An Exploratory Flow Reactor Study of H2S Oxidation at 30-100 Bar

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Published in:
International Journal of Chemical Kinetics

Link to article, DOI:
10.1002/kin.21055

Publication date:
2017

Document Version
Peer reviewed version

Link back to DTU Orbit

Citation (APA):
An Exploratory Flow Reactor Study of H₂S Oxidation at 30–100 Bar

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Received 4 July 2016; revised 30 September 2016; accepted 1 October 2016
DOI 10.1002/kin.21055
Published online 9 November 2016 in Wiley Online Library (wileyonlinelibrary.com).

ABSTRACT: Hydrogen sulfide oxidation experiments were conducted in O₂/N₂ at high pressure (30 and 100 bar) under oxidizing and stoichiometric conditions. Temperatures ranged from 450 to 925 K, with residence times of 3–20 s. Under stoichiometric conditions, the oxidation of H₂S was initiated at 600 K and almost completed at 900 K. Under oxidizing conditions, the onset temperature for reaction was 500–550 K, depending on pressure and residence time, with full oxidation to SO₂ at 550–600 K. Similar results were obtained in quartz and alumina tubes, indicating little influence of surface chemistry. The data were interpreted in terms of a detailed chemical kinetic model. The rate constants for selected reactions, including SH + O₂ ↔ SO₂ + H, were determined from ab initio calculations. Modeling predictions generally overpredicted the temperature for onset of reaction. Calculations were sensitive to reactions of the comparatively unreactive SH radical. Under stoichiometric conditions, the oxidation rate was mostly controlled by the SH + SH branching ratio to form H₂S + S (promoting reaction) and HSSH (terminating). Further work is desirable on the SH + SH recombination and on subsequent reactions in the S₂ subset of the mechanism. Under oxidizing conditions, a high O₂ concentration (augmented by the high pressure) causes the termolecular reaction SH + O₂ + O₂ → HSO + O₃ to become the major consumption step for SH, according to the model. Consequently, calculations become very sensitive to the rate constant and product channels for the H₂S + O₃ reaction, which are currently not well established. © 2016 Wiley Periodicals, Inc. Int J Chem Kinet 49: 37–52, 2017

INTRODUCTION

Hydrogen sulfide (H₂S) is a by-product from cleaning of natural gas and synthesis gas produced from
gasification of coal or biomass, from hydrodesulfurization of light hydrocarbons, and from upgrading heavy oils and coals. It is also released from sulfur-containing fuels during pyrolysis or formed under very reducing conditions in combustion processes. The medium to high-temperature chemistry of H$_2$S has received attention due to its importance in combustion, in the Claus process, and as a potential hydrogen source [1,2]. The H/S system has been investigated for both H$_2$S pyrolysis and H$_2$ sulfidation [3–10]. Despite some remaining uncertainties, available experimental results are described satisfactorily by a detailed reaction mechanism [8].

A wide range of studies have dealt with H$_2$S conversion in the presence of oxygen. Wilson and Hirst [11] discuss atmospheric oxidation of H$_2$S, whereas early investigations of oxidation at elevated temperatures were reviewed by Cullis and Mulcahy [12]. Experimental results have been reported for explosion limits and induction times in static reactors [13–17], oxidation in flow reactors [18], induction times in shock tubes for H$_2$S [19] and H$_2$/H$_2$S mixtures [20], flame speeds [21–24], and structures of premixed [25–32] and diffusion flames [33].

Early modeling efforts for H$_2$S oxidation [19,34] largely had to rely on estimated rate constants for the sulfur subset. More recently, Haynes and coworkers investigated the chemistry of H$_2$S pyrolysis and oxidation in a series of modeling studies [8,18], supported by ab initio calculations for key reactions [35–39]. The mechanism of Zhou et al. [18] was adopted and slightly revised by Mathieu et al. [20] who studied the impact of H$_2$S on ignition of H$_2$ over large pressure (1.6–33 atm) and temperature (1045–1860 K) ranges. Bongartz and Ghoniem [40] used the mechanism of Zhou et al. as a starting mechanism for a more thorough optimization study, using rate constants for 15 reactions in the scheme as parameters when modeling a range of experimental data reported in the literature.

With the exception of the shock tube experiments from Frenklach et al. [19] and Mathieu et al. [20], results for H$_2$S oxidation at elevated pressures are scarce. The objective of the present study is to obtain experimental results for the oxidation of H$_2$S at high pressure (30–100 bar) as a function of temperature (600–925 K) and stoichiometry (lean to stoichiometric) and analyze them in terms of a detailed chemical kinetic model, based on the work of Zhou et al. [18].

