k)\), whereas the last interval is of the form \([k, \infty)\).

**B. Defining the scoring parameter**

Figures 2(a) and 2(b) show the \(k\) intervals for all structures in the ICSD and COD with, respectively, a 2D interval and a 1D interval. In both cases there is a visible cluster of structures in the approximate region \(k_1 \approx 1\) and \(0.1 \leq k_2 \leq 1.5\). The position of the cluster is intuitive from a bonding perspective. First, if the bonding model and covalent radii exactly described the actual bond lengths, the cluster would lie on the line \(k_1 = 1\); the variability in the interval starting points results from the simplicity of the ball-and-stick bonding model. Second, since low-dimensional components are geometrically separated, we should expect a correspondingly wide \(k\) interval; it can be seen that easily exfoliable structures such as graphite, boron nitride, and molybdenum disulfide have wide \(k\) intervals.

We propose a scoring parameter which distills the above observations of the \(k\)-interval plots into a single number:

\[
s(k_1, k_2) = f(k_2) - f(k_1),
\]

where

\[
f(x) = \frac{c \times \max(0, x - 1)^2}{1 + c \times \max(0, x - 1)^2}.
\]

Here, \(c\) is a constant which determines the scale at which a bond is broken. We use \(c = 1/0.15^2\), which is chosen so that \(s(1, 1.15) = s(1.15, \infty) = 0.5\); slightly different values of the parameter will give similar results. Figure 3 illustrates how a \(k\) interval is transformed into a score.

The interval width increases the score, but with diminishing returns as \(k_1\) increases above 1. This avoids the \([k, \infty]\) interval dominating unless \(k_1\) is close to 1, in which case the structure is indeed 3D. Furthermore, \(k\) values below 1 are effectively set to 1; this avoids erroneous low-dimensional classifications when \([k_1, k_2] \approx [0, 1]\). In structures with low-dimensional components, the scoring parameter rewards large intercomponent distances. A further convenient property of the scoring scheme is that the interval scores sum to 1. We have found that the best results are achieved by merging \(k\) intervals with the same types of dimensionality classifications [e.g., intervals of the same color in Fig. 1(b)].
The principal motivation of the scoring scheme is to identify the intuitively correct dimensionality classification, by determining whether a \( k \) interval lies within a cluster of the type shown in Figs. 2(a) and 2(b). Using the scoring parameter, the structures are colored in Figs. 2(c) and 2(d) according to their dimensionality classification.

The scoring scheme is demonstrated for three structures in Fig. 4. The first material, Al\(_2\)O\(_3\), is clearly a bulk crystalline structure. If a single \( k \)-value threshold at \( k \approx 1 \) were used, however, it would result in a misclassification as a layered structure. Similarly, the scoring scheme also ensures that the Ti\(_2\)Cl\(_2\)N\(_2\) structure is correctly identified as a layered structure. The AuTe\(_2\) has an ambiguous classification, lying close to the contour \( s(k_1, k_2) = 0.5 \). In this case the dimensionality classification is sensitive to small changes in the functional form or the parameters of the scoring function. Then, the useful information contained in the scores is not in their exact values, but rather that \( s_2 \) and \( s_3 \) are approximately equal in value; this can be interpreted as a layered structure with a very small interlayer spacing.

It should be emphasized that the scoring is exclusively based on interatomic distances and atomic sizes, and that it simply assumes that longer bonds tend to be weaker than shorter ones. The physical characters of the bonds, i.e., whether they can be considered covalent, ionic, or of dispersion type, are not revealed. Nonetheless, the coarse treatment of bond lengths is justified by the cluster separation in Fig. 2(b). We will show that the scoring scheme allows for identification of interesting materials, whose properties can then be investigated experimentally or using electronic structure methods.

