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Electrochemical Deposition of Aluminum from NaCl-AlCl₃ Melts

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ABSTRACT

Electrochemical deposition of aluminum from NaAlCl₃ melts saturated with NaCl onto a glassy carbon electrode at 175°C has been studied by voltammetry, chronoamperometry, and constant current deposition. The deposition of aluminum was found to proceed via a nucleation/growth mechanism, and the nucleation process was found to be progressive. The morphology of aluminum deposits was examined with photomicroscopy. It was shown that depending on the current density, smooth layers deposited at intermediate current density (below 0.7 mA/cm²), dendritic or porous deposits were formed due to difficulties in electronucleation and could be inhibited by application of pulsed currents and/or addition of manganese chloride into the melt.

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tion of species containing aluminum (12-14). Generally, in melts with less than 57 mole percent (m/o) AlCl₃, the critical current density is very small and dendrites are easily formed. This finding is unfortunate, since at high concentrations of AlCl₃, the melts are difficult to handle because of their high vapor pressures.

There were also reports suggesting that electrode subtrates (8, 11, 15), inorganic and organic additives (10, 13, 16, 17), and superposed alternating current (8) or pulsed current (15) had significant effects on inhibiting aluminum dendrite formation. These findings cannot, of course, fit the mechanism of inhibition of dendrite formation but suggest the governing role of the formation of crystalline nuclei in the deposition (14, 15). In the overvoltage measurement in a NaCl-KCl-AlCl₃ melt, Hayashi et al. (18) found that the main part of cathodic polarization seemed due to crystallization overvoltage in the low polarization potential region. Rolland and Mamantov (19) found that the reduction of AlCl₃ in acidic chloroaluminate melts involved a nucleation process. Apart from this, little work has been done on the electrochemical nucleation and growth of aluminum from that medium. Moreover, almost all previous studies were performed in acidic NaCl-AlCl₃ melts, and little work has been reported in the basic NaAlCl₄ melt saturated with NaCl. This system is of much more interest from the technical point of view because of its low vapor pressure. It was our intention to study the electrochemical nucleation process and to examine how and when the aluminum dendrites form in NaCl-saturated melts at 175°C.

Experimental

Preparation of chemicals.—Purification of AlCl₃ and NaCl and other experimental procedures have been previously described (3, 20). The equimolar NaCl-AlCl₃ melts, which were made from distilled AlCl₃ and analytical grade NaCl, were dryed at 200°C, became slightly yellow after several days of experimentation and some dark particles were always formed both on the aluminum electrode and in the melt. Further purification of the equimolar metals was therefore performed by zone refining carried out in a glass furnace consisting of nine stationary ring heaters. A certain amount of a mixture of 50.1 m/o distilled AlCl₃ and 49.9 m/o dried NaCl was charged into a Pyrex ampul of 200 mm length and 25 mm inner diameter. The ampul was then sealed off under vacuum and heated overnight in a rocking furnace. The heating zones of 35 mm width each were maintained at 150°C. The motion of the charged ampul led to a displacement of 75 mm which was repeated many times. The travel rate of the molten zones was 3 mm/h. The duration of the whole refining procedure was about 2 weeks.

After zone refining, the molten salts were colorless and allowed aluminum electrodes in the subsequent experiments to remain shiny and bright. However, a tiny amount of dark particles still appeared in the melt after some days of use.

Cell assembly and oven.—To facilitate the visual observation of the aluminum deposits, the cells were made of square Pyrex tubes. The working electrode was made of a glassy carbon rod (Type V10 from Le Carbone Lorraine) of 3 mm diameter (area 0.07 cm²), as described earlier (3). Two aluminum rods of 99.999% purity joined to tungsten leads were used as the counter and reference electrode, respectively. The reference electrode was placed in the cell in such a way that it was surrounded mainly by NaAlCl₄ saturated with NaCl (i.e., the composition of the melt before an experiment started). The test cell was placed into a transparent oven of our own construction, heated with circulated air. The aluminum deposits could be observed and photographed directly.

Instruments and measurements.—Voltammograms and potentiostatic current transients were obtained with an electrochemical system built in this laboratory (21) and recorded on a Hewlett-Packard X-Y recorder (Type 7004B). After each set of measurements, the working electrode was kept at a positive potential of 100 mV vs. aluminum for at least half an hour and then left for another half hour to eliminate the concentration polarizations.