**EXPERIMENTAL**

The experimental setup was a laboratory-scale high-pressure laminar flow reactor designed to approximate plug flow. The setup is described in detail elsewhere [41], and only a brief description is provided here. The system was used here for investigation of hydrogen sulfide oxidation at 30 bar and 100 bar pressure, respectively, and temperatures from 450 to 900 K.

The reactions took place in a tubular quartz reactor (inner diameter of 7.5 mm), enclosed in a stainless steel tube that acted as a pressure shell. Using a quartz tube and conducting the experiments at high pressure ensured a minimal contribution from heterogeneous reactions at the reactor wall. However, additional experiments were conducted in an alumina tube (Degussit AL23; inner diameter 6 mm) to assess the importance of surface reactions. The steel tube was placed in a tube oven with three individually controlled electrical heating elements that produced an isothermal reaction zone (±6 K) of 37–47 cm. The temperature profile in the flow reactor was measured by a thermocouple positioned in the void between the quartz/alumina reactor and the steel shell. Results for 30 bar are shown in Fig. 1, whereas the 100-bar profiles are available as the Supporting Information. The system was pressurized from the feed gas cylinders. The reactor pressure was monitored upstream of the reactor by a differential pressure transducer and controlled by a pneumatically actuated pressure-control valve positioned after the reactor. All gases used in the present experiments were high-purity gases or mixtures with certified concentrations. The total flow rate was 2.8 L min$^{-1}$ (STP).

The product analysis was conducted by an online 6890N Agilent Gas Chromatograph (GC-TCD/FID from Agilent Technologies) with an overall relative measurement uncertainty in the range ±2–6%.

![Figure 1](https://via.placeholder.com/150)

**Figure 1** Measured temperature profiles along the reactor axis for 30 bar conditions.
The plug flow assumption was shown by Rasmussen et al. [41] to be a good approximation for the present operating conditions. The uncertainty in the gas temperature due to the effect of heat release from combustion was limited by a high level of dilution.

DETAILED KINETIC MODEL

The detailed chemical kinetic model for H$_2$S oxidation was adopted mostly from the recent study on H$_2$S oxidation by Zhou et al. [18]. The reaction mechanism consists of a H$_2$ subset [42] and a full description of the H/S/O reaction system [8,18,43–47].

Table I shows thermodynamic properties of selected species. Most of the data are drawn from the database of Goos et al. [48]. The heats of formation of key species such as SH, HSO, and HOS are in agreement within the uncertainty with recent high-accuracy theoretical results [51–53].

Selected reactions from the sulfur subset are listed in Table II. Below selected reactions are discussed in some detail, with emphasis on steps that may have a particular significance under the high-pressure conditions of the present study.

Most of the reactions of H$_2$S with the radical pool have been characterized experimentally and theoretically over a wider temperature range. For H$_2$S + H (R2), H$_2$S + O (R3, R4), and H$_2$S + S (R12), we rely on rate constants determined by Marshall and coworkers [54,55,58]. The rate constant for H$_2$S + OH (R5) was taken from the theoretical study by Elling-son and Truhlar [56]; it provides an explanation for the unusual temperature dependence of the reaction, and the calculated value is in good agreement with experiment [64–68]. The H$_2$S + O reaction has two product channels, SH + OH (R3) and HSO + H (R4). The branching fraction $k_3/(k_3 + k_4)$ calculated by Goumri et al. [55] agrees with the reported low-temperature upper limit by Singleton and Cvetanovic [69] and, within the uncertainty, with the high-temperature determination by Tsuchiya et al. [70].

Other consumption steps for H$_2$S have not been characterized experimentally. For the reaction of H$_2$S with HO$_2$,

$$H_2S + HO_2 \rightleftharpoons HS + H_2O_2 \quad (R6b)$$

$$H_2S + HO_2 \rightleftharpoons HSO + H_2O \quad (R7)$$

only a room temperature upper limit has been reported [71]. We have adopted the values for $k_6$ and $k_7$, calculated from transition state theory by Zhou et al. [18]. The rate constants for H$_2$S with O$_2$ and SO$_2$ were also drawn from theoretical work of Haynes and coworkers [18,35,36]. For H$_2$S + O$_2$, we considered in the present work also the spin-forbidden formation of singlet H$_2$SOO peroxyde, followed by the barrierless decomposition path to HSO + OH identified by Motoya et al. [35]. However, the reaction barrier for the initial step is too high for it to be important under the current conditions. Another possibility, as suggested by Starik et al. [72], is that reaction of H$_2$S with the tiny equilibrium population of singlet oxygen may promote...
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The low pressure limit is $1.0254 \times 10^4$ cm s$^{-1}$.
the initiation chemistry; however, this was not investigated in the present work.