The scaled bonding criterion described in Eq. (1) is the same one employed by Ashton et al. [6] in their study of layered materials. An additive bonding criterion of the form \( d_{ij} < r_i + r_j + \Delta \) is used by Mounet et al. [5] and Cheon et al. [14], using van der Waals radii and elemental radii, respectively. In these works, the material dimension is determined by sampling a range of parameter values (either \( k \) or \( \Delta \)) in a fixed interval, which does not easily permit the construction of a scoring parameter. Cluster dimensionalities
are determined using a topology-scaling algorithm (TSA) [6] (also proposed in [14]), which relates the dimension to the number of bonded clusters as a function of the size of a periodic supercell, or using the RDA [5]. Due to the use of a fixed-size supercell, the TSA and RDA can respectively underestimate and overestimate the number of bonded clusters in certain materials with complex geometries. In the methods section we describe a variant of the RDA which correctly assigns all atoms to bonded clusters without the need to specify a supercell. Except for such complicated cases, however, our definition agrees with the TSA and the supercell RDA.

Other methods for identification of layered materials include the analysis of the packing fraction [15,16], identification of layered slab structures [17], and the use of discrepancies between experimental lattice constants and those predicted by density function theory (DFT) [18]. By identifying structures with interlayer sodium atoms, Zhang et al. [19] have investigated promising layered cathode materials for sodium-ion batteries. McKinney et al. [20] have extended this search to general “ionic layered” structures.

C. Mining the ICSD and COD

We have analyzed all materials in the ICSD and COD using the proposed scoring parameter. Figure 5 shows examples of materials with different dimensionalities and high values of the scoring parameter.

The database has been filtered in standard ways [5,21] by removing incomplete and/or defective entries, structures with more than 200 atoms, structures with partial occupancies, theoretically calculated structures, and structures with missing hydrogen atoms. Duplicate structures are removed using the structure matcher function of PYMATGEN [22]. The filtering process reduces the initial set of 585 485 CIF files to 167 767 structures. The filtering statistics are shown in Table I.

An overview of the database is shown in Table II. In this table the materials have all been categorized by the dimension(s) with the largest value of the scoring parameter \( s \). In some cases all \( s \) values may be fairly small and the classification is then rather uncertain. A large number of materials (105 199) are classified as 0D. These are mostly molecular crystals, which we shall not consider any further here.

The second largest category is the 3D materials. Most of these have a single 3D component, but some of them have two components still with large \( s \) values. Two examples [Ag(B(CN)4) and Ca(C(CN)3)2] are shown in Fig. 5. As can be seen from the figure, both materials consist of two identical interpenetrating networks which cannot be disentangled without breaking bonds. The two networks are sufficiently spatially separated to give scoring values above 0.7. (In the figure the two networks are colored red and blue.)

4623 materials are identified as two-dimensional, which is about 2.8% of all materials. This can be compared to for example the study by Mouvet et al. [5] where they find 1825 out of 108 423 materials (or also about 1.7%) of the materials to be easily or potentially exfoliable. About 2% of the materials are classified as 1D.

There are also some materials with several components of different dimensionality. In particular there are 9459 materials which have one or more 0D components in combination with components of higher dimensionality. These correspond to molecules or molecular ions embedded in the higher-dimensional network. Only a few materials combine 1D, 2D, and 3D components. We find 15 materials combining 1D and 2D. There are 22 materials which combine 1D and 3D components. Three of them are shown in Fig. 5.

While we have made every effort to remove inconsistent structures from the database, an automated filtering is not sufficient given the many different types of errors and partial structures present in the ICSD and COD. The numbers presented here should therefore be taken as only approximate.

A database containing the calculated scoring parameters for all dimensionalities for all compounds in the ICSD and COD is available at the Computational Materials Repository [23].

