Conductance switching in a molecular device: The role of side groups and intermolecular interactions

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An important goal in the study of molecular electronics is to identify molecules that can be combined to perform logical functions.1 In particular, Reed and co-workers2–5 have studied the electrical properties of a set of phenyl-ethylene oligomers [known as Tour wires (TW’s)] functionalized with different side groups, and demonstrated that such molecules can show negative differential resistance (NDR)2,3 and a “molecular memory”4,5 effect, in which case molecules can be switched from a low conductance state to a high conductance state by application of a voltage pulse. A fundamental understanding of the microscopic mechanisms governing the NDR and memory effects in these systems is still lacking.

Complementary studies, by Weiss and co-workers,7 of isolated or small bundles of TW’s embedded in a monolayer of spacer alkyl molecules, demonstrated spontaneous conductance switching. While Reed and co-workers only observed switching effects for TW’s with NO2 side groups, Weiss and co-workers reported spontaneous switching of all the molecules in their study but had limited success in inducing switching by using a voltage pulse.7

Theoretical studies of functionalized TW’s, have so far been restricted to isolated molecules. Based on such studies, Seminario et al. proposed that charging of the molecule and subsequent localization/delocalization of molecular orbitals is the microscopic mechanism behind NDR,8,9 while Cornil et al. have found that bias-induced alignment of molecular orbitals on the first and last phenyl rings of twisted TW’s can lead to NDR.10

In this paper we present the first studies of the electrical properties of functionalized TW’s covalently bound to gold surfaces. Most prominent previous studies of related systems include the IV characteristics of a functionalized benzene ring13 and a bare TW.14 Our calculations are based on the TRANSIESTA package,11 which is a nonequilibrium Greens-function-based electronic transport program founded on density-functional theory.12 It treats the entire system self-consistently under finite bias conditions.

The nanopore structures in which NDR and memory effects have been observed are formed by evaporating the molecules onto a small Au surface inside a Si3N4 structure, and then evaporating another Au electrode onto the molecular monolayer.2 There is no detailed information about the structure of the monolayer, and the purpose of this study is not an accurate numerical model of the nanopore experiment. We will form an idealized model of an electrode/monolayer/electrode system and use the insight from our studies to make general statements about the effect of functionalization on the electron-transport properties of TW’s. We will show that while DeVentra et al. found intrinsic NDR for a functionalized benzene ring, the ground-state geometry of a functionalized TW does not show any NDR, and has an IV characteristics similar to the unfunctionalized TW. Instead, our calculations show that functionalization stabilizes twisted conformations within the monolayer, and we propose that NDR and memory effects are related to such conformational changes.

We use an Au(111) surface in a 3×3 unit cell for the electrodes and assume that the TW’s are chemisorbed to the surfaces through strong thiol bonds,15 as illustrated in Fig. 1. Monolayer A consists of a bare TW with thiol end groups, while monolayers B and C have been functionalized with NH2 and NO2 side groups, respectively.

We first investigate the effect of the side groups on the electrical properties of the monolayers. The current through the monolayers is determined by the quantum-mechanical probability for electrons to tunnel from one electrode to the other, and is calculated using the Landauer formula16

$$I = \int \frac{\mu_B}{\mu_L} T(E, V_b) dE,$$

where \(\mu_{L/R} = \pm e V_b\) and \(T(E, V_b)\) is the transmission probability for electrons incident at an energy \(E\) through a device under a potential bias \(V_b\) (\(n.b \ V_b \geq 0\) corresponds to hole injection from the right electrode). The general shape of the zero-bias transmission spectra, shown as \(T(E, V_b=0V)\) in Fig. 2(a), is similar for all three systems. The zero-bias conductance \(G = (e^2/h) T(E = \mu_L, V_b = 0)\), given by \(2.0 \mu S, 2.3 \mu S, \text{ and } 1.9 \mu S\), for monolayers A, B, and C, respectively, is dominated by the tail of a broad highest-occupied-molecular-orbital (HOMO) resonance (at \(E_{\text{HOMO}} = -0.84, -0.65, \text{ and } -0.85\) eV for monolayers A, B, and C). The narrower lowest-unoccupied-molecular-orbital (LUMO) resonance (at \(E_{\text{LUMO}} = 1.56, 1.59, \text{ and } 1.14\) eV for A, B, and C respectively) contributes less to the conductance, and thus holes are the dominant carriers. The
NH\textsubscript{2} group is electron donating, so that $E_{\text{HOMO,B}} - E_{\text{HOMO,A}} \approx 200$ meV while the NO\textsubscript{2} group is electron accepting so $E_{\text{LUMO,C}} - E_{\text{LUMO,A}} \approx -400$ meV.