The reaction of SH with O₂ is a key step in the oxidation of H₂S at high temperature [62]. It has a significant barrier and attempts to measure the rate coefficient at low temperatures have yielded only upper limits, with a value of 2 × 10⁵ cm³ mol⁻¹ s⁻¹ at 298 K from the study of Stachnik and Molina [73] considered to be the most reliable [59]. The reaction, which has been studied theoretically by a number of groups [38,50,55,74–76], has several proposed product channels:

\[
\text{SH} + \text{O}_2 \rightarrow \text{HSO} + \text{O} \quad (\text{R25})
\]

\[
\text{SH} + \text{O}_2 \rightarrow \text{S} + \text{HO}_2 \quad (\text{R26})
\]

\[
\text{SH} + \text{O}_2 \rightarrow \text{SO} + \text{OH} \quad (\text{R27})
\]

\[
\text{SH} + \text{O}_2 \rightarrow \text{SO}_2 + \text{H} \quad (\text{R28})
\]

The theoretical work of Zhou et al. [38] indicated that the reaction proceeds mainly via a four-membered cyclic transition state to form SO + OH (R27) at temperatures below 1000 K, whereas HSO + O (R25) becomes the major product channel above this temperature. However, Garrido et al. [75] identified a new and faster reaction path to form SO₂ + H (R26) via a three-center ring structure.

The existence of the new pathway is supported by the recent theoretical study of Freitas et al. [76]. They characterized a transition state leading from HSOO via an electronically excited state to HSO₂, with a CCSD(T) energy extrapolated to the complete basis set limit of 19.8 kcal mol⁻¹ above HSOO. Assuming this is the bottleneck, SH + O₂ collisions will maintain a small equilibrium population of HSOO, which lies 6.4 kcal mol⁻¹ below SH + O₂ [77]. Accordingly, we evaluated the bimolecular TST for SH + O₂ → products via this TS with a barrier of 13.4 kcal mol⁻¹. Because HSO₂ is formed with energies \( E \) of about 82 kcal mol⁻¹ or more, well above the S–H dissociation threshold \( E₀ \) of about 21 kcal mol⁻¹ [75], we assume the major products are H + SO₂ (see Table II). We justify this via evaluation of the microcanonical TST result \( k(E) = G'(E - E₀)/(h N(E)) \) for first-order dissociation of nascent vibrationally excited HSO₂. \( N(E) \) is the density of states for HSO₂, and \( G'(E - E₀) \) is the sum of states for the TS for HSO₂ dissociation to H + SO₂. This yields \( k(E) = 1.6 \times 10^{14} \) s⁻¹, three orders of magnitude faster than collisional stabilization at 100 atm pressure.

Figure 2 shows an Arrhenius plot for the SH + O₂ reaction. Tsuchiya et al. [62] derived the rate constant from measured concentration profiles of H and O atoms in flash-photolysis/shock tube experiments of H₂S/O₂/Ar mixtures. They could not identify the major product channel with certainty, but in their analysis they assumed formation of HSO + O. The theoretical values derived by Zhou et al. [38] and Garrido et al. [75] are all consistent with the room temperature upper limit by Stachnik and Molina [73] but also

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²Duplicate reaction – the resulting rate constant is the sum of the two expressions.
³Calculated from TST.
⁴Reaction on the triplet surface.
⁵Reevaluation of the PES indicates that the reaction proceeds without a barrier.
⁶The rate constant for (R37) from Zhou [94] was obtained for quartet transition state. For completeness, we did TST calculations also via the doublet TS.
⁷The high-pressure limit A factor was reduced from the value of 3.5E12 calculated by Zhou et al. [39].
well below the values derived by Tsuchiya et al. The present rate constant for the dominating product channel to SO$_2$ + H is almost an order of magnitude higher than the value of Garrido et al., used in the modeling study of Zhou et al. [18].

The apparent agreement of the present rate constant with the values of Tsuchiya et al. is fortuitous since they assumed HSO + O to be the major product channel. Actually, Tsuchiya et al. concluded based on their measurements that SO$_2$ + H was unlikely to be the major product channel for SH + O$_2$. To resolve this issue, we used the present kinetic model to reinterpret selected data from Tsuchiya et al. (Fig. 3). The comparison shows that the present rate constants and product channels for SH + O$_2$ are consistent with their measurements. In fact, the present model predictions of H and O show only a limited sensitivity to the rate constant for the SH + O$_2$ reaction.