D. Physical significance of the scoring parameter

Due to the well-defined identification of the 2D materials, the scoring scheme also serves as a simple predictor of exfoliability. Mouvet et al. [5] have calculated the exfoliation energy (i.e., the binding energy between layers) of 1535 layered materials, and they suggest an energy of 35 meV/Å² as the
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FIG. 5. Smorgasbord of low-dimensional and mixed-dimensional materials, identified by applying the automatic dimensionality classification method to materials in the ICSD.

threshold for “easily exfoliable” materials. They furthermore highlight 11 materials, which they denote as “well-known” 2D materials.

In Fig. 6 we show the calculated exfoliation energies versus the scoring parameter $s_2$. There is a clear correlation between the scoring parameter and the exfoliation energy with essentially all of the high-scoring materials (say $s_2 > 0.7$) having an exfoliation energy below the threshold. The separation of materials of different dimensionality is also clearly seen here by the low density of points in the region $s_2 \approx 0.3–0.5$. The 11 well-known 2D materials are also shown in the figure. All of them, except Bi$_2$Te$_3$, have high-scoring values with $s_2 > 0.7$. Despite its small interlayer distance, Bi$_2$Te$_3$ is nonetheless classified as a 2D material, since $s_2$ is larger than its other scores.

It should be noted that although the exfoliation energy is a highly relevant quantity for the exfoliation process, it is not clear whether an absolute threshold in energy is the best indicator of exfoliability. The exfoliation process involves breaking the bonds between the layers keeping the bonds within the layers intact, so the exfoliation energy should be seen relative to the intralayer bond strengths. While the scoring parameter proposed here does not explicitly involve the energetics, the high-scoring materials have a clear separation between the intra- and intercomponent bond lengths, which can be expected to be a characteristic of easily exfoliable materials.

E. Ranking of low-dimensional materials

In addition to dimensionality classification, the scoring parameter defines an order on materials. We have identified the ten materials in the ICSD and COD with the highest 2D scores, shown in Table III. Widely studied layered structures such as graphene, boron nitride, and magnesium chloride are highly ranked. Some of the remaining structures have much larger unit cells, but are nonetheless clearly van der Waals
TABLE I. Number of structures remaining after each stage of filtering, performed in the order shown. “Defective” structures encompass incorrect CIF files and theoretical structures, and manually identified entries such as misfit compounds, surface structures, and superstructures. Where duplicate structures are found across the two databases, the COD structure is kept.

<table>
<thead>
<tr>
<th>Dimensions</th>
<th>COD</th>
<th>ICSD</th>
</tr>
</thead>
<tbody>
<tr>
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<td>105199</td>
<td></td>
</tr>
<tr>
<td>1</td>
<td>3503</td>
<td>3285</td>
</tr>
<tr>
<td>2</td>
<td>2946</td>
<td>15</td>
</tr>
<tr>
<td>3</td>
<td>3010</td>
<td>22</td>
</tr>
</tbody>
</table>

TABLE II. Number of entries of each dimensionality type found in the ICSD and COD. In the diagonal the number of materials with a single dimension are shown while the off-diagonal entries indicate materials with components of two different dimensionalities. In addition to the single- and bidimensional materials counted here, we have found 16 tridimensional structures with 0D, 1D, and 2D components.

<table>
<thead>
<tr>
<th>Dimensions</th>
<th>COD</th>
<th>ICSD</th>
</tr>
</thead>
<tbody>
<tr>
<td>0</td>
<td>105199</td>
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<td>2946</td>
<td>15</td>
</tr>
<tr>
<td>3</td>
<td>3010</td>
<td>22</td>
</tr>
</tbody>
</table>

bonded layered structures. It should be noted that the detailed ordering of the top materials is sensitive to the detailed choice of the function \( f(x) \) in Eq. (3), whereas the overall classification of the materials is more robust.