To calculate the $I$-$V_b$ spectrum we performed self-consistent calculations for biases in the range $-2.8$ to $2.8$ V in steps of $0.2$ V. We note that the charge on the molecule is not fixed\textsuperscript{11} and adjusts itself to minimize the free energy\textsuperscript{17} as the left/right electrochemical potentials are changed. For all three systems we find that the molecules are close to charge neutrality, and the charge on the functionalized molecules $Q_M$, as determined by a Mulliken population analysis, change by less than $0.03e$ as the bias voltage $V_b$ is varied.\textsuperscript{18} The $I$-$V_b$ spectra, shown in Fig. 2(b), are very similar, increasing slowly at first but then increasing rapidly around $2$ V where the resonances come into alignment with the bias window. The main effect of $V_b$ on $T(E,V_b)$ is to sample more and more of the resonance while resonant peak height and position vary slowly with $V_b$.\textsuperscript{14}

When the molecules interact with the Au(111) electrodes, the molecular levels broaden into a continuum.\textsuperscript{16} The eigenstates of the whole metal/monolayer/metal system consist of scattering states,\textsuperscript{19,20} which are molecular-orbital-like in the molecule, and Bloch-wave-like in the metal slabs. If an orbital is delocalized across the molecule, an electron that enters the molecule at the energy of the orbital has a high probability of reaching the other end, and thus there is a corresponding peak in the transmission probability $T(E,V_b)$, as illustrated in Fig. 2(a). By calculating the continuum eigenstates at the resonance energies ($E_{\text{HOMO}},E_{\text{LUMO}}$), the orbitals that are responsible for current flow through the molecule can be analyzed. The HOMO resonances, illustrated in Fig. 2(c), resemble the HOMO of the isolated molecules while the narrower LUMO resonances, illustrated in Fig. 2(d), resemble the LUMO of the isolated molecules. There are minor differences between the orbitals: the NH\textsubscript{2} side group participates in the HOMO resonance but not in the LUMO resonance, while the NO\textsubscript{2} group participates in the LUMO resonance and not the HOMO resonance, consistent with their respective donating/accepting characters. It is, however, clear from both the shape of the transmission curves in Fig. 2(a) and the orbitals in Figs. 2(c) and 2(d) that the resonant transmission peaks in monolayers $A$, $B$, and $C$ are related to the delocalized nature of the $\pi$ orbitals of the bare TW’s and is not strongly affected by the functionalization.

The energetics of the monolayers are, however, strongly affected by the functionalization. The triple bond between the phenyl rings is rotationally symmetric and thus constitutes an easy axis about which the middle ring can rotate. The differences in total energy as a function of the rotation angle of the middle ring, $\theta$, for the molecules arranged in the $3\times3$ unit-cell of Au(111) are shown in Fig. 3(a). The inter-molecular interaction energy is strongly dependent on the functionalization: there is a cost of 60 meV (180 meV) to rotate the ring by $90^\circ$ in monolayer $A$ ($B$), and the flat molecule is clearly favored. Monolayer $C$ shows a quite different behavior, it acquires a local minimum at $\theta=60^\circ$ and $120^\circ$. Separable calculations show that if we double the distance between the molecules these minima disappear, thus they must be related to intermolecular interactions. Figure 3(b) shows the effective potential within the monolayer, and it reveals the formation of a hydrogen bond between the NO\textsubscript{2} group and a hydrogen atom on the neighboring TW. Similar hydrogen bond formation involving NO\textsubscript{2} groups has also been observed in other molecular layers.\textsuperscript{25} From Fig. 3(b) it is also clear that the minima at $60^\circ$ and $120^\circ$ can be attributed to bond formation with a new neighboring molecule, and these particular angles arise from the symmetry of the monolayer.

There is a barrier for rotation of the middle ring from $\theta = 0$ to $\theta = 60^\circ$ of $\sim 150$ meV, while the barrier from $60^\circ$ to $0^\circ$ is $\sim 65$ meV. We postulate that, as a voltage pulse is applied, these barriers may change, and a transition to a new
molecular conformation with the middle ring twisted may be induced. We have selected a candidate conformation, which we designate monolayer C@60, formed by twisting the middle ring of molecules in monolayer C by 60°, for possible stabilization under applied bias. The transmission spectra for monolayer C@60 is shown in Fig. 4(a) and its I-V$_b$ characteristics in Fig. 4(d). The conductance $G=0.12$ $\mu S$ is roughly 16 times smaller than monolayer C, and the current at $V_b=2$V is approximately five times smaller than in C. To lowest order, $G$ is proportional to the product of the matrix elements between $\pi$ orbitals on neighboring phenyl rings so that a rotation of the middle ring will reduce $G$ by a factor $\sim \cos^4(\theta)=\frac{1}{8}$, in rough agreement with the calculated values. The transmission eigenstates for electrons traveling from left to right at $E_{\text{HOMO}}=0.85$ eV and $E_{\text{LUMO}}=0.99$ eV are illustrated in Fig. 4(b). The HOMO resonance is localized on the phenyl ring nearest to the left electrode while the LUMO state corresponds to an orbital of the middle ring, weakly coupled to the first and last rings.