At the low to medium temperature, high-pressure conditions of the present work, it is of interest whether an adduct formed from SH + O$_2$ may have a sufficient lifetime to react. Theoretical studies [38,50,55,75,76] indicate that the thylperoxyl radical (HSOO) is formed from recombination of SH and O$_2$.

\[
\text{SH} + \text{O}_2(+\text{M}) \rightleftharpoons \text{HSOO}(+\text{M}) \quad (\text{R29})
\]

whereas formation of HSO$_2$ or HOSO is inaccessible.

Turnipseed et al. [78] proposed that reaction (R29) could be partially equilibrated in the atmosphere, with subsequent reactions of with HSOO promoting oxidation of SH. Goumri et al. [50] and Zhou et al. [38] found (R29) to be essentially barrierless, whereas Resende and Ornellas [74] and Ballester and Varandas [79] found barriers of 12 and 8 kcal mol$^{-1}$, respectively. The difference may partly be attributed to the different geometries of HSOO that are separated by only about 1 kcal mol$^{-1}$ in energy [75]. The recent study by Garrido et al. [75] calculated a 7 kcal mol$^{-1}$ barrier to formation of cis-HSOO, whereas formation of a skewed HSOO isomer proceeds without a barrier.

The HSOO adduct is weakly bound, and the reaction is rapidly equilibrated. This means that the concentration of HSOO will remain very low, even at the pressures of the present study. Dissociation of HSOO to HSO + O has a significant barrier and is not competitive. However, a fast reaction of HSOO and O$_2$ (the only abundant reactant under our conditions) could possibly be important,

\[
\text{HSOO} + \text{O}_2 \rightleftharpoons \text{HSO} + \text{O}_3
\]

There are no data for this step in the literature, but the reverse reaction, HSO + O$_3$, has been studied experimentally at low temperatures [63,80–82]. The reaction apparently has two major product channels,

\[
\text{HSO} + \text{O}_3 \rightleftharpoons \text{SH} + \text{O}_2 + \text{O}_2
\]
\[
\text{HSO} + O_3 \rightleftharpoons \text{HSO}_2 + O_2, \quad \text{(R41)}
\]

with a branching fraction reported to be roughly 50\% at 298 K [59]. Conceivably, the \( \text{SH} + O_2 + O_3 \) channel (R40) is a sequence of the two steps,

\[
\text{HSO} + O_3 \rightleftharpoons \text{HSOO} + O_2
\]

since \( \text{HSOO} \) decomposes readily even at 298 K. Taking both steps in the reverse direction, we get the sequence \( \text{SH} \rightleftharpoons \text{HSOO} \rightleftharpoons \text{HSO} + O_3 \), which could conceivably be important under the conditions of our work. In the reaction mechanism, we have adopted the one-step reaction (R40), as recommended by Atkinson et al. [59]. However, owing to the fast equilibration of (R29), modeling predictions are not sensitive to whether (R40) is put in as a single reaction or divided into two steps.

Consequently, \( \text{SH} + O_2 + O_2 \rightarrow \text{HSO} + O_3 \) (R40b) may be a source of ozone under the conditions of the present work. Ozone is much more reactive than molecular oxygen and may promote reaction. Following other recent studies on low-temperature sulfur chemistry [83,84], we include an \( O_3 \) reaction subset with rate constants mostly from Atkinson et al. [59]. Ozone may interact with the O/H radical pool or react with sulfur species, primarily \( \text{H}_2\text{S} \) or \( \text{SH} \). The \( \text{H}_2\text{S} + O_3 \) reaction has been studied both experimentally [85–91] and theoretically [57]. Experimental work [88–90] indicates that the overall reaction is complex, involving a radical-forming initial step followed, after an induction time, by a free radical mechanism. Reaction orders of 0–0.5 in \( \text{H}_2\text{S} \) and 1.5 in \( O_3 \) have been reported [86,87,89,90]. The reaction appears to be sensitive to surfaces, and reported results show a considerable scatter.

The most reliable measurement of \( \text{H}_2\text{S} + O_3 \) is believed to be the room temperature upper limit by Becker et al. [88]. The recent theoretical study by Mousavipour et al. [57] indicates that \( \text{SO}_2 + \text{H}_2\text{O} \) (R10) is the dominating product channel for \( \text{H}_2\text{S} + O_3 \), along with a (very) minor channel forming \( \text{HOSO} + OH \) (R11). We have tentatively adopted the rate constants from Mousavipour et al., even though the suggestion of (R10) as the dominant channel seems to be inconsistent with the experimental observations.

The rate constant for \( \text{SH} + O_3 \) (R30) has been measured at low temperature [80,82,92], and values are in good agreement, indicating a fast reaction with a low activation energy. We have adopted the recommendation by Atkinson et al. [59]. The recent theoretical study by Resende and Ornellas [93] indicates a significantly higher barrier to reaction than derived from experiment. However, this is contradicted by our own preliminary analysis and more work is desirable on this reaction.