Similarly a list of the highest-scoring 1D materials is provided in Table IV. We shall not discuss these materials in depth here, but briefly characterize the top five entries with two or three different chemical elements. For all of these the one-dimensional or chainlike character has already been recognized. NPF\(_2\) consists of chains of alternating nitrogen and phosphorous atoms with the fluorine atoms bound to the phosphorous. The chains can also close on themselves forming ring-shaped molecules. SO\(_3\) is an asbestos-like structure made up of corner-linked SO\(_4\) tetrahedra forming spiraling chains, while the chains in the SiS\(_2\) structure consists of edge-sharing tetrahedra. The two ruthenium compounds form chains of ruthenium atoms. Ru(CO)\(_4\) is constructed from planar units with ruthenium in the middle and CO molecules attached with a fourfold rotation symmetry. These units are then stacked forming chains of ruthenium. RuCl\(_3\) is in the \( \beta \) phase also called the ZrCl\(_3\) structure. Here again the ruthenium atoms form linear chains, but with the chlorine atoms connecting two adjacent ruthenium atoms. It can be noted that RuCl\(_3\) also appears as a strongly layered material (\( s_2 = 0.933 \)) in the \( \alpha \) phase with prototype RhBr\(_3\). We have performed density functional calculations for these highly 1D compounds using the GPAW code [24,25] and the Atomic Simulation Environment (ASE) [26,27]. The three compounds NPF\(_2\), SO\(_3\), and SiS\(_2\) are found to be nonmagnetic large band gap semiconductors.

In the two compounds with ruthenium chains the distance between the ruthenium atoms are in fact comparable to the bond distance in ruthenium bulk metal. However, the strong couplings to the attached atoms and molecules lead to opening of band gaps. According to the DFT calculations Ru(CO)\(_4\) is nonmagnetic while RuCl\(_3\) is found to be antiferromagnetic. Details of the calculations can be found in the Supplemental Material [28].

The scoring approach also allows for identification of materials with several components of different dimensionality. Five of these are shown in Fig. 5. The two 1D+2D materials and the last two 1D+3D materials all involve alkali atoms (Na or Li) decorated with water molecules as the 1D component. Note that in the case of Na\(_2\)(H\(_2\)O)\(_3\)(Mn(NCS)\(_3\))\(_2\) the chains penetrate the 2D layers in the framework, while in the case of Na\(_3\)HP\(_2\)O\(_7\)(H\(_2\)O)\(_3\) the chains run parallel to the 2D components.

In these materials charge transfer takes place with approximately one electron per alkali atom donated to the 2D or 3D framework. We have investigated this by performing DFT calculations followed by a Bader analysis, [29,30] where an electronic charge is associated with each atom based on a natural per-atom division of the electronic density. For the compounds Na\(_2\)(H\(_2\)O)\(_3\)(Mn(NCS)\(_3\))\(_2\) and Na\(_2\)(H\(_2\)O)\(_3\)Ni\(_6\)(CO\(_3\))\(_6\)(OH)\(_6\) the chains consist of Na(H\(_2\)O)\(_3\) units with an electron transfer of 0.85 and 0.90 electrons per unit, respectively. Similarly in the case of Li(N\(_3\))(H\(_2\)O), the charge transfer is 0.85 per Li(H\(_2\)O)\(_3\) unit. Na\(_3\)HP\(_2\)O\(_7\)(H\(_2\)O)\(_3\) contains chains of Na(H\(_2\)O)\(_4\) molecular ions with a charge transfer of 0.80 electrons per unit.

The charge transfer in these systems illustrates that the geometrically defined scoring parameter does not only identify components which are exclusively van der Waals bonded to each other. Charge transfer may take place between spatially separated components giving rise to bonding of a more ionic character.

The last compound in Fig. 5 of mixed dimensionality is S\(_2\)Ta\(_2\)P\(_2\). It has a very intriguing structure. It consists of spiraling sulfur chains penetrating a 3D network constructed of TaPS\(_6\) building blocks. The 3D network itself consists of two identical interpenetrating components. The ability of Ta-P-S compounds to form tunnels has previously been reported [31] and sulfur spirals appear in several compounds. The present compound does not exhibit any charge transfer between the components. Figure 7 shows the calculated density of states for the S\(_2\)Ta\(_2\)P\(_2\) compound projected onto the 1D and 3D components. The two components are seen to exhibit different band gaps. This opens the possibility for selectively exciting electrons in one of the components using light with an appropriate wavelength.