To investigate whether monolayer C@60 could be stabilized with respect to monolayer C as the bias voltage is increased, we show in Fig. 4(b) the Kohn-Sham total energy of monolayers C and C@60. At $V_b=0$, monolayer C is lower in energy than monolayer C@60. However, when the bias is increased above 2 V, monolayer C@60 becomes lower in energy, leading to an expected conformational change. While there are issues surrounding the evaluation of total energies for such systems,$^{17,23}$ Fig. 4(b) serves to illustrate that different conformations may be stabilized with respect to each other by an external bias. In this case, the calculation suggests a transformation of monolayer C to monolayer C@60 at around 2V. The exact value of the switching voltage will depend on the coverage, detailed structure, and size of the monolayer, as well as how the potential drops across the molecules, which can be affected by electrode coupling.$^{14}$

The local minima in Fig. 3(a) are stabilized by the NO$_2$ side group, which correlates well with that the molecular memory effect is only observed in monolayers with NO$_2$ side groups.$^4$ We find an energy barrier for rotation within an ideal (3 $\times$ 3) monolayer of 65–150 meV. This energy barrier will depend on the distance between the molecules in the monolayer, and we found that when doubling the distance between the molecules the barrier disappears. Most likely, the distance between the molecules in the nanopore experiments$^2$–$^5$ and not in experiments of single molecules. In the scanning tunneling microscopy experiments of Weiss and co-workers$^7$ on single or small bundles of molecules, the conductance switching may be related to different

![FIG. 3.](image1.png) **FIG. 3.** (Color online) (a) Energy vs rotation angle for molecules A, B, and C in Au(111) 3 $\times$ 3 unit-cell. The energy is calculated within the Perdew, Burke, and Ernzerhof approximation for the exchange-correlation functional (Ref. 21). (b) Contour plot of the effective potential between TW’s with NO$_2$ side groups. Note the bond formation between the O atom and the H atom on the neighboring TW.

![FIG. 4.](image2.png) **FIG. 4.** (Color online) (a) Zero-bias transmission $T(E,V_b=0)$ for monolayer C@60. (b) Transmission eigenchannels corresponding to HOMO and LUMO resonances of C@60. (c) Total energy of monolayer a C (dark gray or blue) and monolayer C@60 (light gray or magenta) as a function of bias potential. (d) I-V$_b$ characteristics of monolayer C (dark gray or blue) and monolayer C@60 (light gray or magenta).
conformations formed by steric interactions between the TW’s and the insulating alkyl molecules. This could explain why the observed switching behavior was independent of the functionalization of the molecules.

In conclusion, we find that functionalization of TW’s has a stronger effect on the energetics of the monolayers than on the orbitals responsible for current transport, and a better understanding of the intermolecular interactions in such molecules could hopefully be exploited in order to design molecular electronic devices with specific properties.

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12 In this study we have used both the local-density approximation (LDA), Refs. 25 and 26, and the generalized gradient approximation of Peredw, Burke, and Ernzerhof (PBE), Ref. 21. We find that the electrical properties are rather similar in PBE and LDA, while structural energies are slightly different. Unless otherwise stated, the presented results are obtained within LDA.
15 We use a double $\xi$–polarization basis set for the organic molecule, and single $\xi$ for the gold $s$, $p$, and $d$ channel. Core electrons are described by Troullier-Martins pseudopotentials. The molecular geometry is obtained by first optimizing the geometry of the free molecule with H atoms attached to the sulfur atoms, then placing the molecule at the fcc positions with an Au-S distance of 2.45 Å, Ref. 27, and finally relaxing the molecule in the sandwich structure with fixed Au atoms.
18 Since the charging energy is inversely proportional to the distance between the molecules, we expect that doubling the distance between the molecules can at most double the charge on the molecules.
23 The statistically most likely geometry is that which minimizes the free energy: $E - \mu_L Q_L - \mu_R Q_R$, where $Q_L$ and $Q_R$ are the charges injected from the left and right electrodes (Ref. 17). The quantities $Q_L$ and $Q_R$ are, for technical reasons, very expensive to calculate. Furthermore, in order to conclusively determine which geometry has the lowest free energy, we should allow for relaxations of the atomic coordinates under the external bias.