Owing to the low reactivity of the \( \text{SH} \) radical, the \( \text{SH} + \text{SH} \) reaction becomes important, even at oxidizing conditions. This reaction has two major product channels:

\[
\text{SH} + \text{SH} \rightleftharpoons \text{H}_2\text{S} + \text{S} \quad \text{(R12b)}
\]

\[
\text{SH} + \text{SH}(+\text{M}) \rightleftharpoons \text{HSSH}(+\text{M}) \quad \text{(R43)}
\]

The rate constant for the \( \text{H}_2\text{S} + \text{S} \) reaction (R12), which is believed to be largely independent of pressure, was taken from the combined experimental and theoretical study of Gao et al. [58]. The high- and low-pressure limits, as well as the falloff behavior, for the \( \text{SH} + \text{SH} \) recombination step (R43) were initially drawn from the RRKM study of Zhou et al. [39]. However, for this reaction rate parameters are more uncertain, since extrapolation is required. In the present work, we modified the high-pressure limit \( k_{43,\infty} \) to improve agreement with experiment under stoichiometric conditions. This is discussed in more detail below. The reaction feeds into the \( S_2\text{H}_3 \) subset, which was drawn largely from Haynes and coworkers [8,39].

Reactions of sulfur radicals with \( O_2 \) can be important for the generation of chain carriers, even under conditions with fairly low concentrations of oxygen. In addition to \( \text{SH} + O_3 \) discussed above, reactions of \( \text{HOSO} + \text{S}, \text{HSS} \) with \( O_3 \) should be considered. The rate constants for the chain branching steps \( \text{SO}_2 + O_2 \) and \( \text{S} + O_2 \) are well established [83,84], but values for \( \text{HSO} + O_2 \) and \( \text{HSS} + O_2 \) are more uncertain.

For \( \text{HSO} + O_2 \), the only reliable measurement is a room temperature upper limit of \( 1.2 \times 10^7 \text{ cm}^3\text{ mol}^{-1}\text{ s}^{-1} \) from Lovejoy et al. [81]. The reaction has several possible product channels,

\[
\text{HSO} + O_2 \rightleftharpoons \text{SO} + \text{HO}_2 \quad \text{(R37)}
\]

\[
\text{HSO} + O_2 \rightleftharpoons \text{SO}_2 + \text{OH} \quad \text{(R38)}
\]

\[
\text{HSO} + O_2 \rightleftharpoons \text{H} + \text{HSO}_2 + \text{O} \quad \text{(R39)}
\]

According to theoretical work, the fastest channel is (R38) [18], whereas (R39) is quite slow [46]. However, also (R37) has a significant impact on the system,
because the subsequent reaction of SO with O2 leads to formation of atomic oxygen.

Prior investigations of HSO chemistry with O2 relied on G3 energies at CASSCF geometries [94]. Investigation of the HSO + O2 transition states using CBS-QB3 theory leads to differences in the rate constants of up to a factor of 10. This is consistent with the target energy accuracies of G3 and CBS-QB3 theory for stable species, and we provisionally suggest uncertainties in these reaction barriers of ca. 2.5 kcal mol$^{-1}$, which corresponds roughly to an order of magnitude in the rate constant.

For HSS + O2, the original energy calculations were based on MP2/6-31G(d) geometries [18]. We reevaluated the TSs using B3LYP/6-311G(2d,d,p) theory, followed by CBS-QB3 for single-point energies. Based on this, for the doublet TS for HSS + O2 → HSO + SO, we get a barrier of 39.2 kcal mol$^{-1}$, and for the two quartet TSs that we attribute to HSS + O2 → S2 + HO2, we get 18.1 and 10.7 kcal mol$^{-1}$, respectively. In addition, we identified a new doublet TS for HSS + O2 → S2 + HO2 with a barrier of 8.7 kcal mol$^{-1}$. For this lowest barrier pathway, our TST calculation (including Eckart tunneling and a hindered rotor) yields a rate constant which is about an order of magnitude faster than the previous estimate.

RESULTS AND DISCUSSION

Experiments for H2S oxidation with high dilution in N2 (or CO2) as a function of temperature from 450 to 900 K were conducted under stoichiometric and oxidizing conditions. Table III lists the experimental conditions. The fuel–air equivalence ratio $\phi$ ranged from 0.9 to 0.03. In the following, the experimental results are compared with modeling predictions. Calculations shown in the figures, conducted using the CHEMKIN PRO software package [95], were restricted to the isothermal zone. Simulations with the full measured temperature profile were similar.