Aluminum deposits were achieved by electrolysis at constant current, which was delivered by a chronamperogram built in this laboratory. To ensure that sufficient metal was deposited to provide a visible layer of aluminum deposits, a total charge of ca. 30 C/cm² was employed in each run, corresponding theoretically to a 0.01 mm thickness of uniformly plated aluminum. The duration of each deposition experiment varied from several minutes to more than 20h, depending on the current density applied. After deposition, the aluminum deposits were observed by using a Zeiss Jena Technical 2 Stereo Microscope equipped with an Olympus OM-2 SLR camera and a Schott KL 1500 cold light source. The morphology of the aluminum deposits was examined by microphotography and the specific volume of the smooth deposits was measured as averages from the pictures.

Results and Discussion

Voltammograms.—Figure 1 shows voltammograms obtained in both slightly acidic (Fig. 1a) and slightly basic (saturated with NaCl, Fig. 1b) NaCl-AlCl₃ melts at 175°C. The small current peak A in Fig. 1a may be attributed to the AlCl₃ reduction which starts at ca. -50 mV vs. the aluminum reference. Rolland and Mamantov (19) found the same overpotential value in their electrochemical study. The peak was followed by the AlCl₃ reduction. In the voltammogram from the basic melts (Fig. 1b), the current peak disappeared. In this case only AlCl₄⁻ reduction occurred starting at ca. -65 mV vs. aluminum. In the cathodic branches of both voltammograms, cross-over loops appeared on reversing sweeps characterizing the presence of a nucleation/growth process. Integration of both the cathodic and the anodic currents of the voltammograms obtained from the basic melts gave a charge ratio, \( \frac{Q_{Al}}{Q_{Cl}} \), between 1.01-1.05, indicating that the reduction of AlCl₄⁻ behaves reversibly in a chemical sense.

Chronoamperograms.—Figure 2 shows a set of potentiostatic current-time transients obtained at different overpotentials in NaAlCl₄ saturated with NaCl. For all overpotentials applied, there appears to be a peak current transient corresponding to the charging of the electric double layer. For overpotentials below 65 mV vs. the Al reference, there is no reduction current. A nucleation overpotential of at least 65 mV is required for any significant electrodeposition to take place. This is in agreement with the overpotential observed in voltammetric measurement for the same electrode and the same NaCl-saturated melt. For overpotentials larger than 70 mV, the capacitive charging current was followed by a rising current due to the formation of new nuclei. As the nuclei grew, the overlap of nuclei or neighboring diffusion zones gave rise to a current maximum followed by a falling portion of current. The time to reach the maximum current decreased as the applied overpotentials increased (Fig. 2a).

![Fig. 1. Voltammograms on glassy carbon electrodes at 175°C; sweep rates were 10 mV s⁻¹. (a) 49.7 m/o NaCl-50.3 m/o AlCl₃ melt. (b) NaAlCl₄ saturated with NaCl.](image-url)
Fig. 2. Potentiostatic current transients for deposition of aluminum onto a glassy carbon electrode (3 mm diam) from a NaCl-saturated melt (a) at low overpotentials (mV), and (b) at high overpotentials (mV). Applied overpotentials (mV) vs. aluminum reference are indicated in the figure. For clarity not all measured current transients are shown. Electrode area, 0.07 cm$^2$, temperature, 175°C.

When large overpotentials are applied, the time required to reach the maximum current becomes so short that the capacitive current peak is not distinguishable from the rising portion of the current transients due to the faradaic process. After the first maximum in the transients, the current rises again, and a second current maximum appears thereafter when overpotentials larger than 120 mV are applied (Fig. 2b). As nucleation proceeds continuously, new nuclei are formed on top of the first deposited layer long before it has covered the whole area of the electrode surface, i.e., multinuclear multilayer growth occurs (22).