III. CONCLUSION

We have defined a simple geometric scoring parameter to identify materials of particular dimensionality. The parameter provides an estimate of the degree to which a given dimensionality is present in the compound. The parameter is easy to calculate and can be applied to large materials databases.
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As mentioned in the introduction several computational 2D materials databases are presently under construction while 1D materials and materials of mixed dimensionality have received much less attention. The present approach allows for simple identification of existing 1D or mixed-dimensional materials, which can form templates that can be used to construct larger computational databases for materials of a given dimensionality.

IV . METHODS

A. Component dimensionality

A material will in general consist of several clusters of bonded atoms. Such clusters we term the components of the material. The components may have different dimensionalities and they should therefore be investigated separately.

Our definition of material dimensionality of a component is as follows: select an atom in the component, with atomic coordinates \( x_1 \). Let \( X = \{ x_1, x_2, x_3, \ldots, x_i \} \) denote the set of atoms to which the first atom is bonded, and which have the same fractional coordinates but in different unit cells, i.e., \( x_i = x_1 + C \cdot h_i \), where \( C \) is the unit cell description and \( h_i \) is an integer vector. Then, the component dimensionality is the rank of the subspace spanned by \( X \):

\[
\text{dim}(X) = \text{rank}(\{x_2 - x_1, x_3 - x_1, \ldots, x_i - x_1\}). \tag{4}
\]

This definition (illustrated in Fig. 8) accommodates both corrugation and thickness. While \( X \) is an infinite set for all but 0D components, \( \text{dim}(X) \) can be determined in a finite number of steps by exploiting the periodicity of a material.

As described above, determination of the dimension of a material requires an analysis of its constituent bonded clusters, or components. To find the dimension of a component, the rank determination algorithm (RDA) of Mounet et al. [5] uses a supercell of fixed size with open boundary conditions. If the supercell is too small, the number of components might be overestimated.

Conversely, the topological scaling algorithm (TSA) of Ashton et al. [6] uses periodic unit cells, which can underestimate the number of components by forming improper connections between them. By improper connections, we mean components which are disconnected in the infinite crystal but are connected due to the periodic cell chosen.

TABLE III. Top ten “most 2D” materials in the ICSD and COD, as ordered by the interval scoring method in Eq. (2). Since the databases contain many layered polymers with large interlayer spacings, we have not included structures containing hydrogen in this list.

<table>
<thead>
<tr>
<th>Source</th>
<th>ID</th>
<th>Compound</th>
<th>2D</th>
<th>0D</th>
<th>02D</th>
<th>1D</th>
<th>2D</th>
<th>3D</th>
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<td>ICSD</td>
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<td>Rb(Au(CF₃SO₃)₄)</td>
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<td>2.313</td>
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<tr>
<td>ICSD</td>
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<td>Sr(CF₃SO₂)₂N₂</td>
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<td>1.968</td>
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<tr>
<td>ICSD</td>
<td>2242431</td>
<td>Cs(N(SO₂CF₃)₂)₂</td>
<td>0.977</td>
<td>0.965</td>
<td>1.968</td>
<td></td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

TABLE IV. Top ten “most 1D” hydrogen-free materials in the ICSD and COD, as ordered by the interval scoring method in Eq. (2).