Figure 4 shows results for H2S oxidation under stoichiometric conditions at 30 bar. Experimental data are shown as symbols, modeling predictions as solid lines. The dashed line denotes predictions with the original value of $k_{13,\infty}$ (SH + SH (+M) ⇌ HSSH (+M)) from Zhou et al. [39]. Inlet composition: 756 ppm H2S, 1290 ppm O2, balance N2 ($\phi = 0.88$, black symbols) or 750 ppm H2S, 1190 ppm O2, balance CO2 ($\phi = 0.94$, smaller red symbols). The residence time in the isothermal zone is calculated from $\tau (s) = 3520/T (K)$.
Figure 5  Pathway diagram for H\textsubscript{2}S under stoichiometric conditions (corresponding to Fig. 4) and 700 K.

conditions. This is expected, since interaction of CO\textsubscript{2} with the O/H radical pool [96] or with sulfur radicals would be expected to occur only at higher temperatures than those of the present study.

Two sets of modeling predictions are shown. The solid lines denote calculations with the basis mechanism. In this mechanism, we reduced the $A$ factor of the high-pressure limit for SH + SH ($+$M) $\rightleftharpoons$ HSSH ($+$M) (R43) by a factor of four to improve agreement with experiment. This change is considered to be within the uncertainty in the high-pressure limit, for which there are no experimental results. The dashed line shows predictions using the value of $k_{43,\infty}$ calculated by Zhou et al. [39], resulting in a shift upwards in the temperature for onset of reaction of about 100 K.

The modification of $k_{43,\infty}$ allows a satisfactory prediction of the temperature for 50\% conversion of H\textsubscript{2}S. Still, important differences are seen when comparing experimental results and modeling predictions. Most importantly, the temperature for onset of reaction is overpredicted; the calculations indicate a steeper gradient in the H\textsubscript{2}S concentration profile than observed.

Figure 5 shows a reaction path diagram for H\textsubscript{2}S oxidation at 700 K and stoichiometric conditions, corresponding to Fig. 4. Hydrogen sulfide is consumed by reaction with the O/H radical pool (R3, R5) to form SH. The SH radical is largely consumed by its self-reaction, forming either H\textsubscript{2}S + S (R12b) or HSSH (R43). The atomic S is oxidized to SO\textsubscript{2} in the chain-branching sequence S + O\textsubscript{2} $\rightarrow$ SO + O (R34), SO + O\textsubscript{2} $\rightarrow$ SO\textsubscript{2} + O (R36). HSSH, on the other hand, is oxidized to S\textsubscript{2} through the sequence HSSH + SH $\rightarrow$ HSS + H\textsubscript{2}S (R45), HSS + SH $\rightarrow$ S\textsubscript{2} + H\textsubscript{2}S (R49), and HSS + O\textsubscript{2} $\rightarrow$ S\textsubscript{2} + HO\textsubscript{2} (R47). According to the current calculations, formation of oxidized sulfur species from the H\textsubscript{2}S\textsubscript{2} is very limited under these conditions.

Figure 6 shows sensitivity coefficients for H\textsubscript{2}S under the same conditions (stoichiometric, 700 K). The analysis shows that the predicted H\textsubscript{2}S concentration is mainly sensitive to the branching fraction for the SH + SH reaction. Formation of H\textsubscript{2}S + S (R12b) strongly promotes the oxidation rate due to the subsequent chain-branching reactions of S and SO with O\textsubscript{2} (R34, R36), whereas formation of HSSH (R43) is chain terminating. None of the reactions in the S\textsubscript{2} subset is seen to promote oxidation; predictions, however, are sensitive to the terminating steps H\textsubscript{2}S + S (R44) and HSS + SH (R49). We attribute the discrepancy between experimental results and modeling predictions mostly to uncertainties in the S\textsubscript{2} chemistry.

Figure 7 shows results under oxidizing condition at 30 bar pressure. The initiation temperature is 520 K, approximately 100 K lower than at stoichiometric conditions. Contrary to the behavior in Fig. 4, the concentration gradient of H\textsubscript{2}S with respect to temperature is steep and full oxidation to SO\textsubscript{2} is obtained already at 575 K. This indicates that a strong chain-branching mechanism is active already at low temperature.

The model (solid lines) predicts a significantly higher temperature for onset of reaction, about 625 K. However, the concentration profiles are similar to those observed experimentally, just shifted 100 K to higher temperatures. The dashed line shows predictions conducted with the H\textsubscript{2}S + O\textsubscript{3} reaction (R10, R11) omitted. Reaction (R10), which is the dominating product channel, serves in effect as a chain-terminating step because it takes out reactive ozone without providing chain carriers. Omission of this step serves to bring the predicted onset temperature in closer accordance with observations. However, this change causes a smaller gradient in H\textsubscript{2}S than observed above the onset temperature and leads to an underestimation of SO\textsubscript{2}, caused by prediction of significant amounts of SO\textsubscript{3}.