For a three-dimensional nucleation, expressions for the theoretical current transients describing both the rising and the falling parts have been obtained for both instantaneous and progressive nucleations (23, 24). A convenient criterion for distinguishing between the two mechanisms is to represent the experimental data in the dimensionless plot, $(i/m)^2$ vs. $t_{real}/t_m$, where $i_m$ and $t_m$ are the coordinates of the current maximum on the current transients. This analysis can be made by comparing the experimental data with theoretical plots resulting from Eq. [1] and [2] for instantaneous and progressive nucleations, respectively (24)

$$(i/m)^2 = 1.9542 (t/t_m)^{-1.5} \{1 - \exp [-1.2564 (t/t_m)]\}^2 \tag{1}$$

$$(i/m)^2 = 1.2254 (t/t_m)^{-1} \{1 - \exp [-2.3367 (t/t_m)]\}^2 \tag{2}$$

It is found that the dimensionless plot of our experimental data fits to the theoretical curve corresponding to progressive nucleation except for low values of $t/t_m$. As demonstrated (25), the experimental current transients are affected by a certain time delay, $t_0$. The exact comparison is possible only after correction of the time scale by defining

$$t_{real} = t - t_0 \tag{3}$$

Figure 3 shows the graphical analysis of the rising portion of current transients. A linear dependence of $i^{2/3}$ on time is found, indicating the progressive nature of the nucleation process as required by the following equation (25)

$$i = \frac{2}{3} zF (2Dc)^{1/2} (M/p)^{1/3} N_t^{1/3}$$

Here $zF$ is the molar charge of the deposited species, $D$ is its diffusion coefficient, $c$ is its concentration in mol cm$^{-3}$, $M$ is its molecular weight, and $p$ is the density of the electrolyte. This is in accordance with the conclusion obtained from the dimensionless analysis. The straight lines of $i^{2/3}$ vs. $t$ intercept the time axis at $t_{real} = t - t_0 = 0$, i.e., $t = t_0$. Therefore, the delay time, $t_0$, needed for the time correction can be obtained. After correction of the time scale with $t_0$ obtained this way, the dimensionless plot of the experimental data is well superimposed on the theoretical curve, corresponding to a progressive nucleation process (Eq. [2]), as shown in Fig. 4.

Aluminum dendrite formation.—Aluminum deposits obtained by electrochemical reduction from NaA1C14 melt saturated with NaCl can vary greatly in their outer appearance depending on the depositing conditions. They can be smooth, rough, dendritic, or porous as the current densities increase. Several types of surface morphology of alu-


dendrites are shown. Electrode area, 0.07 cm$^2$, temperature, 175°C.

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variation of aluminum grain sizes with current density of thus be attributed to the lack of nuclei originally formed that the grain size is determined primarily by the number electronucleation of aluminum in the deposition pro-
ing study (14). The assumption of the predominant role of cleation/growth mechanism. The difficulty in electronucle-
lation of the spongy deposits at lower current density can progressive character of the nucleation process.

A critical thickness of aluminum deposits, as well as the variation of aluminum grain sizes with current density of deposition, was also observed in an aluminum electroplating study (14). The assumption of the predominant role of the electronucleation of aluminum in the deposition pro-

All of the results cited were obtained by using iron or steel substrates. Therefore the difference in critical current densities between the cited literatures and ours may be partly attributed to the different composition of the melt and to the different electrode substrates used. Moreover,
Nayak et al. (11) found that electrodeposition of aluminum on brass from 36.4 m/o NaCl-63.6 m/o AlCl3 led to non-dendritic crystal deposits at current densities of up to 53 mA/cm². Pretreatment of a mild steel substrate by electropolishing in the electrolyte itself was also found to be very helpful in eliminating dendrite formation (12). The fact that the substrates have such a pronounced effect on aluminum deposition also suggests the importance of the nucleation process. It seems correct to say that even in the fact that the substrates have such a pronounced effect on the electronucleation process may play a certain role in dendrite formation.

The smooth aluminum deposits obtained at intermediate current densities between 2 and 10 mA/cm² are not compact. Figure 7 shows the results of the average volume of the aluminum deposits per coulomb vs. the current density that was applied. The volumes were all measured on photographs of the deposits. The deposits obtained from NaAlCl₄ saturated with NaCl by constant current deposition look porous. The values for the volume of aluminum are about five times or more larger than the theoretical value of 0.0345 mm³/C for a compact layer of aluminum. The larger values obtained at low and high current densities during constant current deposition reflect the formation of two types of deposits, viz., the electronucleation-controlled formation of spongy deposits at low current densities and the diffusion-controlled formation of dendrites at high current densities.