<table>
<thead>
<tr>
<th>Source</th>
<th>ID</th>
<th>Compound</th>
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<th>k₁</th>
<th>k₂</th>
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The number of components is dependent on the size (along the \( t \) axis) of the periodic cell. In the formulation given in Eq. (5) components for an array of different sizes. In this case, the correct size of the periodic cell is given by lcm \((n, m \mod n)\). In order to avoid improper connections between components, a periodic cell of size \( n \) is needed. This is further complicated for cells containing multiple \( n \)-helices of different sizes. In this case, the correct size of the periodic cell is given by \( \text{lcm}(n_1, n_2, n_3, \ldots) \), where \( n_i \) denotes the number of components in the \( i \)th \( n \)-helix. For example, a structure containing a 5-helix, a 6-helix, and a 7-helix requires a periodic cell of length 210. This cell is so large that it is unlikely that it would be tested using the existing methods.

FIG. 7. Density of states projected onto the 1D and 3D components of the \( \text{S}_2\text{Ta}_3\text{P}_4 \) compound. The energy is relative to the valence band maximum (VBM) and the vertical dashed lines indicate the band edges of the smallest band gap. The electronic spectrum is calculated using the GLLB-SC exchange-correlation functional.

FIG. 8. Cut-out of a periodic corrugated 2D component. The component dimensionality can be found by selecting any atom in the component, and identifying all other atoms in the bonded cluster with the same fractional coordinates, shown here in red. The rank of the subspace spanned by these atoms (here, a plane) determines the dimensionality.

The purpose of the modified method (described in Algorithm 2) is to identify intervals in \( k \) in which the dimensionality classification is constant.
The infinite structure contains 8 components. Any number of repetitions, connections between components.

The input to the algorithm (line 1) is the set of all possible edges, sorted according to their $k$ values, from lowest to highest. Each element in this set, $(k, i, j) \in E$, contains the $k$ value of the edge and the indices, $i$ and $j$, of the vertices it connects. Periodic boundary conditions must be taken into account when generating this set. Due to the periodicity this set is infinitely large; the relevant (finite) subset, however, can be generated incrementally.

The algorithm proceeds by inserting edges from $E$ into the graph, one by one (line 5). For every edge in the primitive cell, the corresponding number of edges are inserted into the supercell. Connected components in both the primitive cell and the supercell are identified after each edge insertion (line 6), from which a component histogram is calculated (line 8). The histogram, $h$, contains the number of 0D, 1D, 2D, and 3D components present. Prior to any edge insertion, only 0D components are present, which is reflected in the initial state of the histogram (line 3). If an edge insertion produces a change in the component histogram (line 9), the $k$-interval is added (line 10) to the set of results (line 4). The algorithm terminates when the histogram consists only a single 3D component (line 12). A 3D interval is added to the results (line 4). The algorithm changes the classification.

The algorithms developed here are included in the ASE library.

Algorithm 1. Pseudocode for calculating component dimensionality.

```python
1: procedure CALCULATEDIMENSIONALITY(E, c)
2:  s := ∅
3:  v := {∅ ∀i ∈ 1…n}
4:  Q := [{c, ∅}]
5:  while Q ≠ ∅ do
6:    if (i, j) ∈ s then
7:      Q := Q \ Q
8:      if rank (v) > rank (v) then
9:        v := v ∪ {j}
10:       end if
11:     end if
12:   for (j, Δ) ∈ E do
13:     if rank (v) > rank (v) then
14:       Q := Q \ Q
15:       if rank (v) > rank (v) then
16:         Q := Q \ Q
17:       end if
18:     end if
19:   end for
20: end while
21: return rank (v)
22: end procedure
```

Algorithm 2. Pseudocode for finding all dimensionality intervals.

```python
1: procedure FINDINTERVALS (E)
2:  kprev := 0
3:  hprev := [n atoms, 0, 0, 0],
4:  R := ∅
5:  for (k, i, j) ∈ E do
6:    if h ≠ hprev then
7:      R := R \ (kprev, k, hprev)
8:    end if
9:    if h = [0, 0, 0] then
10:       return R \ (k, ∞, h)
11:     end if
12:    kprev := k
13:    hprev := h
14: end for
15: end procedure
```
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