Figure 8 shows results under oxidizing condition at 100 bar. The larger black symbols represent experimental results conducted in a quartz tube, whereas the smaller red symbols denote data obtained in an
alumina tube. The hydrogen sulfide oxidation is initiated at around 475 K and completed at 550 K. The earlier onset, compared to data at 30 bar, is partly caused by a longer residence time at 100 bar. The behavior is similar to that at 30 bar, except that the initiation temperature is shifted to lower values.

Despite some deviation in residence time, the experimental results from the alumina tube agree well with those of the quartz tube. This is an indication that surface effects, if present, are of minor significance under the conditions with high pressure. This issue is discussed further below.

The modeling results show deviations similar to those at 30 bar. Again, the full model overpredicts the onset temperature. Omission of (R10) serves to reduce the predicted onset temperature significantly but reduces also the H$_2$S concentration gradient and the SO$_2$ level.

Figure 9 shows a reaction path diagram for H$_2$S oxidation at 600 K and oxidizing conditions, corresponding to Fig. 8. Oxidation paths are very different from those under stoichiometric conditions. A major fraction of the hydrogen sulfide is consumed by reaction with ozone (R10) to form SO$_2$ directly. A similar amount reacts with O/H radicals (R5, R6b) to form SH, which is consumed by reaction with two oxygen...
molecules (R40b), yielding HSO and O. The HSO reacts mostly with O to form SO. Owing to the high oxygen concentrations, S species are only formed in minor amounts.

Figure 10 shows sensitivity coefficients for H2S under lean conditions (100 bar, 600 K). The analysis shows that consumption of H2S is promoted by chain-propagating or -branching reactions, primarily H2S + HO (R6b), HSO + O (R38), and H2S + O (R4). Even though it consumes H2S, the H2S + O → SO2 + H2O (R10) is a major inhibiting step because it forms stable products. It competes with the minor radical producing channel HOSO + OH (R11), which has a negative sensitivity coefficient.

It is known that oxidation of H2S is sensitive to surface reactions, and it is important to assess whether heterogeneous effects can explain some of the differences between measurements and modeling predictions under the present conditions. Adesina et al. [9] report that for thermolysis of hydrogen sulfide (i.e., under oxygen-free conditions) there are no indications of surface effects on quartz at 1073 K. However, under oxidizing conditions significant heterogeneous effects have been reported. An increase of the surface to volume ratio in glass reactors serves to inhibit H2S oxidation and explosion at low pressure and temperatures of 500–600 K [13,14]. Under these conditions, the main impact of the surface seems to be to remove chain carriers, while surface initiation involving a reaction between H2S and O is less important. More recently, Zhou et al. [18,94,100] investigated the impact of surface effects on H2S oxidation in quartz flow reactors at atmospheric pressure. They found that a 30-fold increase in the surface area of the reactor slightly enhanced H2S consumption in the 650–950 K range, while it inhibited formation of H2S [94,100]. With a B2O3 coating of the quartz surface, the system became much less reactive [18,94,100].

Figure 11 compares modeling predictions with the experimental data of Zhou et al. [18], obtained with
Figure 12  Ignition delay time of H$_2$S/air (4–22% H$_2$S, 16–20% O$_2$) and H$_2$S/H$_2$O/air (4–21% H$_2$S, 2–13% H$_2$O) mixtures versus temperature. Pressure was varied between 29 and 47 atm, whereas the equivalence ratio $\phi$ changed from 0.3 to 2. Open symbols present experimental data from Frenklach et al. [19], whereas the closed symbols are the results of the present simulations.

and without coating of the quartz reactor surface. The pressure is lower, but the temperature range and composition are comparable to those of the present work. The figure illustrates the considerable impact of the surface condition. Two sets of predictions are shown, conducted with the present model (solid lines) as well as with the basis mechanism of Zhou et al. [18] (dashed lines). Under these conditions, the two mechanisms result in very similar concentration profiles for H$_2$S. Both models predict the sharp onset of reaction observed in the coated reactor, but the calculated onset temperature is closer to that reported for the uncoated reactor.

Surfaces may initiate reaction and promote oxidation by catalyzing fuel conversion or inhibit oxidation by acting as a sink for radicals. The difference between the impact of the quartz surface between the batch and flow reactor experiments may be attributed to differences in temperature and pressure that could change the balance between surface initiation/oxidation and loss of radicals on the wall. Based on the available results, we believe that the impact of heterogeneous reactions in the high-pressure quartz reactor used in the present work is limited. This is supported by the agreement between results obtained in the quartz and alumina reactors.