It seems of importance to distinguish the two mechanisms from one another. Consequently, different techniques may be used to inhibit the formation of individual kinds of dendrites. For instance, to avoid dendritic formation caused by the slow nucleation process at low current density, a pulsed current shown in Fig. 8 was used. A single pulse of current (i₀) ten times as large as the deposition current (iₑ) and with a duration of 10s was applied at the beginning of the experiment in order to accelerate the nucleation process. A small deposition current (iₑ) was then used to avoid the dendrite formation caused by the diffusion-controlled process afterward. In this way, the spongy deposits can be completely eliminated. The volume of the aluminum deposits obtained by the pulsed current deposition is also shown in Fig. 7. It can be seen that the compactness of the aluminum deposits was much improved by this pulse technique, especially at low current densities. Apparent differences between aluminum deposits obtained at the same deposition current density (1.14 mA/cm²) with and without the initial current pulse are shown in Fig. 6b and c. Furthermore, the addition of manganese chloride into the melt can also significantly improve the quality of deposits, as seen in Fig. 6d. Detailed studies are now under way to explore the effect of manganese chloride addition on aluminum deposition.

Conclusions

Electrochemical deposition of aluminum onto glassy carbon electrodes from NaAlCl₄ saturated with NaCl at 175°C has been investigated and found to proceed via a nucleation/growth mechanism. The nucleation process was found to be progressive. Several types of aluminum deposits can be electrodeposited from the melt, mainly depending on the current densities applied. It is observed that only at intermediate current densities, between 2 and 10 mA/cm², can a smooth layer of aluminum be formed. The smooth deposits are still about five times more voluminous than the theoretical value. At high current densities, above ca. 15 mA/cm², needle-like or fir-tree-like dendrites or very porous deposits were often obtained. The deposits were attributable to the slow diffusion process of species containing aluminum. At lower current densities, below ca. 0.7 mA/cm², aluminum was deposited in the form of spongy deposits, apparently due to the slow nucleation process. The spongy deposits can be inhibited by using pulsed current deposition or by adding manganese chloride to the melt.

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REFERENCES

Analysis of the Response of Ion Sensing FETs with a Chemically Modified Gate Insulator

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ABSTRACT

This paper is devoted to the analysis of the physicochemical parameters influencing the response of ion sensing FETs obtained by chemical grafting of the silica gate insulator. The site binding theory has been applied to the modified silica/electrolyte interface; for the first time the equilibria involving H⁺ and M⁺ ions have been taken into account. The established theoretical expression ψ = ΔpM shows that the sensitivity of the ion sensing FETs is better when the number of sensitive sites and the complexation constant of these sites are higher. The experimental response of the Ag⁺ sensing ISFET was fitted with this model. The percentage of cyanografted sites is determined to be 8%; their complexation constant is ~0.9.

Site Binding Theory Applied to Grafted Silica Surface

Surface reactions.—The two types of sites on the grafted silica surface are amphoteric hydroxyl sites, called SOH, and the ionophore grafted sites called S-X. It is assumed that no surface complexation of M⁺ occurs with hydroxyl groups, as has been observed in the case of Ag⁺ ions, the flatband potential of SiO₂/electrolyte only being slightly shifted in the presence of Ag⁺ ions (10). The S-X sites are assumed to present no amphoteric properties as has been shown through the electrokinetic properties of CN-grafted silica in colloidal aqueous suspensions (16). Thus, each type of site reacts with one specific ion: H⁺ with the amphoteric SOH sites and M⁺ with the S-X sites. The formation of surface complexes between charged surface sites (SOH⁺ and SO⁻) and counterions in the solution will be neglected. Such surface complexes have been shown to have a negligible effect on the surface potential (14). The equations describing the surface reactions of the two sites are:

Amphoteric reactions of SOH.—

\[
\text{SOH}^+ + \text{H}^+ \rightleftharpoons \text{SOH} + \text{H}_2\text{O} \quad \text{(1)} \\
\text{SOH} + \text{H}_2\text{O} \rightleftharpoons \text{SOH}^+ + \text{H}_2\text{O} \quad \text{(2)}
\]

Surface complexation of M⁺ with S-X sites.—

\[
\text{S-X} + \text{M}^+ \rightleftharpoons \text{S-X} \cdot \text{M}^+ \quad \text{(3)}
\]

In these equations, the quantities in parentheses are ionic activities, in the case of bulk concentrations, or numbers of sites per unit surface area in the case of surface concentrations. The activity coefficients for the surface con-

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