To extend the evaluation of the present model, predictions are compared also to selected data obtained at high temperature and pressure in shock tubes. Figure 12 shows data from Frenklach et al. [19] on ignition delays of H$_2$S in air, with and without presence of water vapor. The experiments were conducted at 34 bar and temperatures in the range 950–1200 K. The present model underestimates the ignition delays by factors of 2–6, most pronounced under dry conditions and at high temperatures.

Figure 13 shows the results of a sensitivity analysis for the ignition delay time under typical conditions (12% H$_2$S in air, 38 bar, and 1060 K). The sensitivity coefficients are calculated as $(\Delta \tau / \tau) / \Delta k / k$, so a negative coefficient indicates a promoting effect. Similar to the flow reactor results, the modeling predictions are very sensitive to the branching fraction for the SH + SH reaction. Formation of H$_2$S + S (R12b) strongly promotes the oxidation rate due to the subsequent chain-branching reactions of S and SO with O$_2$ (R34, R36) while formation of HSSH (R43) is chain terminating. Also a number of the reactions in the S$_2$ subset, such as the terminating steps H$_2$S + S (R44) and HSS + SH (R49), show significant sensitivity coefficients. No attempt was made to improve modeling predictions for these conditions; further work is desirable to reduce uncertainties in the S$_2$ chemistry subset.

Figure 14 compares modeling predictions with ignition delay data from Mathieu et al. [20] for mixtures of H$_2$ and H$_2$S, obtained at pressures of about 34 atm and 1045–1860 K. Under these conditions, the radical pool is determined to a larger extent by the H$_2$/O$_2$ subset of the mechanism, and the agreement with experiment is better than for the conditions of Frenklach et al.
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Figure 13  Sensitivity of ignition delay time to the reaction rate constants (corresponding to Fig. 8). Results are shown for 12% H$_2$S mixed with air ($\phi = 0.43$) at 38 bar and 1060 K.

(Fig. 12). For low concentrations of H$_2$S (100 and 400 ppm), the predictions compare well to the measurements, but the model systematically underpredicts the ignition delays for higher concentrations of H$_2$S (1600 ppm).

CONCLUSIONS

Hydrogen sulfide oxidation experiments were conducted at high pressure (30 and 100 bar) under oxidizing and stoichiometric conditions, respectively, and temperatures ranging from 450 to 925 K. The H$_2$S oxidation behavior depended strongly on the stoichiometry. Under stoichiometric conditions, the oxidation of H$_2$S was initiated at 600 K, with the consumption rate increasing only slowly with temperature up to 900 K. Under oxidizing conditions, the onset temperature for reaction was 500–550 K, depending on pressure and residence time, with a steep gradient in H$_2$S above this temperature. The data were interpreted in terms of a detailed chemical kinetic model. The rate constants for selected reactions, including SH + O$_2$ ⇄ SO$_2$ + H, were determined from ab initio calculations. Modeling predictions generally overpredicted the temperature for onset of reaction. Calculations were sensitive to reactions of the comparatively unreactive SH radical. Under stoichiometric conditions, the oxidation rate was mostly controlled by the SH + SH branching ratio to form H$_2$S + S (promoting reaction) and HSSH (terminating). Further work is desirable on the SH + SH recombination and on subsequent reactions in the S$_2$ subset of the mechanism. Under oxidizing conditions, a high O$_2$ concentration (augmented by the high pressure) causes the reaction SH + O$_2$ + O$_3$ → HSO + O$_3$ to become the major consumption step for SH,
Figure 14 Ignition delay time of $\text{H}_2\text{S}/\text{H}_2\text{O}_2/\text{Ar}$ mixtures ($\phi = 0.50–0.58$) versus temperature. The $\text{H}_2\text{S}$ concentrations are 100, 400, and 1600 ppm, respectively, with 1% $\text{H}_2$ and 1% $\text{O}_2$. Symbols denote experimental data from Mathieu et al. [20], whereas lines represent simulations conducted at the average pressure of 34.1 atm.

According to the model. Consequently, calculations become very sensitive to the rate constant and product channels for the $\text{H}_2\text{S} + \text{O}_3$ reaction, which are currently not well established.

The work is part of the CHEC (Combustion and Harmful Emission Control) research program at DTU Chemical Engineering. YS wishes to acknowledge funding from CSC (China Scholarship Council). PM thanks the R. A. Welch Foundation (grant B-1174) for support.

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